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Evolution of Crystallographic Structures and Phases in Micropyretically Formed Boron Rich Boron Carbide – a New Material System

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1. Introduction

The 'boron- carbon system' is a typical representative of hard refractory nonmetallic low Z-ceramic. Among various boron to carbon stoichiometry, the ratio higher than 2, are known as boron carbides. These offer many attractive advantage compared to any commercially available hard materials [1-2]. Beside the powder forms, the sintered shapes find wide spread application in ballistic, biomedical application and component engineering. In principle the binary B-C system exist with a number of stable and metastable configurations [3]. Many classical and non classical processes have been considered [4 -12] for producing boron carbides. Among them the researchers have reported on existence of rhombohedral boron carbide with stoichiometry ratio (B/C) near to 4:1. The widely adopted carbothermal process requires rapid heating of all particle separately and individually up to temperature range 2473K to 2773K to achieve significant conversion of input material in 5 hours to 24 hours. In almost all carbothermal processes it is reported to have agglomerated carbon rich boron carbide with boron value up to 78.8%. As per the literature surveys, hardly any systematic investigation has been reported for existence of multiple phases in a single step solid state processing path. It is possible that selection or tailoring of properties for particular use can be made on basis of different stoichiometric phases of boron carbide [13]. The presence of boron rich boron carbide can compensate the boron loss during boron carbide based solid body formation processes and thus preserves or improves the functional requirement. 'Vapor deposition' and 'boron doping' processes are the well studied processing methods which have confirmed the 'non stoichiometric existence of boron carbides' [3, 14,15]. With reference to above facts the research article discusses the structure and analytical development of many phases of boron carbide formed in a 'B₂O₃ - Mg - C' micropyretic process with bulk boron to carbon ratio as 5.042:1[16].

1.1 Crystal structure and stoichiometry

In variety of boron based systems, the structure of low boron containing systems are determined by the metal lattices or semi metallic lattices. For boron-rich compounds, the

structures are determined by their boron atoms which form covalent boron-boron lattices [17]. In these systems the non stoichiometry phenomenon arises from the existence of a broad free energy minimum in the free energy vs. composition diagram. This is facilitated either by the defects in boron chains or because of vacancies in the covalent lattices with a slight shift in lattice spacing [18-19]. The presence of these non-stoichiometric phases complicates the crystal structure elucidation. Different modifications of elementary boron and numerous compounds of it belonging to boron rich lattices exhibit a close similarity in their structures. The structure consisting of complicated three dimensional network of B12 or B11C (in case of B-C system) icosahedra or related aggregates of atoms as essential common structural elements. The structure of rhombohedral boron carbide is a modification of alpha-B12 structure [20 -21]. In general the structure of B-C system is an arrangement of regular do-decahedrons at the nodes of a rhombohedral Bravais lattice (R-3M space group) as shown in Fig.1.

