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# pH-Sensitive Nitroxide Radicals for Studying Inorganic and Organo-Inorganic Materials and Systems

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Additional information is available at the end of the chapter

<http://dx.doi.org/10.5772/39119>

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## 1. Introduction

Many inorganic materials are widely used as adsorbents and catalysts. For example, silica gels efficiently absorb vapors and gases in chemical reactors and are applied as filters for the purification of mineral oils and water (Buyanov, 1998). Different aluminum oxide modifications show good adsorption and catalytic properties in many organic reactions due to the presence of active sites on their surface (Lisichkin et al., 2003). TiO<sub>2</sub> gels are widely used in heterogeneous catalysis owing to their enhanced chemical stability, the accessibility of active sites on their surface throughout the reaction volume, simplicity of reaction product separation, and the feasibility of repeated regeneration (Petrov et al., 1998).

The efficiency of these materials mainly depends on the degree of surface development, texture and structural characteristics, availability of active centers, and possibly medium acidity near these centers. The specific surface diminishes on drying in any procedure for the preparation of xerogels. It can be maintained constant by a number of methods, including the use of additions (Ur'ev & Potanin, 1992), for example, powder cellulose (PC) (Shishmakov et al., 2007). Cellulose is a linear high molecular polysaccharide, which forms rigid chain structures due to the inter-molecular hydrogen bonding. It functions in nature as an agent that imparts high mechanical stability to plant tissues (Nikitin, 1962). The deposition of SiO<sub>2</sub> and TiO<sub>2</sub> xerogels on the PC surface affords composite materials (CMs) with a high dispersity of particles (Shishmakov et al., 2010).

The use of hybrid organo-inorganic materials as supports is a new area in the development of new metal-containing catalytic materials. It allows to obtain supports with specific surfaces that are capable of retaining the metallic component of a catalytic system more strongly.

Chitosan, poly-D- $\beta$ -glucosamine, is a commercially available amino polymer that is a perfect complexing agent, due to the strong donor properties of both the amino and hydroxyl groups (Varma et al., 2004). Chitosan is thus widely used in obtaining various catalytic materials, including those containing  $\text{Au}^0$  that are used in the hydroamination of alkenes (Corma et al., 2007);  $\text{Pd}^0$  used for the reduction of ketones (Yin et al., 1999); the  $\text{Pd}^0$ - $\text{Ni}^0$  bimetallic system, used for carbonylation (Zhang & Xia, 2003); Os (VIII), used for hydroxylation (Huang et al., 2003);  $\text{Co}^{2+}$ , used for hydration (Xue et al., 2004); and  $\text{Cu}^{2+}$ , used for the oxidation of catecholamines (Paradossi et al., 1998).

$\text{SiO}_2$  is usually used as the inorganic component for these systems. The obtained hybrid materials are used to create sorbents of 3d-metal ions (Liu et al., 2002); to immobilize enzymes (Airoldi & Monteiro, 2000); as a solid phase for the liquid chromatography of organic compounds (Budanova et al., 2001), including enantiomers (Senso et al., 1999); and to improve the mechanical properties of other polymers (Yeh et al., 2007). Other oxides in combination with chitosan allow us to obtain biosensors based on  $\text{ZnO}$  substrate (Khan et al., 2008), selective sorbents of fluoride ions based on  $\text{Al}_2\text{O}_3$  substrate (Viswanathana & Meenakshib, 2010) and magnetic materials based on  $\text{Fe}_3\text{O}_4$  substrate (Li et al., 2008). Using an organic polymer (e.g., cellulose) as a substrate also has advantages in the sorption of metal ions (Corma et al., 2007). Metal-containing hybrid organo-inorganic materials can also be used as antibacterial composites (Mei et al., 2009), as sorbents of proteins (Shi et al., 2003), and as pervaporation membranes (Varghese et al., 2010).

Nanostructured metal oxides, which are distinguished by extremely developed surface and porosity of particles, are new promising materials for different fields of science and technology, especially, for heterogeneous catalysis and chemistry of adsorption phenomena (Zakharova et al., 2005).

Many sorption and catalytic processes are pH-dependent. Therefore, the determination of acidity and other acid-base characteristics in pores of inorganic, organo-inorganic materials is of great practical interest, since the catalytic and adsorption properties of solid-phase objects are affected by not only the chemical nature of solutions, but also specific conditions inside pores and on the surface of these materials. The mobility of liquid molecules in pores of inorganic sorbents was investigated by some authors using the spin probe method (Borbat et al., 1990; Martini et al., 1985). Recently, a new method was developed for the determination of medium acidity in pores of solids ( $\text{pH}_{\text{int}}$ ) by means of pH-sensitive nitroxide radicals (NRs) as spin probes (Molochnikov et al., 1996; Zamaraev et al., 1995). In recent years, this method was used to measure  $\text{pH}_{\text{int}}$  in micropores of various cross-linked organic polyelectrolytes (ion-exchange resins and films) (Molochnikov et al., 1996, 2004) and in pores of some zeolites and kaolin (Zamaraev et al., 1995). We found that  $\text{pH}_{\text{int}}$  inside sorbents differ from the pH of external solutions by 0.8–2.1 units (Molochnikov et al., 1996). The method developed allowed us to study the processes of sorption and hydrolysis in ion-exchange resins and the catalytic properties of  $\text{Cu}^{2+}$ -containing carboxyl cation exchangers (Kovaleva et al., 2000), to determine ionization constants of functional groups and to give a critical estimation to the regularities previously found for the behavior of adsorbents in aqueous media.

pH-sensitive nitroxide radicals (NR) as labels were also used to determine surface electrical potential (SEP) of different biological objects like phospholipids (SLP) - derivatives of 1,2-dipalmitoyl-sn-glycero-3-phosphothioethanol (PTE) (Voinov et al., 2009) and the mixed bilayers composed of dimyristoylphosphatidylglycerol and dimyristoylphosphatidylcholine (Khramtsov & Weiner, 1988).

This work is aimed to review the applications of pH-sensitive NR as probes and labels for determination of local acidic and electrochemical characteristics of inorganic and organo-inorganic materials and systems such as :

- $\text{Al}_2\text{O}_3$  (non-modified and modified with  $\text{F}^-$  and  $\text{SO}_4^{2-}$  ions) and  $\text{TiO}_2$  as hydrogels and nanopowders doped with  $\text{Cu(II)}$  (Molochnikov et al., 2007) ;
- pure and  $\text{Cu}^{2+}$ -containing solid-phase composites based on nanostructured  $\text{SiO}_2$  and  $\text{TiO}_2$  and powder cellulose (Shishmakov et al., 2010; Parshina et al., 2011) ;
- pure and  $\text{Co}^{2+}$ -containing hybrid organo-inorganic materials based on the chitosan- $\text{SiO}_2$ , chitosan- $\text{Al}_2\text{O}_3$ , and -chitosan-cellulose systems (Mekhaev et al., 2011a, 2011b).

