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# o-Phenylenediamine encapsulated silver nanoparticles and their applications for organic light-emitting devices

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## 1. Introduction

Quantum-size (<10 nm) metal nanoparticles (NPs) have attracted a great deal of interest for a variety of applications (Gittins & Caruso, 2001; Brust & Kiely, 2002), many of which require that the particles be water-soluble (Ackerson, et al. 2005; Kim, et al., 2005; Bao, et al. 2003; Manna, et al. 2001) or remain suspended in water with no loss of their physical or chemical properties over extended periods of time. In addition, many applications require colloidal precipitation (Li, et al. 2005; Pan & Rothberg, 2005; Haes & Van Duyne, 2004) of these NPs to allow them to be film-cast using suitable solvents. The choice of synthetic methods and precursor elements are vitally important for determining the phase (water-soluble or colloidal precipitation) of the NPs. Water-based syntheses of NPs, however, are fraught with problems that result from ionic interactions, which are typically overcome by using very low reactant concentrations (ca.  $1 \times 10^{-4}$  M) (Ahmadi, et al. 1996). Moreover, diffusion, which is much faster in aqueous solutions, can lead to aggregated metal NPs in the absence of a stabilizer (Rele, et al. 2004). In contrast, particles synthesized in suitable organic solvents can be prepared at relatively high concentrations of reactants (Sun & Xia, 2002), with predefined sizes and shapes (Gree & O'Brien, 1999; Pileni, 1998), and with improved monodispersity relative to that of NPs prepared in aqueous solutions.

Here, we describe the preparation of stable, monodisperse, quantum-sized silver nanoparticles (NPs) in water soluble form through a simple one-pot procedure that uses glycerol for reducing  $\text{Ag}^+$  in the presence of o-phenylenediamine (o-PDA) (Habib, et al., 2006).

On the other hand, organic light emitting devices (OLEDs) have attracted great attention for flat panel display (FPD) applications due to their sophisticated device structure that is simpler than liquid crystal displays (LCD) as well as their fast response time and wide viewing angle which are similar to the typical cathode ray tube (CRT) display (Hung & Chen, 2002).

OLEDs are commonly fabricated on glass substrates coated with a thin film of indium tin oxide (ITO). The ITO glass is generally employed as the anode in OLEDs because of its good

electrical conductivity, high transparency over the visible region, ease of patterning, and relatively high work function. However, it has been suggested that the inefficiency of hole injection from bare ITO, poor adhesion at the ITO-organic film interface, and the diffusion of materials across the ITO-organic interface have led to the poor device performance of OLEDs (Kim, et al. 1998; Kugler, et al. 1998; Mason, et al. 1999). As a result, great effort has been made to modify the ITO anode, such as various treatments of the ITO surface and the deposition of very thin buffer layers between the ITO and the hole transport layer in order to enhance the OLED performances. In particular, poly(3,4-ethylene dioxythiophene)/polystyrene sulfonic acid (PEDOT/PSS) is in the limelight because of high conductivity, stability, and transparency (Groenedaal, et al. 2000). In this aspect, we also investigated the performance of OLEDs by systematic incorporation of these Ag NPs in PEDOT/PSS matrix, which is used as a hole-injection layer for OLEDs (Park, et al. 2007).

## 2. Experimental

All chemical used were of analytical grade and purchased from Sigma-Aldrich. In a typical synthesis, *o*-PDA flakes (0.40 g) were dissolved in glycerol (24 mL) by heating at 140–160 °C for 30 min under vigorous magnetic stirring. The formation of a transparent light-green solution (0.154 M) indicated that *o*-PDA had completely dissolved. Meanwhile, AgNO<sub>3</sub> (1.0 g) was dissolved in glycerol (8 mL) over 15 min at 90–100 °C under magnetic stirring; once a clear solution (0.735 M) had formed, the stirring was continued for another 15 min. The hot colloidal Ag(0) solution was then added dropwise continually into the solution of *o*-PDA. Heating of the reaction mixture was continued at 140–160 °C for 15 min and then the reaction mixture was quenched in cold (0–2 °C) water [1.6 L; 50 times the solvent volume (32 ml)]. We adopted a direct quenching technique (pouring the reaction mixture in cold water) to prevent aggregation because of highly concentrated particles being aggregated due to slow cooling. The colloidal mixture was maintained under static conditions for 48 h (room temperature) to effect complete precipitation. The precipitated product was separated by decantation and washed several times with water. The decanted silver solution was then subjected to centrifugation, but no further precipitation or phase separation occurred. We termed this phase as a ‘water-soluble’. Details of the single-phase solution product are described elsewhere (Habib, et al., 2006).