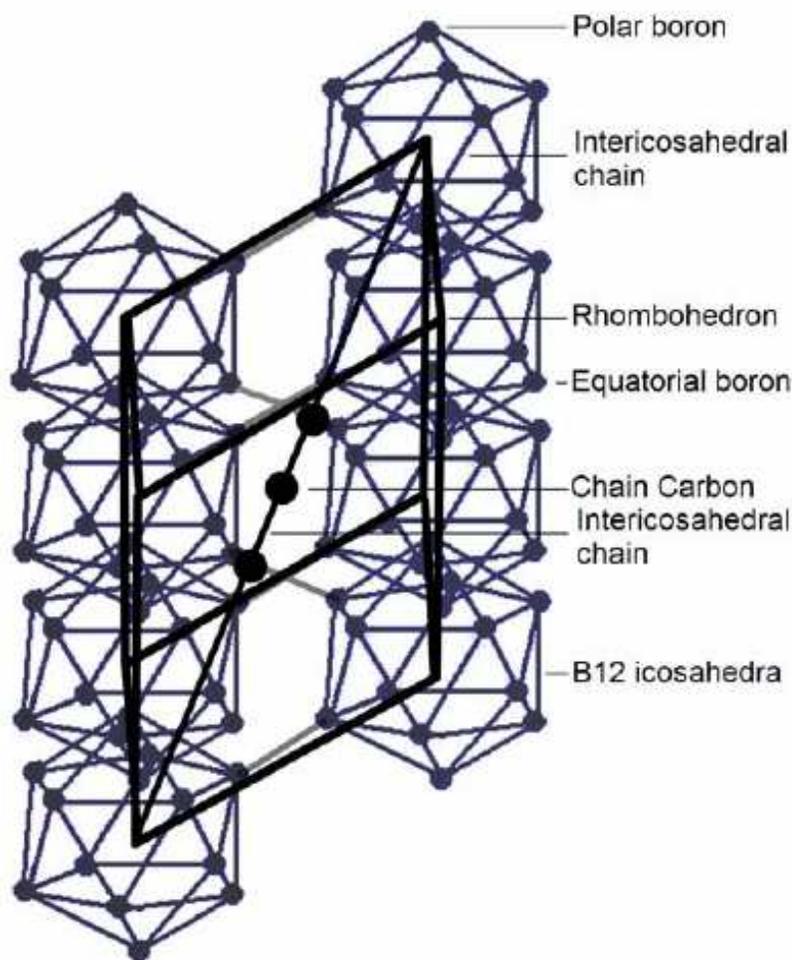


Fig. 1. Rhombohedral crystal structure of a stoichiometric boron carbide , B₄C

The boron rich icosahedral structures reside at the vertices of a rhombohedron. Four sites are available for a total of 15 'B and C' atoms in the crystal structure. Thus the B and C atoms can interchange each other within both icosahedral and inter icosahedral atom chains. This constitutes the basis for the large homogeneity range for B-C system. Mostly the C-atoms are

along the diagonal of the rhombohedron forming the inter-icosahedral chains in chain configuration. The widely studied three carbon atoms as in $B_{12}C_3$ along the intericosahedral chain can be substituted all or in parts by vacancies or several types of atoms, thus forming the boron rich and carbon rich end of the stoichiometry [22]. The case of tetragonal structural unit for boron carbides is little different. It consists of four boron icosahedra in a tetragonal arrangement with two carbon atom and the two remaining boron atoms occupying inter-icosahedral sites.

2. Experimental

A thermal explosion process mode was used to form boron carbide using stoichiometric amount of high pure boron oxide with excess carbon and magnesium considering B_4C as the expected product [16]. The input material were mixed in a high energy planetary ball mill, cold compacted and heat treated in presence of positive non-oxidizing gas pressure. The chemically purified products were collected and characterized in as formed as well as in crushed condition. The characterization techniques like X-ray diffraction (Cu anode), X-ray fluorescence techniques along with scanning electron microscopy were used to establish the products. The XRD pattern for the multi phases was compared with a high pure boron carbide standard with B/C ratio as 4:1. The diffraction patterns were recorded using Bragg - Brentano geometry at the room temperature over the angular range 8° to 90° in the step of 0.002° with 2 second as step time. A number of secondary experiments, graphical simulation and calculations were performed to confirm the presence of poly crystalline phases. These were carried out especially where the percentages of phases were less and when the pattern breaking was critical because of peak over laps, experimental, compositional, structural, machinery and algorithms limitations [23]. Lattice constants were calculated from the peak positions in the XRD profiles using peaks existing up to 50° of two theta values.

3. Result and discussion

The effect of an limited penetrable boundary material such as carbon, on the reaction dynamics of a double species (B-C) reaction process in the low exothermic B-O-Mg-C system with diffusive reactants were studied [24]. The SEM micrographs showing the as formed and as chemically purified powders have been presented in the Fig.2. The purity was more than 99.97% with boron to carbon ratio determined as 5.042 [16]. The observed x-ray diffraction pattern has been shown in Figure-3 with a comparative account with a standard obtained by thermally and chemically purifying a commercially available B_4C . As observed in the Fig.3 a number of new peaks were found compared to the B_4C standard.

Phase detection and quantification: In case of XRD analysis of rhombohedral boron carbides the strong (021) Debye ring is ideally located in between 36.764° to 38.116° and the ring (104) from 34.0760° to 35.284° of 2-theta values. The relative peak positions of boron carbides depend on the carbon contents [25]. Beside this fact, the data analysis of a multi phase pattern is always a complicated process. It happens specially in presence of factors like peak overlaps, continuous back grounds, variation in fixing of peak parameters, fixed algorithm, absorption, porosity, axial divergence of beam, radiation types, alignment problems, disorders, crystallinity, sample preparation, grain size and its uniformity [22, 25]. Such a problem becomes critical when the structural variation are narrow as found in r-boron carbides. Considering the above discussed factors, during subsequent approaches and

by search-match comparison with JCPDS-ICDD data the determined phases were as $B_{13}C_2$, $B_{48}B_2C_2$, $B_{41.11}C_{4.45}$, $B_{12}C_3$, B_8C , $B(T)$ and $B(Rh)$ with JCPDS cards numbers as 33-0225, 71-1107, 71-0363, 35-0798, 26-0232, 31-0206 and 78-1571 respectively. The phases were quantified in a semi quantitative manner using the combination of normalized RIR method [26], integrated intensities and mass absorption ratios. Through the semi-quantitative analysis give theoretically exact result, there are several source of as- discussed errors which prohibit an accurate result. However in any case the result is at least approximately correct [Philips electronic publication, Nov 2001]. Stepwise standard workouts were carried out to confirm the accuracy of the quantification [27].