## 2. Experimental

### 2.1. Objects of study

$\alpha\text{-Al}_2\text{O}_3$  (basic aluminum oxide),  $\gamma\text{-Al}_2\text{O}_3$  and its acid-modified ( $\text{HF}$  and  $\text{H}_2\text{SO}_4$ ) derivatives were supplied by A. M. Volodin. The IK-02-200 type  $\gamma\text{-Al}_2\text{O}_3$  was synthesized by the calcination of aluminum hydroxide at  $600^\circ\text{C}$ .  $\gamma\text{-Al}_2\text{O}_3$  was modified through the sample impregnation with acids followed by calcination at  $600^\circ\text{C}$  that resulted in changes in the acidic properties of its surface (the phase composition and specific surface area of the samples remained unchanged).  $\alpha\text{-Al}_2\text{O}_3$  was prepared by the long-term heating of  $\gamma\text{-Al}_2\text{O}_3$  to  $1300^\circ\text{C}$ .  $\gamma\text{-Al}_2\text{O}_3$  and  $\alpha\text{-Al}_2\text{O}_3$  had specific surface areas of 220 and  $145\text{ m}^2/\text{g}$ , and an average pore diameter of 6 nm, respectively. The structural characteristics of the matrices were determined from the isotherms of nitrogen adsorption at 77 K measured on a Micromeritics ASAP 1400 volumetric setup and by mercury porosimetry with a Micromeritics Pore Size 9300 setup in the Institute of Catalysis, Siberian Division, Russian Academy of Sciences.

The technique of synthesis of  **$\text{TiO}_2$  hydrogel** through the hydrolysis of a tetrabutyl titanate solution in methanol with water at room temperature and under intensive agitation is given in (Shishmakov et al., 2003). The precipitate was washed out with water until no butanol in washing water was observed and heated to  $100^\circ\text{C}$ . The resulting product was  **$\text{TiO}_2$  xerogel**.

**$\text{SiO}_2$  hydrogel** was synthesised by dissolving 10 ml of  $\text{Na}_2\text{SiO}_3$  (TU 6-15-433-92) in 30 ml of  $\text{H}_2\text{O}$ . Then the hydrolysis of  $\text{Na}_2\text{SiO}_3$  solution in 30 ml of 10%  $\text{HCl}$  solution was pursued under intensive agitation. During the reaction of condensation a gelatinous  $\text{SiO}_2$  gel and  $\text{NaCl}$  are formed. The precipitate was filtered off until no chlorine ions in washing water was observed, and dried at  $100^\circ\text{C}$  during 24h until it attained a constant weight. The resulting product was  **$\text{SiO}_2$  xerogel**.

**Powder cellulose (PC)** was obtained by hydrolysis of cellulose sulfate (Baikal Cellulose plant, TU OP 13-02794 88-08-91) in 2.5 N hydrochloric acid at 100°C. The hydrolysis was carried out for 2 h. The resulting product was washed on a filter with distilled water to the neutral pH of the washing water and dried at 100°C.

Composite materials (CMs) based on nanostructured TiO<sub>2</sub> and PC called as TiO<sub>2</sub>-PC xerogels of 70, 53 and 43 % wt. TiO<sub>2</sub> were prepared by diluting 3 g of tetrabutoxytitanium and 0.5 ; 1 and 1.5 g of PC, respectively, in 3 mL of methanol. The hydrolysis was pursued in 10 mL of water at 20°C under intensive agitation, resulting in the condensation of TiO<sub>2</sub> (PC didn't participate in condensation). The TiO<sub>2</sub> particles formed were deposited on a surface of PC.

**Composite materials (CMs) based on nanostructured SiO<sub>2</sub> and PC called as SiO<sub>2</sub>-PC xerogels of 68, 52 and 35% wt.** were prepared from the solutions of 10 ; 5 ; 5 ml of Na<sub>2</sub>SiO<sub>3</sub> and 30 ; 15 ; 15 ml of H<sub>2</sub>O which were modified by introducing 2 ; 2 ; 4 g of PC, respectively. The hydrolysis of the first solution were performed in 30 mL of 10 %HCL, and that of other ones was done in 15 mL of 10% HCL.

The precipitates of the CMs prepared were washed out with hot water, filtered and dried at 100 °C during 24h until they attained a constant weight.

The specific surface ( $S_{sp}$ ) of the synthesized samples was measured using a SORBIMS instrument (ZAO Meta, Novosibirsk) and calculated by the BET procedure. The data are given in Table 1 (Parshina et al., 2011).

	TiO <sub>2</sub>				SiO <sub>2</sub>		
PC, %	0	30	47	57	0	32	65
$S_{sp}, m^2/g$	66	177	226.4	261.9	29.5	145	239

**Table 1.** The specific surface of the pure xerogels of TiO<sub>2</sub> and SiO<sub>2</sub> and the composites with different percentage of PC

Since the specific surface of powder cellulose did not exceed 1 m<sup>2</sup>/g , the growth of  $S_{sp}$  was caused by the fragmentation of TiO<sub>2</sub> and SiO<sub>2</sub> particles deposited on the PC surface during the synthesis of CMs. According to the absolute values of  $S_{sp}$ , the procedure used for the preparation of CMs afforded dioxides with a high degree of dispersity.

Powdered samples **nanostructured TiO<sub>2</sub>** were prepared through heating a sol for 1 h at 200°C followed by washing with distilled water to remove residual acid used for the sol stabilization and drying at room temperature. The specific surface area of the samples was 240 m<sup>2</sup>/g, and the average particle diameter ranged from 4 to 5 nm (Poznyak et al., 1999).

Microcrystalline cellulose (MCC) with an ash content of 0.16% and a humidity of 1.1% produced by JSC Polyex; Basic aluminum oxide; BS-50 silica and chitosan produced by JSC Sonat (Moscow) were used to obtain chitosan-containing ahybride organo-inorganic systems. The degree of deacetylation of chitosan (DD) determined by 1H NMR spectroscopy, its molecular weight as determined by viscosimetry and the ash content were found to be 0.84, 250 kDa and 0.19%, respectively (Mekhaev et al., 2011a). The BS-50 type silica had a specific surface area of 45 m<sup>2</sup>/g and an average diameter of pores of 15 nm (Mekhaev et al., 2011a).

**The hybrid chitosan–(SiO<sub>2</sub>, Al<sub>2</sub>O<sub>3</sub>, cellulose) systems** were obtained by depositing chitosan on the support surface.

0.3 g (1.8 mmol) of chitosan was dissolved in 14.5 ml of water containing 0.22 ml (3.84 mmol) of acetic acid with constant stirring. The substrate in quantities of 3 g was then added, and the solution was stirred for 30 min more.

1 M NaOH solution was added to the suspension under stirring until the pH value reached 13. The precipitate was filtered, washed until the pH value was 7, and dried at 60°C until it attained a constant weight.

CHN analysis was performed using an automatic analyzer PerkinElmer, Inc. The data are given in Table 1. IR spectra of diffuse reflection were recorded using the PerkinElmer Spectrum One spectrometer.