PEDOT/PSS (2.8 wt% dispersed in H<sub>2</sub>O) was purchased from Aldrich. It was filtered through a 0.45 µm membrane filter (Millipore, PVDF) before preparing a hybrid solution. Tris(8-hydroxyquinolino)aluminum (Alq<sub>3</sub>) and *N,N'*-diphenyl-*N,N'*-bis(3-methylphenyl)-1,1-biphenyl-4,4'-diamine (TPD) were purchased from Tokyo Kasei Kogyo Co., Ltd. and used without further purification.

Hybrid PEDOT/PSS–Ag solutions were prepared by dispersing a given mass ratios of *o*-PDA encapsulated Ag NPs in the aqueous PEDOT/PSS solution in an ultrasonic bath for 4 h. Three different weight ratios (5, 10, and 20 wt%) of these Ag NPs were employed to investigate the influence of composition. For the fabrication of OLEDs, ITO coated glass substrates were patterned in a strip of size (2 × 30 mm) and were cleaned ultrasonically in deionized water using a non-phosphorus detergent and then rinsed with ethanol before drying. A layer of PEDOT/PSS–Ag nanocomposite (thickness = 40 nm) was spin-coated onto the pre-cleaned ITO coated glass at 4000 rpm for 60 s and then soft-baked at 100 °C for 6 h in a vacuum oven. On top of the PEDOT/PSS–Ag nanocomposite layer, a 50 nm TPD

layer and a 60 nm thick Alq<sub>3</sub> layer were deposited as a hole transporting layer and an emission layer, respectively, and finally cathode layer of aluminum (Al) was deposited by vacuum evaporation ( $5 \times 10^{-6}$  Torr) technique (as shown in Fig. 1), defining the active emission area of 4 mm<sup>2</sup>. For comparison, the device with the following structure was also prepared; PEDOT/PSS (thickness = 40 nm) / TPD (thickness = 50 nm)/ Alq<sub>3</sub> (thickness = 60 nm)/ Al (see Fig. 1).

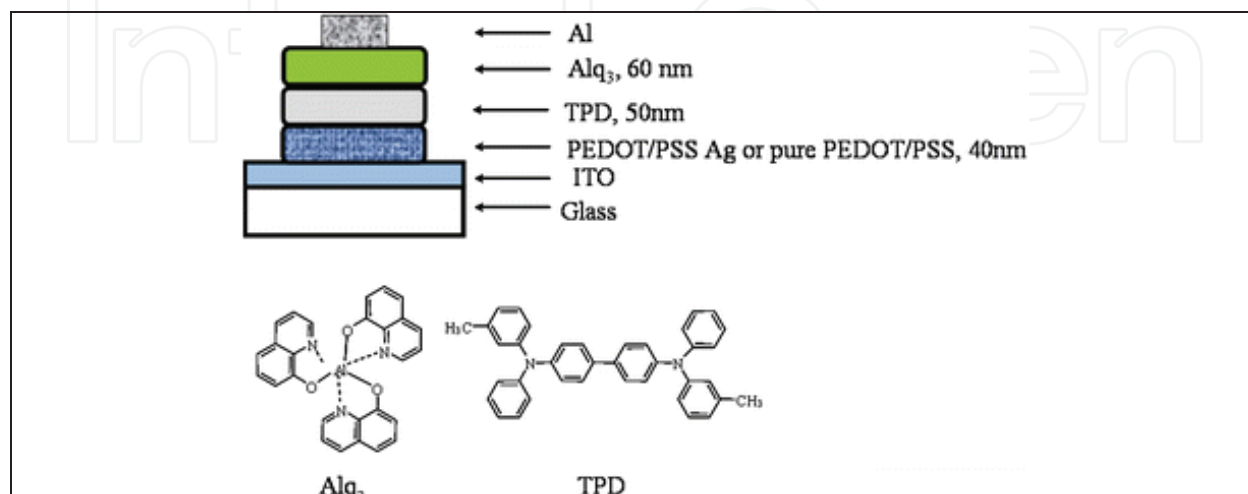


Fig. 1. Schematic cross-sectional structure of OLED used in this study and chemical structures of Alq<sub>3</sub> and TPD (Park, et al. 2007).