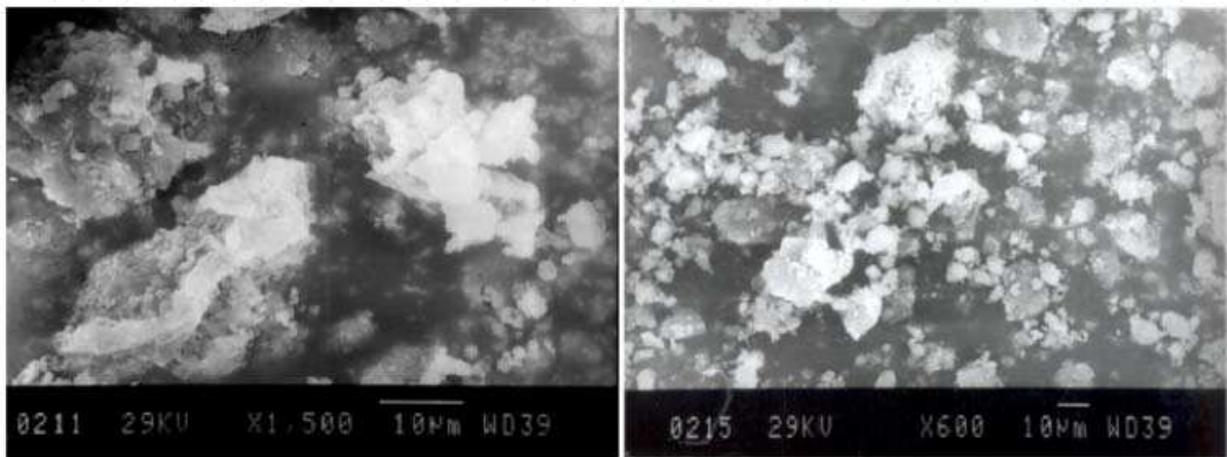


Fig. 2. SEM micrograph of (a) as formed product (b) the chemically purified product

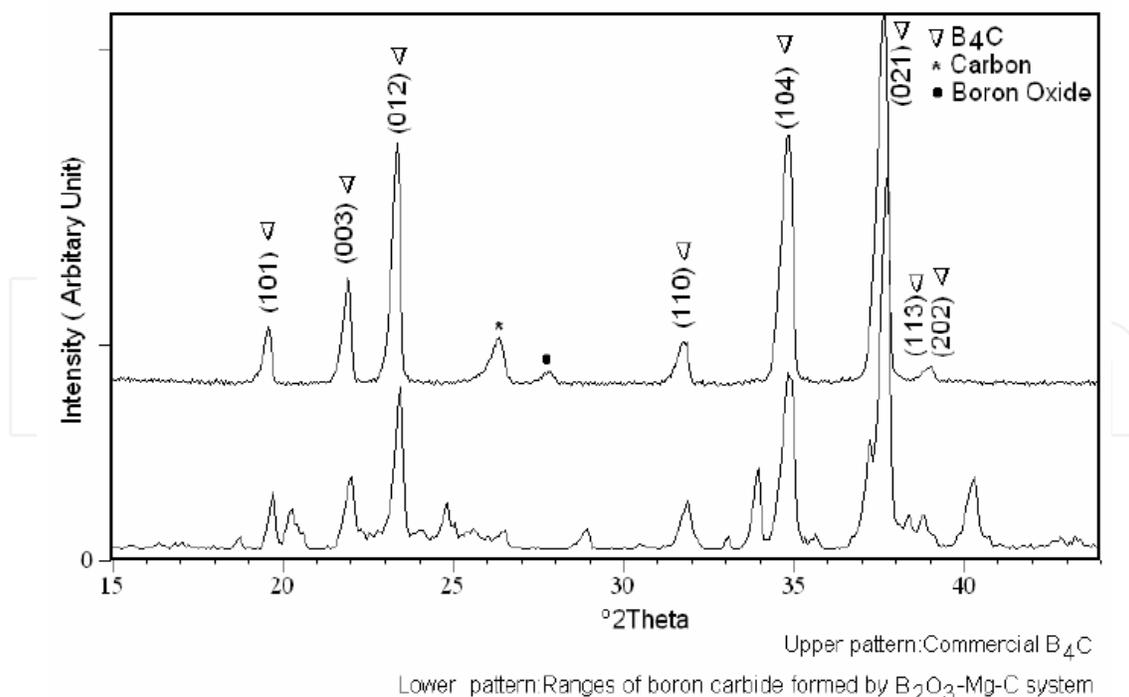


Fig. 3. XRD peak comparison of a carbothermal and the micropyretic BC product

In this study, the peak shapes were determined using individual intensity convolution and peak shape functions (mainly a set of Pseudo-Voigt functions). The intensity of boron rich

peaks were found to be weak in compared to carbon rich boron carbides which have also been confirmed by a work of Disbudak and Silvestre et al. [28]. A variation in peak positions as observed for various rhombohedral boron carbides are explained to be the substitution of a carbon atom by a boron atom in the inter or intra icosahedrons [19,21,22,28]. The phase detection has also been confirmed by a multiphase simulation programme and the results have been represented by Fig. 4.