System	C	H	N	Cl	Co	Formula
MCC-Chitosan	43.08 (43.11)	6.42 (6.35)	0.37 (0.39)	-	-	10C <sub>6</sub> H <sub>10</sub> O <sub>5</sub> · 0.5C <sub>6</sub> H <sub>11</sub> NO <sub>4</sub> · 3H <sub>2</sub> O
Al <sub>2</sub> O <sub>3</sub> -Chitosan	4.07 (4.01)	1.04 (0.61)	0.46 (0.78)	-	-	16Al <sub>2</sub> O <sub>3</sub> · C <sub>6</sub> H <sub>11</sub> NO <sub>4</sub>
SiO <sub>2</sub> -Chitosan	4.50 (4.49)	0.94 (0.68)	0.51 (0.87)	-	-	24SiO <sub>2</sub> · C <sub>6</sub> H <sub>11</sub> NO <sub>4</sub>
MCC-Chitosan- Co <sup>2+</sup> (I)	43.12 (43.12)	6.15 (6.04)	0.59 (0.40)	1.30 (1.28)	1.38 (1.33)	10C <sub>6</sub> H <sub>10</sub> O <sub>5</sub> · 0.5C <sub>6</sub> H <sub>11</sub> NO <sub>4</sub> · Co <sub>0.4</sub> (OH) <sub>0.17</sub> Cl <sub>0.65</sub>
Al <sub>2</sub> O <sub>3</sub> -Chitosan- Co <sup>2+</sup> (II)	3.64 (3.82)	0.79 (0.58)	0.47 (0.74)	2.69 (2.63)	2.18 (2.19)	16Al <sub>2</sub> O <sub>3</sub> · C <sub>6</sub> H <sub>11</sub> NO <sub>4</sub> · 0.7CoCl <sub>2</sub>
SiO <sub>2</sub> -Chitosan- Co <sup>2+</sup> (III)	4.02 (4.26)	0.89 (0.65)	0.58 (0.82)	2.03 (2.02)	3.07 (3.07)	16SiO <sub>2</sub> · C <sub>6</sub> H <sub>11</sub> NO <sub>4</sub> · Co <sub>0.88</sub> Cl <sub>0.96</sub>

**Table 2.** Composition (%) of hybrid systems (calculated values are shown in brackets )

The surface area of the samples was determined by nitrogen adsorption in accordance with the BET method using a TriStar 3000 V.6.03A instrument. The instrumental error was 0.1 m<sup>2</sup>/g. The size of particles was estimated under the assumption that the particles were spherical.

The surface area (*S<sub>sp</sub>*) and the diameter of particles (*D*) were found to be 28.9 m<sup>2</sup>/g and 47 nm ; 123.9 m<sup>2</sup>/g and 7 nm ; 2.4 m<sup>2</sup>/g and 818 nm for chitosan-SiO<sub>2</sub> , chitosan-Al<sub>2</sub>O<sub>3</sub> and chitosan-MCC hybriide systems, respectively.

## 2.2. Saturation of samples with Cu<sup>2+</sup> and Co<sup>2+</sup> ions

### 2.2.1. Saturation of samples with Cu<sup>2+</sup> Ions

A 0.1 M NaNO<sub>3</sub> solution (10 ml) was added into weighed samples (200 mg) of nanostructured TiO<sub>2</sub>, and the samples were kept for one week at a constant solution pH (5.5) held by adding dilute NaOH and HNO<sub>3</sub> solutions. The sorption of Cu<sup>2+</sup> ions on nanostructured TiO<sub>2</sub> was performed by exposing samples in Cu(NO<sub>3</sub>)<sub>2</sub> solutions (10 ml)



with concentrations of  $10^{-4}$ ,  $10^{-3}$ , and  $10^{-2}$  mol/L and ionic strength ( $\mu$ ) of 0.1, which was adjusted using  $\text{NaNO}_3$ . Solution pH equal to 4.3 was maintained by the titration with small volumes of  $\text{NaOH}$  and  $\text{HNO}_3$  solutions. After the equilibrium was established, the residual amount of  $\text{Cu}^{2+}$  ions the equilibrium solutions was measured to determine the amount of sorbed  $\text{Cu}^{2+}$ . Then,  $\text{TiO}_2$  was separated from the solutions by centrifugation. The samples were washed twice with a 0.1 M  $\text{NaNO}_3$  solution (pH 4.3) to remove adsorbed  $\text{Cu}^{2+}$  -ions.

$\text{Cu}^{2+}$  ions were sorbed on  $\text{TiO}_2$  hydrogel from  $\text{CuCl}_2$  and  $\text{Cu}(\text{NO}_3)_2$  with subsequent its removal by filtration and drying at  $20^\circ\text{C}$  for 3 days upto constant weights of the precipitates. A volume of solution and a mass of hydrogel were changed to vary the content of  $\text{Cu}^{2+}$  ions in the phase of the studied  $\text{TiO}_2$  hydrogel, which was determined by the atomic absorption method on a Perkin Elmer 403 spectrometer. As the ESR spectra of  $\text{Cu}^{2+}$ -containing hydrated gels are difficult to record, hydrogel samples filtered and dried at room temperature were used. Preliminary experiments were performed to select the sample drying conditions preventing the structural changes of the complexes formed.

**$\text{Cu}^{2+}$  - containing composites based nanostructured  $\text{SiO}_2$ ,  $\text{TiO}_2$ , and cellulose powder** were prepared by sorption of  $\text{Cu}^{2+}$  ions on a hydrogel from an aqueous solution of  $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ . The volumes of hydrogels of the samples studied were calculated from the masses of xerogels obtained by hydrogels drying. For preparation of  $\text{Cu}^{2+}$ -containing  $\text{TiO}_2$  and  $\text{SiO}_2$  xerogels and the related CMs, a 0.01 g  $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$  containing 0.059 mmol of  $\text{Cu}^{2+}$  ions and 2 ml of  $\text{H}_2\text{O}$  and the calculated volumes of hydrogels prepared were added into flasks. The samples were kept in the contact with a  $\text{Cu}^{2+}$ -containing solution about 24 h upto the equilibrium was established. Then, the residual amounts of  $\text{Cu}^{2+}$  ions in the equilibrium solutions was measured to determine the amount of sorbed  $\text{Cu}^{2+}$  ions. The initial and residual amounts of  $\text{Cu}^{2+}$  -ions in a solution were measured using colorimeter KFK-2MP.

### 2.2.2. *Synthesis and characterization of the Cobalt-Containing Chitosan hybrid systems*

**Cobalt-Containing Chitosan-Supported Systems** were synthesised through stirring a mixture containing 0.24 g of  $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ , 2 g of chitosan-supported hybrid system and 20 ml of ethanol under reflux condenser for 24 h. The obtained cake was filtered off, rinsed with ethanol (15 ml  $\times$  3 times) and dried at room temperature until the weight became constant. The elemental compositions of the hybrid system surfaces were determined using an analytical setup based on a VEGA II LMH scanning electron microscope and an INCA ENERGY energy dispersive microanalysis system (Mekhaev et al., 2011a, 2011b). The data are shown in Table 2.