The current density-voltage-luminance (J-V-L) characteristics of OLEDs were measured using a customized device measurement system equipped with a photomultiplier tube (PMT, Hamamatsu Photonics Co.) and an electrometer (SMU 236, Keithley). All devices were mounted in a dark sample chamber for the J-V-L measurements in order to get rid of any influence of ambient light.

### 3. Results and Discussion

Figure 2 displays a typical TEM image of the water-soluble Ag NPs. The nanoparticles exhibit a bimodal distribution, with distinct size ranges of 2-3 and 5-6 nm (Habib, et al.) (Habib, et al., 2006). We verified the TEM results through an analysis of the sample's UV-Vis absorption spectra. The water-soluble Ag NPs displayed a strong plasmon absorbance ( $\lambda_{\text{max}} = 448$  nm) possessing a bimodal-type pattern (Taubert, et al. 2003); XRD data (Figure 3) of Ag NPs displayed strong Bragg reflections corresponding to the (111), (200), (220), and (311) reflection planes of fcc metallic silver.

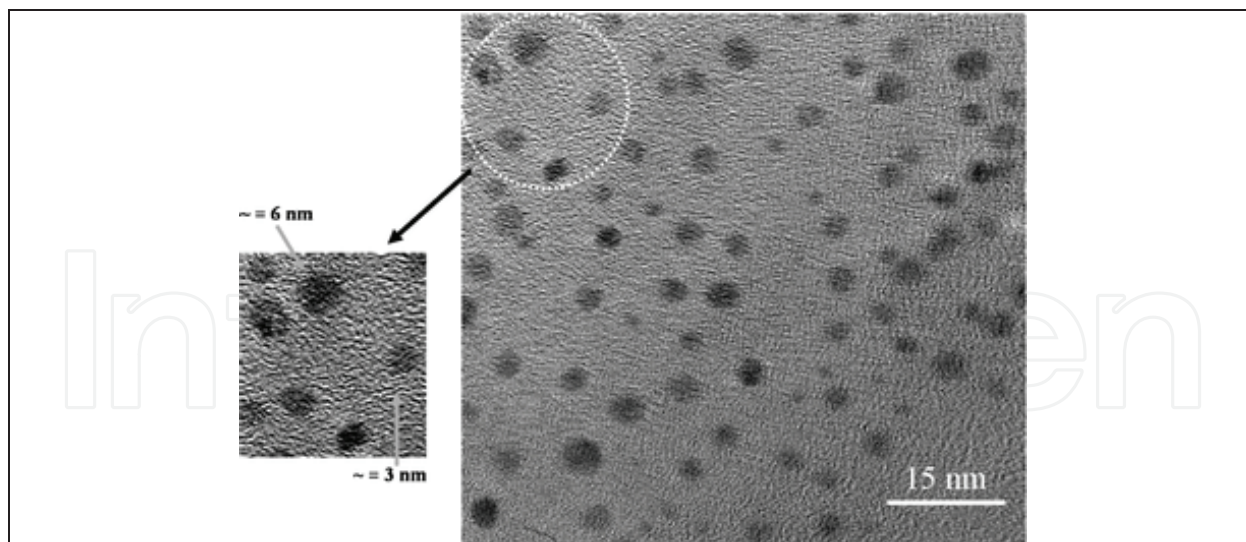


Fig. 2. TEM image of the water-soluble Ag NPs (Park, et al. 2007).