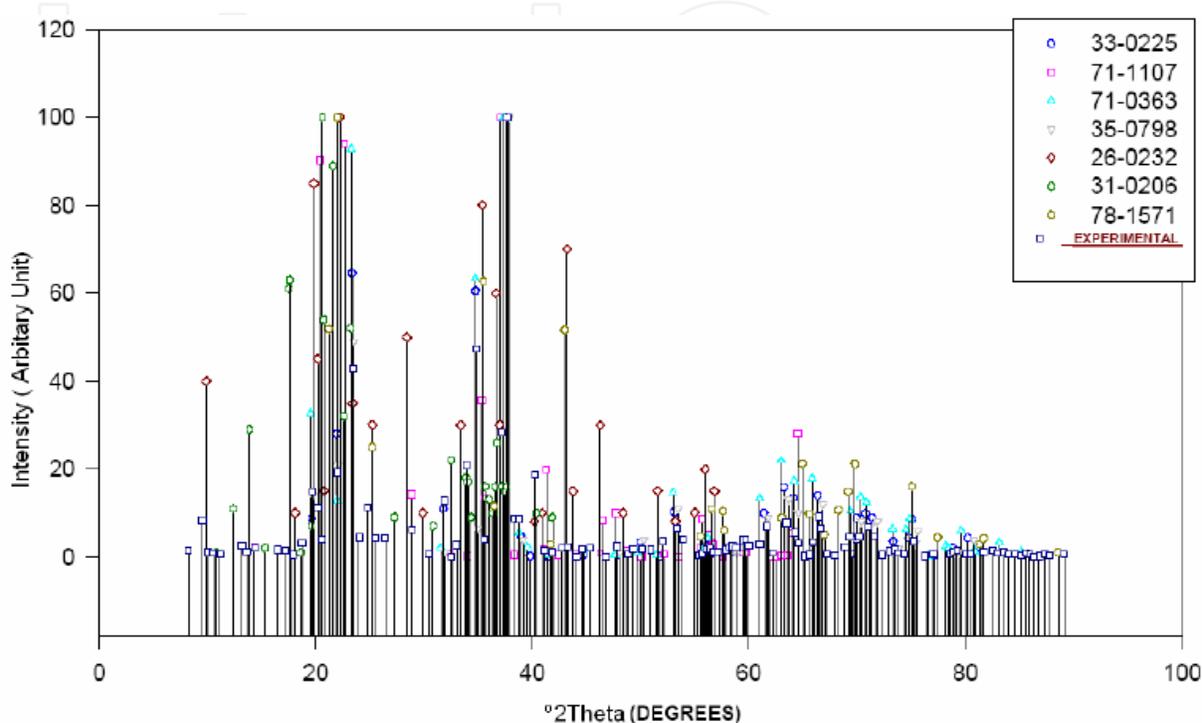


Fig. 4. Overlapped graph of experimental and simulated stick patterns

3.1 Structure estimation

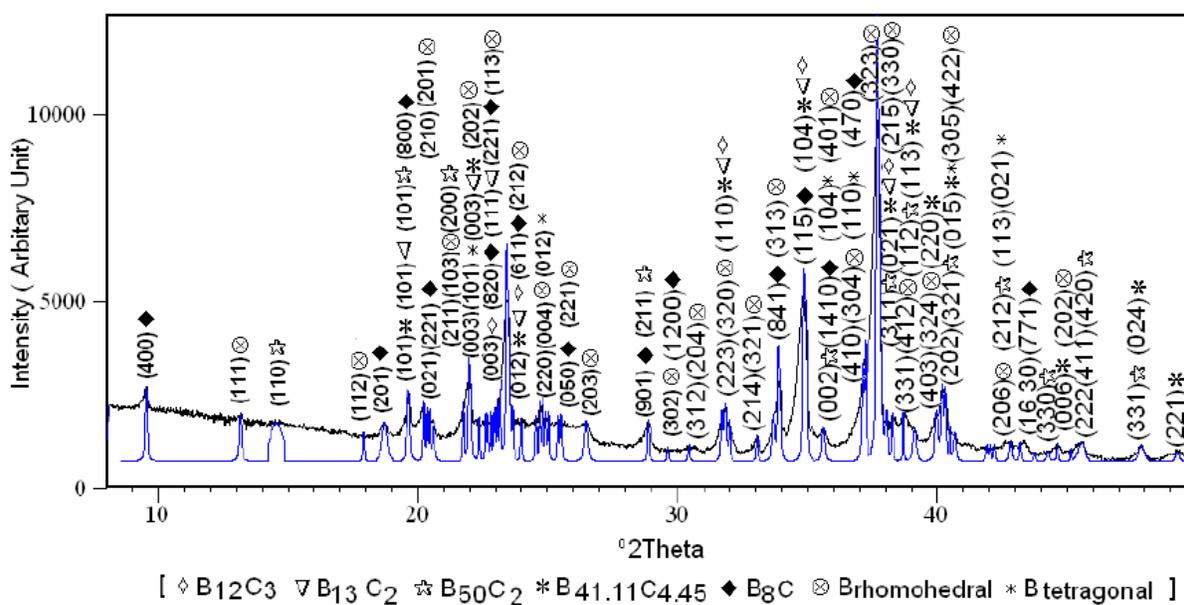


Fig. 5. Selected range of the multiphase, ab-initio indexed pattern

The classical ways to obtain accurate unit cell parameters are full pattern decomposition and full profile based Rietveld refinement [29]. Since some of the parameter required for a Rietveld refinement are not known for the nonstoichiometric phases, the experimental pattern [8° to 90° range, 0.002° step size] was structurally analyzed by defining the ab initio indexing to the determined 2-theta peak positions within 8° to 50° range (Fig.5). In total 42 peaks were used to find best lattice parameter sets for the multi phase pattern. The results were summarized as in the table-1 and table-2 which predict the existence of both tetragonal and rhombohedral borons along with rhombohedral, tetragonal and orthorhombic phases for boron carbides.

Phases	Major2θ (ideal)	Major2θ(Observed)	(hkl)	Obs. Intensity	Calculated BxCy	Mass absorption coefficient	Space Group	SemiQ (%)
B ₁₂ C ₃	37.819°	37.747°	(021)	130.22	B ₄ C	3.589	R-3M	52.59
B ₁₃ C ₂	37.604°	37.677°	(021)	114.29	B _{6.5} C	3.416	R-3M	42.41
B ₄₈ B ₁ C ₂	37.059°	37.136°	(311)	32.12	B ₂₅ C	3.165	P4 ₂ /nmm	02.30
B _{4.11} C _{4.15}	37.324°	37.324°	(021)	29.23	B _{9.24} C	3.320	R-3M	01.40
B ₈ C	22.206°	22.013°	(820)	58.49	B ₈ C	3.354	P222 ₁	00.80