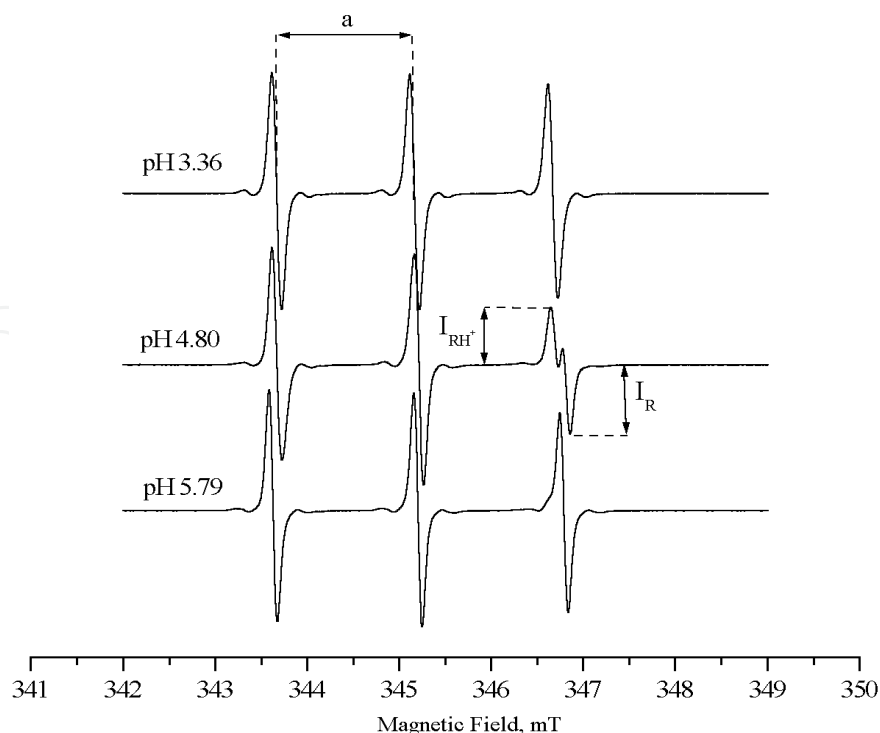
## 2.3. pH probes

The pH values of solutions inside pores and near the surface of the studied inorganic and organo-inorganic materials were determined using spin probes, namely, pH-sensitive NRs of the imidazoline (R1, R2) and imidazolidine (R3) types (Table 3), which were synthesized at the Novosibirsk Institute of Organic Chemistry, Siberian Branch, Russian Academy of Sciences (Volodarskii et al., 1988; Khramtsov et al., 1998; Kirilyuk et al., 2005).

## 2.4. Recording and processing of the ESR spectra of NR

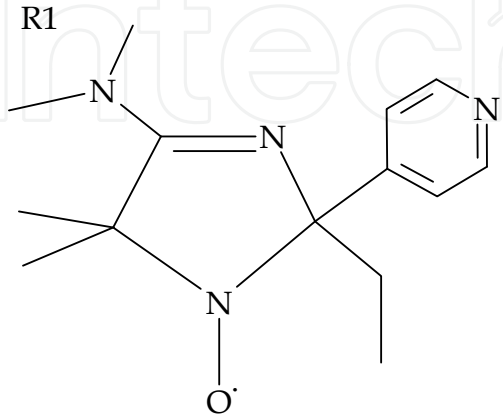
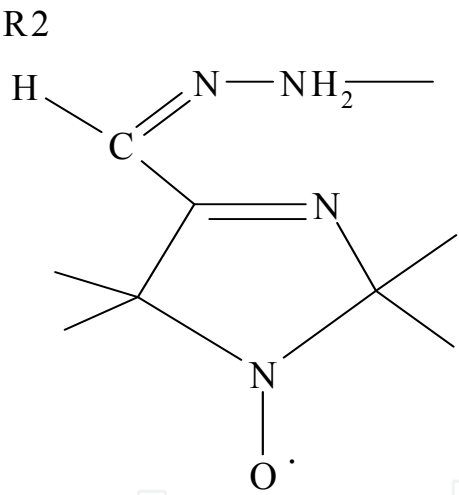
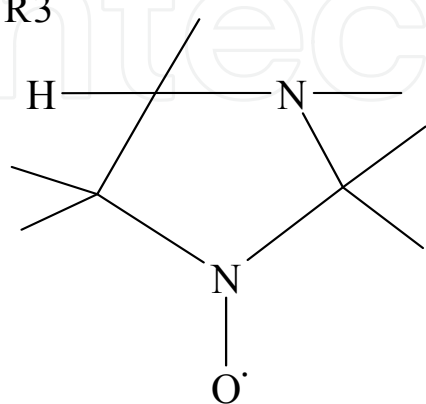
The ESR spectra were recorded on a PS 100.X ESR spectrometer (ADANI, Belarus) in a three-centimeter (X) wavelength range at room temperature. Quartz sample holders with an internal diameter of 3.5 mm were used for solid samples. Solution spectra were recorded using quartz capillaries.

Figure 1 shows characteristic ESR spectra of the pH-sensitive NR in aqueous solutions. According to the ESR theory, isotropic signals are induced by the fast-motivated NR molecules (correlation times of  $10^{-10}$  s and less) and present the triplet of fine lines. Depending on solution pH, NR can be in protonated ( $\text{RH}^+$ ), deprotonated (R), or intermediate (mixed) forms. Because hyperfine splitting constants  $a_N$  for  $\text{RH}^+$  and R forms of the radicals are different (Table 3), in their ESR spectra, the distance  $a$  between the low and central-field component of the triplet increases gradually with pH of a solution, from the values characteristic of the  $\text{RH}^+$  form to those typical of the R form (Fig. 1). This characteristic is a superposition of hyperfine splitting constants ( $a_N$ ) characterizing protonated and deprotonated forms of the nitroxide radical. From the results of measuring  $a$  values during titration, the calibration curves reflecting the dependences  $a$  vs. pH were plotted for each NR used (for example, see Fig. 2, 3 curve 1). In order to plot the calibration curves, NR solutions ( $10^{-4}$  mol/l ;  $\mu = 0.1$  ) were titrated either with dilute HCl and KOH solutions (used for all the samples, excepting hybriide systems)( Molochnikov et al.,2007 ; Parshina et al., 2011 ; Shishmakov et al., 2010) or citrate-phosphate (pH 3.5-7.8) and citrate-salt (pH 1.6-4.8) buffer solutions (used for hybriide systems) (Mekhaev et al., 2011a, 2011b). to vary the pH within the range of NR sensitivity of 2.5 – 7.5.

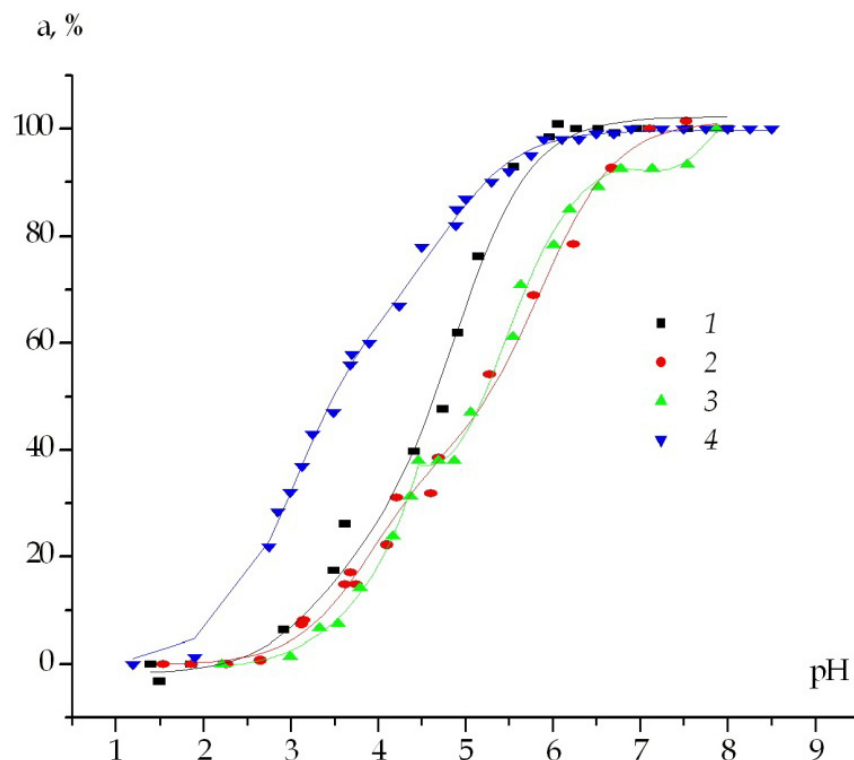


**Figure 1.** The ESR spectra of the aqueous solution of NR R3 at different pH in the X range of wavelengths at 293K.  $I_{\text{RH}^+}$  and  $I_{\text{R}}$  are the intensities of ESR peaks for  $\text{RH}^+$  and R forms of the radical, respectively