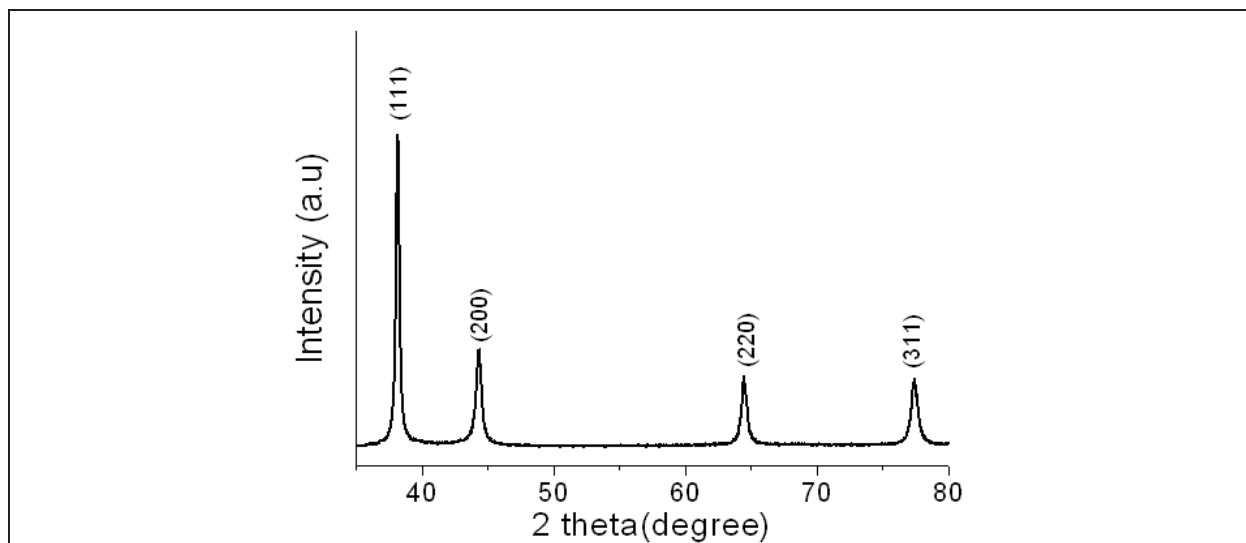


Fig. 3. XRD pattern of Ag NPs, displaying strong Bragg reflections corresponding to fcc metallic silver (Habib, et al. 2006).

The UV-Vis spectrum (Figure 4) of silver sol (prior to addition into the *o*-PDA solution) indicates that initially the particle sizes are <2 nm (Wilson, et al. 2005; Kim, et al. 2004), they may be stabilized temporarily through complexation with hydroxypyruvic acid (i.e., oxidized glycerol) and glycerol (Rele, et al. 2004). Relatively larger particles were formed after adding the mixture into the *o*-PDA solution. This finding may indicate that when the small (<2 nm) particles entered into the comparatively hot (140–160 °C) *o*-PDA solution, they became uncapped from their temporary complexes and underwent a melting-like process to form relatively larger particles (Pastoriza-Santos & Liz-Marzan, 2002).