B _(T)	22.579°	22.590°	(221)	11.90	-	3.06	P4 ₁	rest
B _(Rh)	21.982°	22.020°	(101)	29.54	-	3.06	R-3M	rest

Table 1. Observed experimental values with standard available values for the multiphase pattern

3.2 Explanation to observed result

Most of the reported solid state processes [10,12,30] produce single phase boron carbide. The discussed experimentation process is initially a solid state process. With increase of temperature the (B₂O₃-C-Mg) system with high carbon acts like a 'colloidal system'. Local melting in the vicinity of the reacting species has an appreciable effect on the equilibrium interfacial concentrations and the diffusion co-efficient in contacting phases. The decrease in the exothermicity because of presence of low melting glassy B₂O₃, is adequately compensated by the external heating. The excess external heat sustained the reactive process to attain the adiabatic temperature and the colloidal condition assisted in suppression of initial reduction products like B₂O₃ and thus the other sub oxides or gaseous species [31]. These along with the prevailing reaction conditions in the "high pressure changed environment" prevented the formation of 'MgC₂ or MgB₂'. As a result, the expected reactive interactions were achieved. The products were mostly B-C and Mg-O based compounds. Presence of carbon as buffer reduced the formation of MgO₂ , Mg₂B₂O₆ with a self adjustment to CO activity in the closed furnace environment.