Radical	pK <sub>a</sub> (± 0.1)	g-factor (± 0.0001)		a <sub>N</sub> (± 0.006 mT)	
		R	RH <sup>+</sup>	R	RH <sup>+</sup>
<p>R1</p> 	3.15 4.89	2.0048	2.0051	1.520	1.390
<p>R2</p> 	3.55	2.0048	2.0051	1.590	1.515
<p>R3</p> 	4.70	2.0048	2.0051	1.590	1.485

**Table 3.** ESR paramters and pK<sub>a</sub> values of nitroxide radicals used



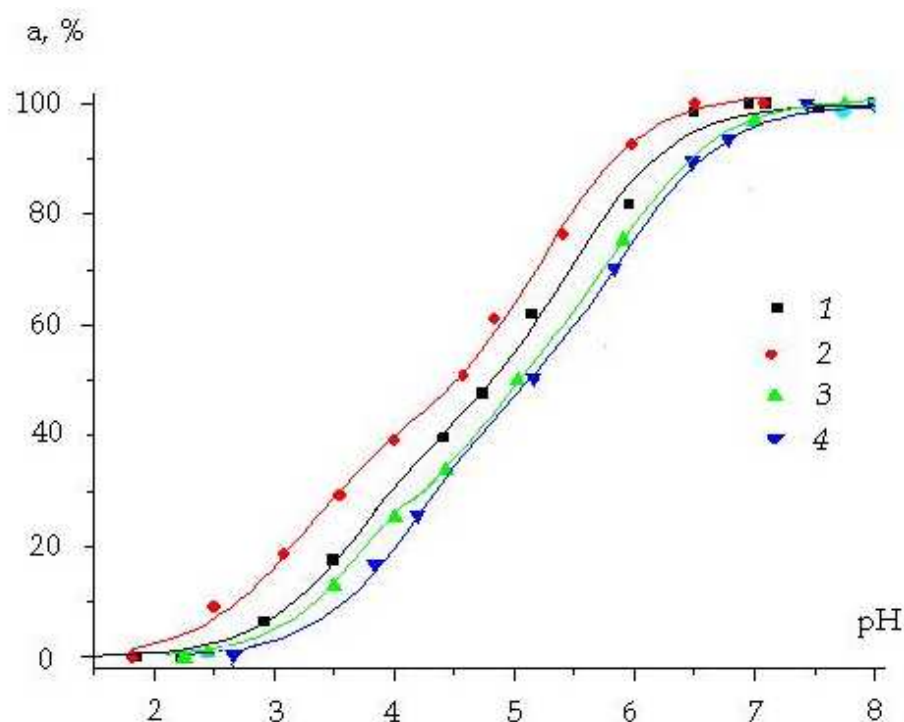
**Figure 2.** Titration curves for NR R1 in bulk aqueous solution (calibration curve) (1),  $\alpha$ - $\text{Al}_2\text{O}_3$  (2), the BS-50  $\text{SiO}_2$  (3) and  $\gamma$ - $\text{Al}_2\text{O}_3$  (4).  $a, \% = ((a - a_{\text{NRH}^+}) / (a_{\text{NR}} - a_{\text{NRH}^+})) \times 100\%$

## 2.5. Determination of pH in the pore and near the sample surface using pH-sensitive spin probes

An aqueous KCl solution (10 ml) with an ionic strength of 0.1 was added to an oxide sample (200 mg) and the mixture was allowed to stand for a preset time. Then, the solution was thoroughly decanted and an NR solution ( $10^{-4}$  mol/L,  $\mu = 0.1$ ) was added to the sample. In some cases, required initial pH values of radical solution were obtained by preliminary mixing of HCl and KOH solutions. After the equilibrium was established, the suspension was titrated with HCl and KOH ( $\text{HNO}_3$  and NaOH) solutions to plot the titration curve for the NR present in the sample.

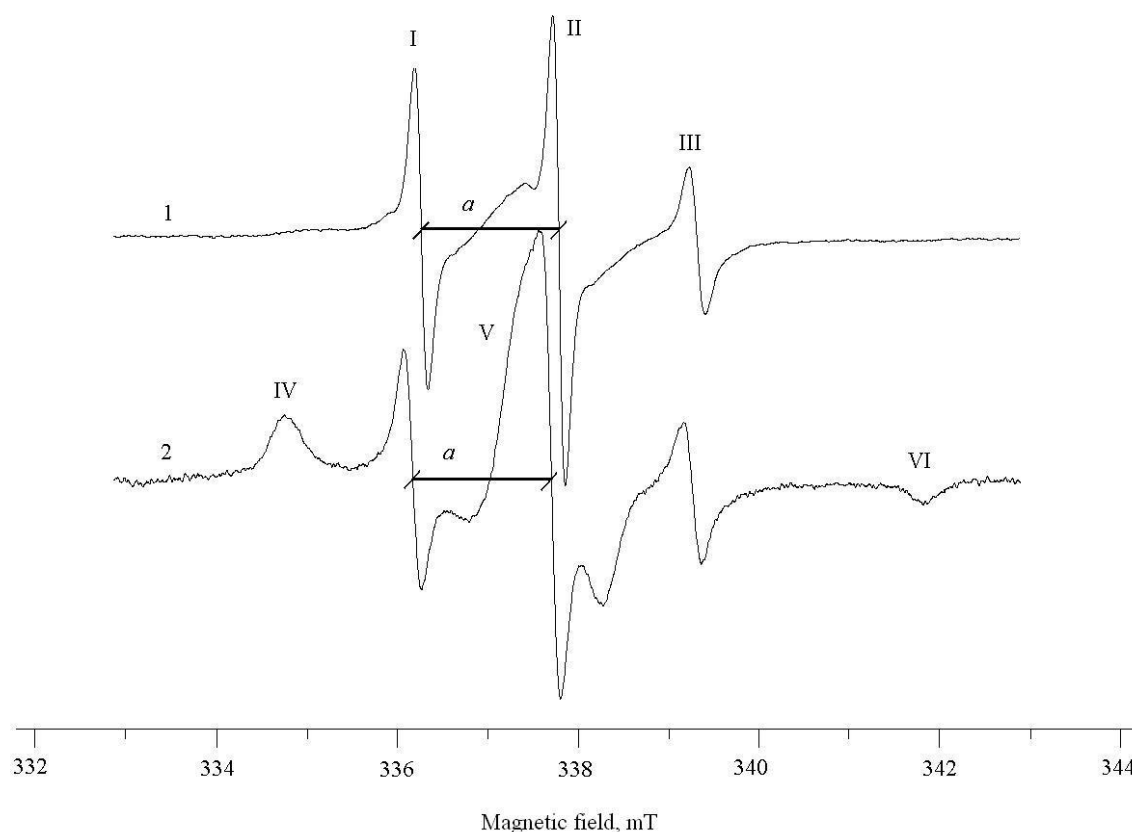
For chitosan cobalt-containing hybrid systems and solid-phase composites based on  $\text{SiO}_2$ ,  $\text{TiO}_2$  and cellulose powder the method of multiply batches was used: 0.05 g of sample was kept in 5 ml of buffer aqueous solution containing nitroxide radicals for 2 days (established experimentally). The solution was then decanted.