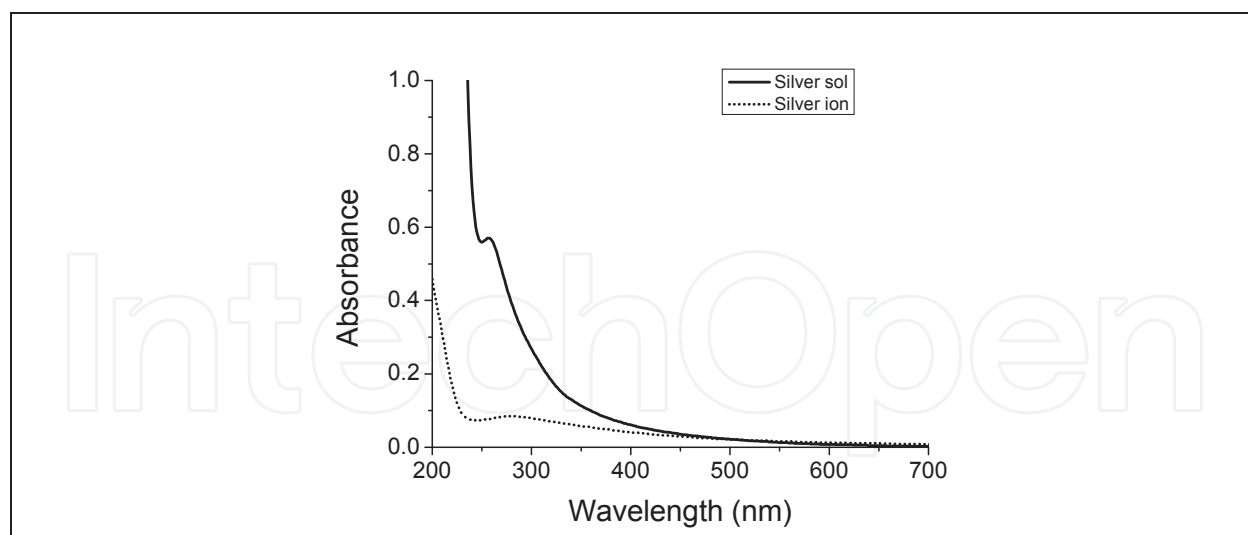


Fig. 4. The UV-Vis spectrum of silver sol prior to addition into *o*-PDA solution, which exhibits a monotonically increasing absorbance toward higher energies because, for very small particles, the distinctive plasmon band is replaced by a featureless absorbance, which has been compared with silver ion sol (Habib, et al. 2006).

Figure 5 presents the UV-vis absorption spectra of the PEDOT/PSS and the PEDOT/PSS-Ag nanocomposites with different concentrations of Ag NPs. *o*-PDA encapsulated water-soluble Ag NPs display a surface plasmon absorption band centered at 448 nm, which is large red shift compared to the uncapped metallic silver that has an intense plasmon absorption band in the visible region at  $\approx 390$  nm (He et al. 2001; Kumar et al. 2003). The UV-vis absorption coefficient (at wavelength  $\approx 450$  nm) of the PEDOT/PSS-Ag nanocomposite increases with increasing the concentration of Ag NPs, where pure PEDOT/PSS does not show any plasmon band. This result indicates that *o*-PDA encapsulated silver NPs are well embedded in the composite layer.

Current density (J)-voltage (V) and luminance (L)-voltage (V) characteristics of various devices fabricated in this work are shown in Fig. 6 (Park et al. 2007). As expected, the TPD/Alq3 device (see (a) curves in Figs. 6A and 6B) without PEDOT/PSS layer exhibits the poorest performance among the devices. It is noteworthy that the performance of the OLEDs is significantly improved (see (c), (d), and (e) curves in Figs. 6A and B) when the PEDOT/PSS-Ag nanocomposite layer (with the loading concentrations of Ag NPs in the range of 5–20 wt%) is used compared to the OLED (see (b) curves in Figs. 6A and B) using PEDOT/PSS only as a hole injection layer. The devices with 10 or 20 wt% Ag NPs and PEDOT/PSS in the hole injection layer exhibited high brightness ((Fig. 6B). The turn-on voltages, 6.0, 7.8, 8.0, and 8.7 V, were obtained for the loading of 20, 10, 5, 0 wt% Ag, respectively; where the turn-on voltage, 14 V was measured for the OLED without PEDOT/PSS. These results demonstrate that the turn-on voltage decreases with increasing the Ag NPs' concentration. In this work, the maximum luminance of the device fabricated with 20 wt% Ag nanoparticles in the PEDOT/PSS matrix as a hole injection layer was ca. 9,000 cd/m<sup>2</sup> at around 16 V (see (e) curve in Fig. 6B).