Phases	$a_{exp}(a_{estab}) \text{ \AA}^{\circ}$	$b_{exp}(b_{estab}) \text{ \AA}^{\circ}$	$c_{exp}(c_{estab}) \text{ \AA}^{\circ}$	$\alpha (^{\circ})$	$\beta (^{\circ})$	$\gamma (^{\circ})$	Synd.FOM	Sp.gr. sym.	$V_{exp}(V_{estab}) (\text{ \AA}^3)$	Data type
$B_{12}C_3$	5.616(2) [5.6003(5)]	5.616(2) [5.6003(5)]	12.129(8) [12.086(2)]	90 90 120 90 90 120			28.905 [58]	R-3M(166)H R-3M(166)H	331.291 [328.270]	Experimental 1984,Exper.
$B_{13}C_2$	5.877(4) [5.633]	5.877(4) [5.633]	12.17(1) [12.164]	90 90 120 90 90 120			12.1365 [103]	R-3M(166)H R-3M(166)H	364.017 334.260	Experimental 1980, Calcu.
$B_{48}B_2C_2$	8.770(7) [8.722]	7.241(7) [8.722]	6.176(6) [5.080]	90 90 90 90 90 90			13 5.701	$P4_2/nmm(134)$ $P4_2/nmm(134)$	392.250 386.450	Experimental 1997,Calcu..
$B_{41.11}C_{4.45}$	5.211(4) [5.672]	4.847 [5.672]	8.962(9) [12.1428]	90 90 112.06 90 90 120			23.000 -	R-3M(166)H R-3M(166)H	211.000 338.320	Experimental 1997,Calcu..
B_8C	20.809 [35.909]	4.503(3) [17.653]	4.109 [5.094]	90 90 90 90 90 90			14 2	$222_1(17)$ -	386.66 3229.09	Experimental Single crystal
$B_{(I)}$	9.443(9) [10.140]	9.443(9) [10.140]	8.322(7) [14.170]	90 90 90 90 90 90			5.744 73	$P4_1(76)$ $P4_1(76)$	749.954 1456.95	Experimental Calculated.
$B_{(Rh)}$	3.2776 [4.927(3)]	3.2776 [4.927(3)]	36.38 [12.564(7)]	90 90 120 - - -			9.9937 7.745428	R-3 M(166)H R-3 M(166)H	336.8905 264.13	Experimental Calculated.

Notation : *exp.(experimental finding) *estb.(established value from literatures) * FOM (Figure of merit) *Calcu.(Calculated)

Table 2. Crystal structure values as deduced from the indexed peaks in the multi-phase pattern

Considering the experimental results and the above analysis, a non-equilibrium mechanism is proposed. The global behavior of such a process depends on limiting factors like, the number of particles that interact locally in a single step and the dynamics of the reacting phases. Considering the report of combustion process achievements (maximum adiabatic temperature, combustion propagation rate, rate of heating, and thermal gradient) and the behavior of inputs, the phase formation in the process is proposed to have occurred in a non isothermal condition with existence of normal diffusive to ballistic diffusive behaviors. The diffusion length, inter effect of particles growth were found to be less and the MgO growth is not determined to be higher [figure-2]. Considering the above discussions, ignoring the Langmuir adsorption and basing on Fick's second law, a self consistent equation is defined for the process as

$$\frac{\partial \bar{\rho}_r}{\partial t} = D_T \nabla^2 \bar{\rho}_r - 2k \cdot \bar{\rho}_r^2$$

where k represents the slower kinetic term, D_T as diffusion constant at corresponding temperature, $\bar{\rho}_r$ is the average density of reacting phases and $\frac{\partial \bar{\rho}_r}{\partial t}$ is the rate of decay of reacting phases. Under present condition the product phase concentration is dependent more on the mean free path rather on 'k' or ' D_T '. Since the spatial dimension of the process path is low i.e. two dimensional, at any time, reacting phases those are close to each other react at a faster rate compared to widely separated one. This explains the formation of ' mix phases' with the importance of process input homogeneity. The limit to achieve theoretical yield value is much dependent on 'before process interaction homogeneity'. Thus the

possibilities associated with a simple type of propagating 'non equilibrium heat wave' is generalized as in the Fig.6.