The pH values of the equilibrium solutions ( $\text{pH}_{\text{ext}}$ ) over the samples were measured using a Mettler Toledo pH meter (Switzerland) with an accuracy of 0.01 units. The samples separated from the solutions by centrifugation or filtration were placed into unsealed quartz ampoules and their ESR spectra were recorded. After measuring the  $a$  distances in the ESR spectra of corresponding radicals located in the samples (Fig. 4), the  $\text{pH}_{\text{int}}$  values of the studied materials were determined using the calibration curves (Fig. 2,3).



**Figure 3.** Titration curves for NR R1 in bulk aqueous solution (calibration curve) (1), PC (2),  $\text{SiO}_2$  (3) and  $\text{TiO}_2$  (4) xerogels.  $a, \% = ((a - a_{\text{NRH}^+}) / (a_{\text{NR}} - a_{\text{NRH}^+})) \times 100\%$

As can be seen from Fig. 4, the ESR spectra of NR in the samples studied represent the superpositions of three components of an isotropic signal of the probes in aqueous solutions inside pores and a spectrum of the probes immobilized on the surface of the objects studied. For determination of  $\text{pH}_{\text{int}}$  values, only the isotropic signals in the ESR spectra were used.



I, II, II – components of an isotropic spectrum ; IV, V, VI – components of the spectrum of the immobilized probes

**Figure 4.** The ESR spectra of NR R1 in the samples of CMs SiO<sub>2</sub>-PC (68% wt.SiO<sub>2</sub>) at pH 7.8 (1) and pH 3.7(2)

### 3. Results and discussion. Acid-base equilibria of inorganic and organo-inorganic materials and systems

#### 3.1. Inorganic oxide materials

##### 3.1.1. Pure oxides

The comparison of the titration curves for NR R1 radical in  $\gamma$ -Al<sub>2</sub>O<sub>3</sub>,  $\alpha$ -Al<sub>2</sub>O<sub>3</sub>, the BS-50 type SiO<sub>2</sub> and TiO<sub>2</sub> and SiO<sub>2</sub> xerogels with the calibration curve of this NR (Fig. 2, 3) indicates that for  $\gamma$ -Al<sub>2</sub>O<sub>3</sub>, the curve is shifted to the left, and for other samples, to the right. As was shown previously for organic sorbents (Molochnikov et al., 2004), the left-side shift of the titration curves for the NRs used is characteristic of anionites in acidic solutions in the pH range, where amino groups bind hydrogen ions, for example,

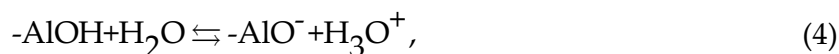
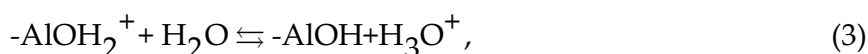


where  $P$  is the sorbent matrix. At the same time, the shift in the titration curves to the right is related to the dissociation of acidic functional groups of cationites, for example, a carboxylic one,



Thus, the shift of the titration curve of R1 to the left observed for  $\gamma\text{-Al}_2\text{O}_3$  indicates the binding of  $\text{H}^+$  ions with the surface ( $\text{pH}_{\text{intr}}$  is higher than solution pH) and to a positive charge of  $\gamma\text{-Al}_2\text{O}_3$  surface, while the right-hand side shift of the titration curve of this NR in  $\text{SiO}_2$  gel suggests the release of  $\text{H}^+$  ions and, consequently, a negative surface charge.

According to the data reported in (Hubbard, 2002) the acidic dissociation of hydrated aluminum oxide occurs in two stages, which can be described by the following equations:



The  $\text{pK}_a$  values of these equilibria ( $\text{pK}_{a1} = 5.87$  and  $\text{pK}_{a2} = 7.50$ ) are given in (Lidin et al., 1987). Using NR, we can determine only pH values relevant to the ascending part of its S-shaped calibration curve (Fig. 2, 3 curve 1), i.e., the range, in which variations in  $a$  parameter correspond to variations in pH. The range of pH values appropriate to the ascending part of the calibration curve for a given NR is referred to as the zone of its sensitivity. The sensitivity zone of NR R1 falls in the pH range from 2.5 to 6.5 (Fig. 2, 3). If the above data on  $\text{pK}_a$  reflect (even if qualitatively) the  $\text{pK}_a$  of functional group dissociation for  $\gamma\text{-Al}_2\text{O}_3$  studied, then the comparison of the sensitivity zone of NR R1 with  $\text{pK}_a$  values indicates that the use of NR allows one to study only the pH range in which  $\gamma\text{-Al}_2\text{O}_3$  either has a positive charge or is electrically neutral. On the other hand, there are data on the point of zero charge (PZC) for  $\text{Al}_2\text{O}_3$  in the literature.

For example, as was shown previously, the PZC of an  $\text{Al}_2\text{O}_3$  sample kept in water for a long time was equal to 9.2, and for  $\text{Al}_2\text{O}_3$  treated at  $1400^\circ\text{C}$ , the PZC decreased to 6.7 (Robinson et al., 1964). The first value corresponds to  $\gamma\text{-Al}_2\text{O}_3$ , and the second one, to  $\alpha\text{-Al}_2\text{O}_3$ . For hydrated  $\text{Al}_2\text{O}_3$  the PZC was found to be equal to 8, and it decreases as a result of the modification of  $\text{Al}_2\text{O}_3$  surface with alkyl phosphate-based surfactants (Jeon et al., 1996). Moreover, these data suggest that, in the studied pH range, the surface of  $\gamma\text{-Al}_2\text{O}_3$  has a positive charge that is in accordance with the shift of its titration curve to the left.