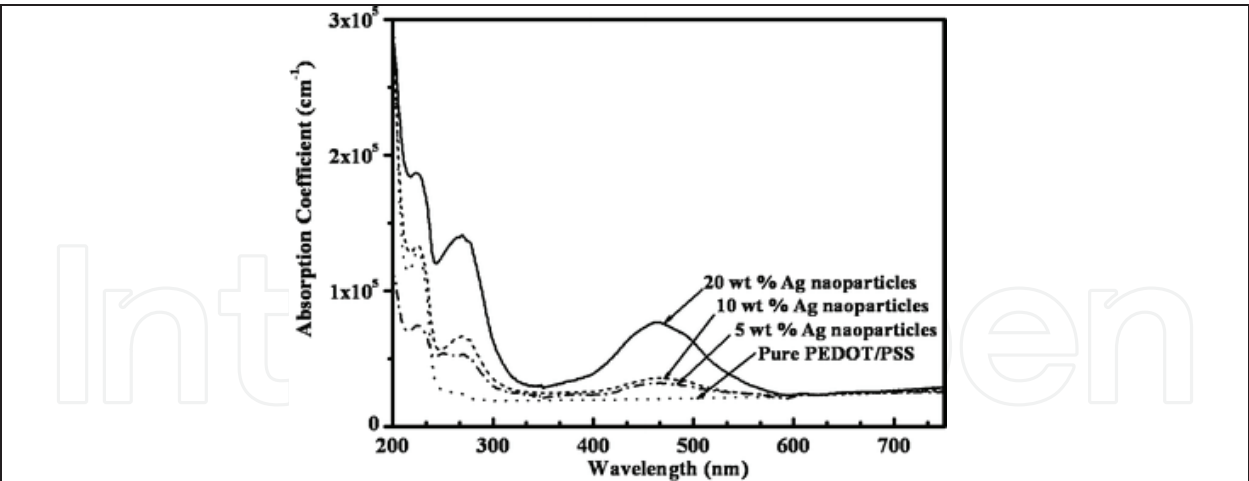


Fig. 5. UV-vis absorption spectra of the PEDOT/PSS and the PEDOT/PSS-Ag nanocomposites with different concentrations of Ag NPs in the PEDOT/PSS matrix (Park, et al. 2007).

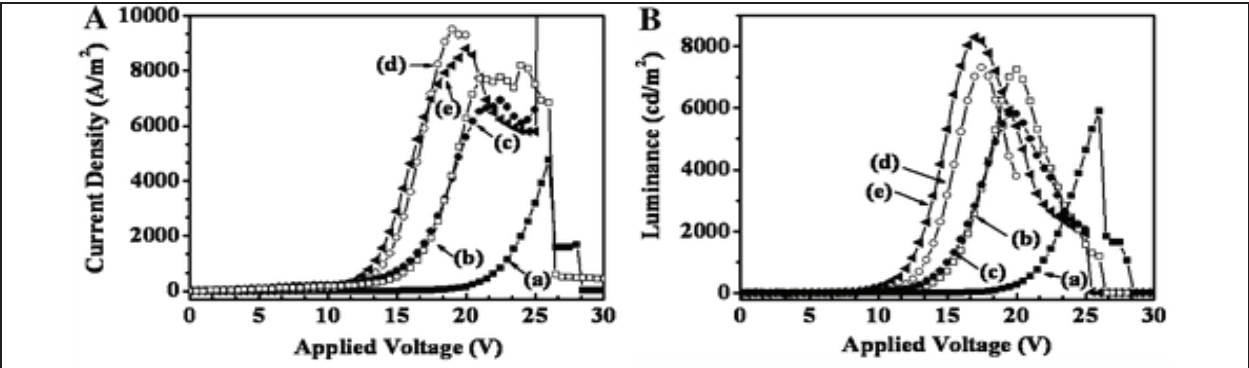


Fig. 6 (A) Current density-voltage and (B) luminance-voltage characteristics of OLEDs made with the PEDOT/PSS and the PEDOT/PSS-Ag nanocomposites with different concentrations of Ag NPs in the PEDOT/PSS matrix in the hole injection layer [(a) OLED without PEDOT/PSS, (b) OLED with PEDOT/PSS only, (c) PEDOT/PSS-Ag nanocomposite with 5 wt% Ag NPs, (d) PEDOT/PSS-Ag nanocomposite with 10 wt% Ag NPs, (e) PEDOT/PSS-Ag nanocomposite with 20 wt% Ag NPs] (Park, et al. 2007).

Further study is, however, needed in order to understand more exact relationship between the optimum concentration and size of the Ag nanoparticles and the device performances.

4. Conclusions

We have synthesized aggregation free quantum-size silver NPs using a high AgNO<sub>3</sub> concentration through a novel technique. The in-situ production of binary phases quantum-size Ag NPs having narrow size distributions will open the new avenue for the applied fields. The nanoparticles, especially, water-soluble Ag NPs are suitable for electronics applications because the  $\pi$ -conjugated aromatic amine molecules are well-coordinated with silver surface. In this sense, we investigated the performance of OLEDs consisting of PEDOT/PSS-Ag nanocomposites with different concentrations of quantum-size silver NPs as a hole injection layer. We found that the OLEDs with PEDOT/PSS-Ag nanocomposites

exhibit improved performance. The turn-on voltage is significantly decreased when increasing the Ag NPs' concentration. In this work, the maximum luminance of the device consisting of 20 wt% Ag NPs in the PEDOT/PSS matrix as a hole injection layer was ca. 9,000 cd/m<sup>2</sup> at around 16 V.

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