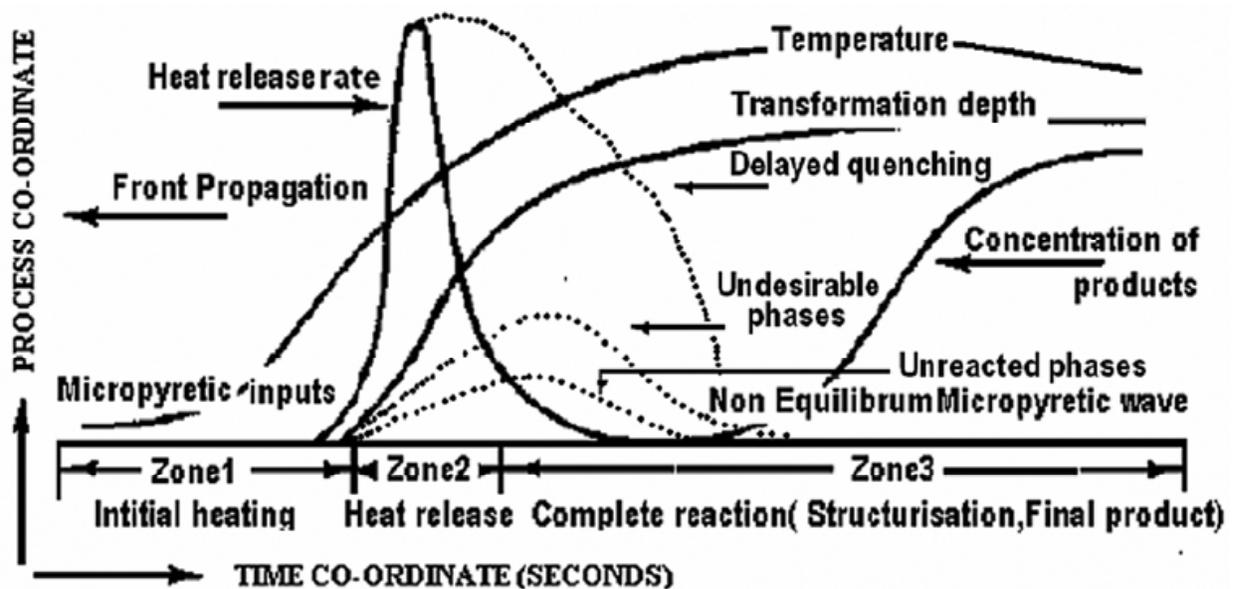


Fig. 6. A typical presentation of associated non equilibrium micropyretic wave

The multiphase boron carbides formation can be explained in the following ways. Above 1400°C with relatively longer reaction time, the formation of β -rhombohedral boron or the tetragonal elementary boron is possible with a three dimensional net work of B12 clusters [32]. Among them β -rhombohedral boron is more stable but its survival as an elementary phase depends on the existing high temperature thermodynamics. The formation of both type borons as found in this study also supported by an earlier published work [32]. Concerning the formation of tetragonal boron, as per Gmelin et al carbon helps in stabilizing the tetragonal boron [33]. In this current high pressure experiment carbon was one of the excessive input. The diffusion of boron and carbon to one another crystal lattices are also possible. However the diffusion length is limited by the involved high temperature kinetics. Again reports of C.E. Lowell et al reveal the maximum solubility of B in C(gr) as 2.35 at% at 2350°C if interstitial formation is neglected [34]. Boron has a diffusivity of $2.028 \times 10^{-11} \text{ m}^2/\text{sec}$ for others lattices at 1000°C. The phase formation results along with the predictions of mentioned literatures favour the formation of B-C lattice on carbon particle with diffusion of C into the developing B-lattice. Thus in boron rich boron carbides the primary bonding is B-B. The arrangement of boron structures depending on its relative activity in a compound can be summarized as in Fig.7.

Earlier the formation of tetragonal boron carbide with less than 10% C-content had been reported by CLJ Adkins and other [35]. As per them this occurs mostly under carbon depletion conditions compared to that exist for the rhombohedral form. The presence of the tetragonal boron, short processing time [Fig. 6], the time based consumption of carbon along with the high diffusion coefficient of boron indicate tetragonal boron carbide is a modification of tetragonal boron. And the conversion process is a consecutive process. The relative percentage of its formation is affected by the prevailing heating conditions in the furnace and the reactants homogeneity. According to Aselage et al. the thermodynamic instability of pure β -rhombohedral boron allows the formation low C-

rhombohedral boron carbide. This supports the experimental result of getting a range of rhombohedral boron carbides and also may be an out come of increased chemical potential of B or reduced chemical potential of C. The presences of larger unit cell of boron carbides suggest the reduced carbon concentration. A comparative study of β - rhombohedral boron and boron carbide spectra with different C-concentration has been reported by Tallant et al. [37]. The main structural difference in-between them is the presence of a CBC linear chain in the boron carbides [26,38,39]. Like boron, rhombohedral boron carbide also has a crystalline icosahedral appearance with B12 cluster as a part of its structure. Thus the above evidences and all results favour the formation of B-C system from elementary boron.