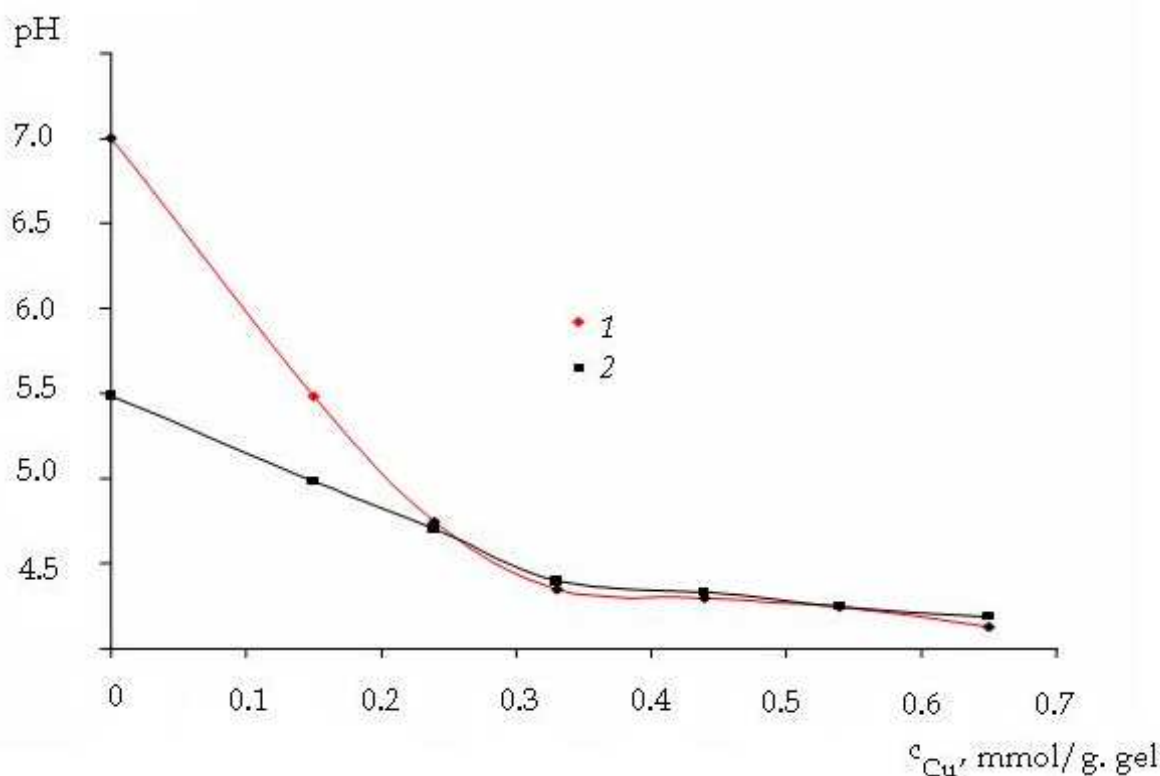
The  $\text{pH}_{\text{int}}$  of the sorbents differ from the  $\text{pH}_{\text{ext}}$  by 0.5–1.5 units (Table 3). Smaller  $\text{pH}_{\text{intr}}$  values compared to pH of the external solutions for all the studied samples (except for  $\gamma\text{-Al}_2\text{O}_3$ ) are indicative of the shift in titration curves of NR R1 to the right relative to the calibration curve, similarly to that demonstrated in Fig. 2, 3 all the samples excepting  $\gamma\text{-Al}_2\text{O}_3$ .

It should be noted that the  $\text{pH}_{\text{int}}$  value determined in this study for the hydrated  $\alpha\text{-Al}_2\text{O}_3$  is in good agreement with the PZC value found in (Robinson et al., 1964). From the data given in Table 3 we notice that in all the studied samples excepting  $\gamma\text{-Al}_2\text{O}_3$  acidic centers

predominate. However, the acidity of these centers is rather weak; their  $pK_a$  of dissociation are presumably close to 6. Only when strongly acidic sulfate residues appear in  $Al_2O_3$ , a drastic decrease in  $pH_{int}$  is observed as a result of their dissociation. It should be noted that the same decrease in  $pH_{int}$  of  $TiO_2$  hydrogel results from the sorption of  $Cu^{2+}$  ions by the hydrogel, which leads to the ion-exchange displacement of  $H^+$  ions (Fig. 5).

Sample	$pH_{ext} (\pm 0.01)$	$pH_{int} (\pm 0.1)$
$\alpha - Al_2O_3$	7.72	6.6
$Al_2O_3 - F^-$	6.9	6.4
$Al_2O_3 - SO_4^{2-}$	4.9	3.8
$\gamma - Al_2O_3$	7.16	>7.3
$TiO_2$ -hydrogel	7.0	5.5
BS-50 type $SiO_2$	7.5	5.7
$TiO_2$ -xerogel	7.0	6.1
$SiO_2$ -xerogel	7.0	6.3

**Table 4.** pH values of the external bulk solution of NR R1 and in the pores and near a surface of the inorganic materials studied



**Figure 5.** The pH values of a  $CuCl_2$  equilibrium solution ( $pH_{ext}$ ) (1) and a solution inside  $TiO_2$  hydrogel ( $pH_{int}$ ) (2) vs. the amount of sorbed  $Cu^{2+}$  ( $c_{Cu}$ )

The titration curves of radicals occurring in organic sorbents are characterized by the presence of horizontal plateaus corresponding to constant  $pH_{int}$  (parameter  $a$  remains unchanged), when the pH of the external solution increases (Molochnikov et al., 2004). The



presence of the horizontal plateaus on the titration curves is related to the consumption of the titrant solution for the neutralization of functional groups of ionites. There was found no horizontal plateau on the titration curve of NR R1 in  $\gamma$ -Al<sub>2</sub>O<sub>3</sub> within its zone of sensitivity (Fig.2). This fact can be explained either by the presence of a small number of acidic functional groups in  $\gamma$ -Al<sub>2</sub>O<sub>3</sub> or by noncoincidence of pK<sub>a</sub> values of these acid groups with the zone of sensitivity of the radical used. Based on our experimental data, it is difficult to make a conclusion in favor of any of these assumptions.

It should be noted that the titration curve of NR R1 in the BS-50 SiO<sub>2</sub> contains two horizontal plateaus corresponding two constant pH<sub>int</sub> determined near the SiO<sub>2</sub> surface when the pH of the external solution is changed (pH<sub>ext</sub>) (Fig.2, curve 3). These plateaus are referred to as the titration process of functional groups on the surface (Golovkina et al., 2008). The pK<sub>a</sub> values for these groups were determined from the titration curves, and were pK<sub>a1</sub> = 4.4–4.7 and pK<sub>a2</sub> = 6.5–6.8. According to (Long et al., 1999 ; Zhao et al., 1997, 1998) amorphous SiO<sub>2</sub> contains groups of three types: silanol, silandiol, and siloxane groups in a ratio of 59.2, 14.7, and 26.1%, respectively. We can therefore assume that the lower and longer horizontal plateau reflects the titration process of silanol functional groups on the SiO<sub>2</sub> surface, and the upper plateau corresponds to silandiol groups. According to (Méndez et al., 2003), the intervals of changes in the dissociation constants are pK<sub>1</sub> = 3.51–4.65 and pK<sub>2</sub> = 6.17–6.84, respectively. The same values of constants of dissociation for silanol and silandiol groups of silica were measured and published in (Nawrocki, 1997 ; Neue, 2000). The presence of silanol and silandiol groups on the surface of the studied SiO<sub>2</sub> was thus identified.

**From two types of the xerogels studied (TiO<sub>2</sub> and SiO<sub>2</sub>)** electrical potential of TiO<sub>2</sub> xerogel is slightly more than that of SiO<sub>2</sub> xerogel (Fig.3). There was found a small horizontal plateau in the range of pH<sub>intr</sub> between 3.75 and 4.25 on the titration curve for NR R1 near the SiO<sub>2</sub> surface. From this plateau pK<sub>a</sub> values for silanol groups were determined. They were found to be equal to  $3.95 \pm 0.07$  within the zone of constant pH<sub>int</sub> values (Molochnikov et al., 1996). The measured values of pK<sub>a</sub> are in good agreement with those determined before for silicas using chromatography (Méndez et al., 2003) and the method of pH probe in different silica-containing samples (Golovkina et al., 2008). At lesser pH values this curve goes near (but slightly righter) than the calibration curve of the above-mentioned NR. As seen from the figure, the surface of SiO<sub>2</sub> xerogel carries a small negative charge which is characteristic for the sample before titrating silanol groups. The lack of horizontal plateaus on the titration curves of NR R1 for the TiO<sub>2</sub> xerogel indicates the fact of titration of its functional groups at pH above 7. It can be explained by the presence of mainly terminal OH<sup>-</sup>-groups with basic properties on a surface of TiO<sub>2</sub> xerogel. The range of their titration lies above the zone of sensitivity of the NR used.