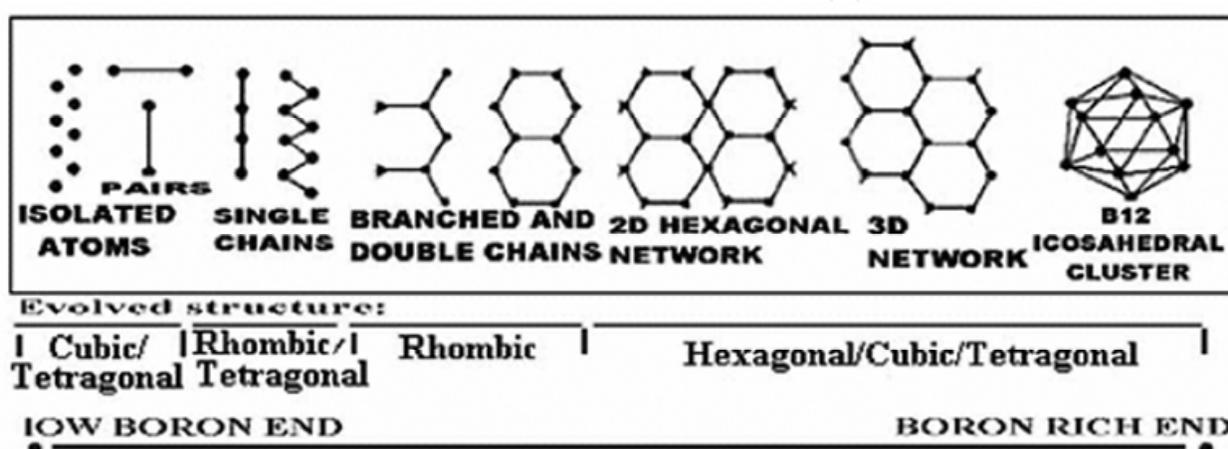


Fig. 7. Relative prediction of structural development in boron based system

4. Conclusion

Experimental data obtained during the investigation and subsequent analysis predict the existence of a range of boron carbides and borons like $B_{13}C_2$, $B_{48}B_2C_2$, $B_{41.1}C_{4.45}$, $B_{12}C_3$, B_8C , B(T) and B (Rh) in the single step solid state process. The other findings of the study are summarized as

1. Both 'rhombohedral and tetragonal' boron are formed as in-situ reactive products.
2. The nature and the stability of this insitu boron decide the type of boron carbide thus formed. In other word, varying multiple heat waves as obtained in the micropyretic process and the prevailing reaction conditions like steady heat wave formation, vortex detachment or quenching (heat flow stability), coupling between the pressure wave and combustion assist the formation of the boron carbides with varying crystallographic structures.
3. Absence of carbon rich boron carbide shows the diffusion of boron to carbon lattice is limited and even show at high temperatures.
4. Tetragonal boron carbide forms under carbon depletion condition with availability of tetragonal borons.
5. Thermodynamic and kinetic instability of pure rhombohedral borons leads to the formation of low carbon boron carbides. The existence of a range of rhombohedral boron carbide is a out come of either the increased chemical potential of r- boron or the decreased chemical potential of the carbon in the local chemical environment.

6. The foamy and less agglomerated crude products obtained from the reactive process indicate the existence of high reaction rate and low partial surface energy. And the reaction mechanics of the process enables the obtaining of both equilibrium and non equilibrium boron based powders of high purity with a low micron size good particle dispersity.
7. Type and quantity of insitu boron controls the percentage and structure formation of boron carbides.

5. References

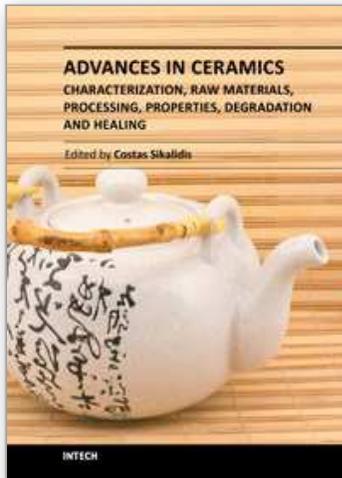
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Processing, Properties, Degradation and Healing**

Edited by Prof. Costas Sikalidis

ISBN 978-953-307-504-4

Hard cover, 370 pages

Publisher InTech

Published online 01, August, 2011

Published in print edition August, 2011

The current book consists of eighteen chapters divided into three sections. Section I includes nine topics in characterization techniques and evaluation of advanced ceramics dealing with newly developed photothermal, ultrasonic and ion sputtering techniques, the neutron irradiation and the properties of ceramics, the existence of a polytypic multi-structured boron carbide, the oxygen isotope exchange between gases and nanoscale oxides and the evaluation of perovskite structures ceramics for sensors and ultrasonic applications. Section II includes six topics in raw materials, processes and mechanical and other properties of conventional and advanced ceramic materials, dealing with the evaluation of local raw materials and various types and forms of wastes for ceramics production, the effect of production parameters on ceramic properties, the evaluation of dental ceramics through application parameters and the reinforcement of ceramics by fibers. Section III, includes three topics in degradation, aging and healing of ceramic materials, dealing with the effect of granite waste addition on artificial and natural degradation bricks, the effect of aging, micro-voids, and self-healing on mechanical properties of glass ceramics and the crack-healing ability of structural ceramics.

How to reference

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R.M. Mohanty, K. Balasubramanian and S.K. Seshadri (2011). Evolution of Crystallographic Structures and Phases in Micropyretically Formed Boron Rich Boron Carbide – a New Material System, *Advances in Ceramics - Characterization, Raw Materials, Processing, Properties, Degradation and Healing*, Prof. Costas Sikalidis (Ed.), ISBN: 978-953-307-504-4, InTech, Available from: <http://www.intechopen.com/books/advances-in-ceramics-characterization-raw-materials-processing-properties-degradation-and-healing/evolution-of-crystallographic-structures-and-phases-in-micropyretically-formed-boron-rich-boron-carb>

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