When studying acidity inside organic sorbents or hydrogels, pH probes are placed into water, which penetrates into a solid due to sample swelling. In this case, the probes occur in water environment far away from the chains of organic (synthetic ion-exchange resins) or inorganic (TiO<sub>2</sub> and SiO<sub>2</sub> hydrogel, xerogel) polymers.

The particles of nanostructured oxides in diameter of several nm possess extremely large surface (the ratio of the number of atoms on the surface of a nanoparticle to the number of atoms located inside the particle is equal to 1/7). Therefore, a pH-sensitive NR surrounded by water molecules can approach the hydrated surface of nanoparticles rather closely. It was expected that they as probes and labels would appear to be sensitive to SEP, which would make it possible to estimate its value.

The dependences of NR ionization constants on the surface electric potentials and the polarity of media were reported by Fromherz (Fromherz, 1989) and Khramtsov et al. (Khramtsov et al., 1992) for fluorescent pH indicators and for pH-sensitive NRs, respectively, as follows:

$$\Delta pK_a = \Delta pK_a^{el} + \Delta pK_a^{pol} \quad (5)$$

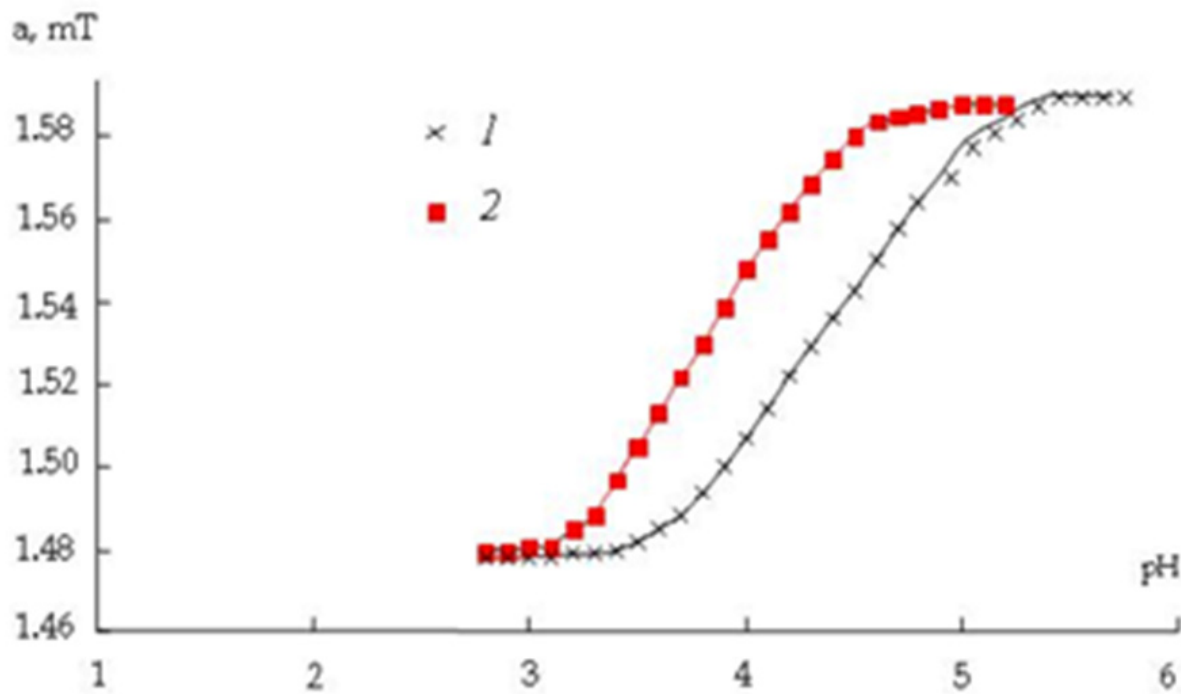
where  $\Delta pK_a$  is the total shift in the  $pK_a$  of an NR occurring in a sample relative to that of the NR in an aqueous solution and  $\Delta pK_a^{el}$  and  $\Delta pK_a^{pol}$  are the shifts in the  $pK_a$  of the NR occurring in the sample, which are due to the electric potential arising on the sample surface and to a change in the polarity of the medium relative to the polarity of the NR aqueous solution, respectively.

$$\Delta pK_a^{el} = -e \times \varphi / 2.3k \times T \quad (6)$$

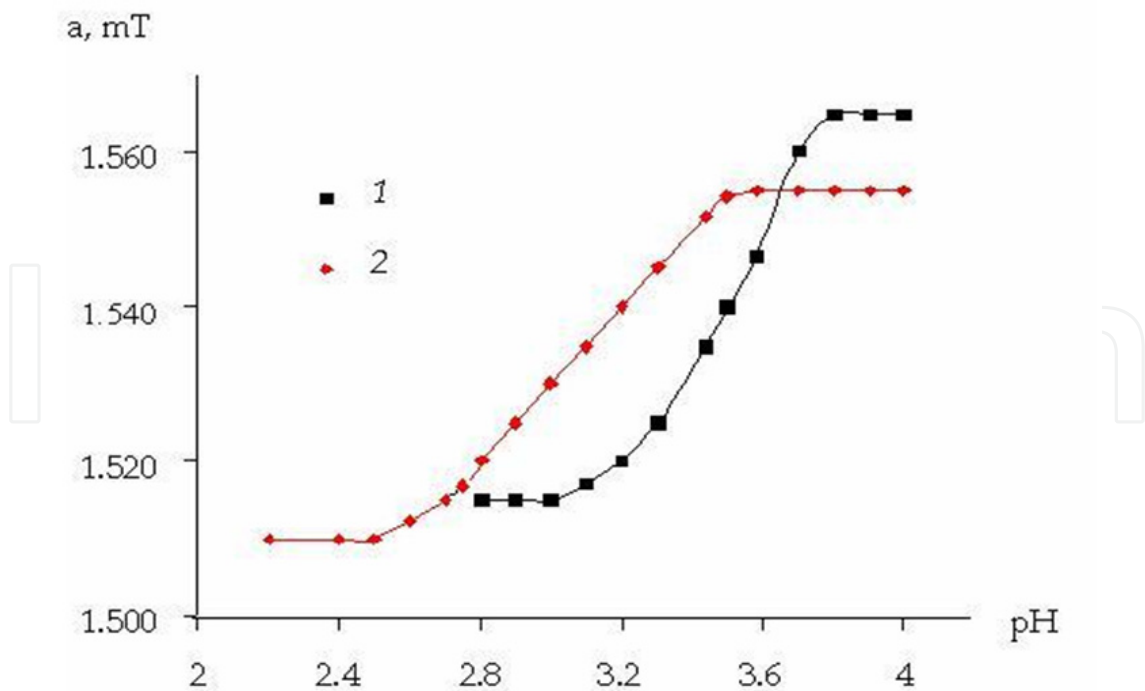
$$\Delta pK_a^{pol} = b \times (\varepsilon - \varepsilon_{HOH}) \quad (7)$$

where  $\varphi$  is the electrostatic potential,  $e$  is the electron charge,  $k$  is the Boltzmann constant,  $T$  is absolute temperature,  $b$  is a proportionality coefficient dependent on the radical nature,  $\varepsilon$  is the medium permittivity, and  $\varepsilon_{HOH}$  is the permittivity of water. Figures 6 and 7 demonstrate the results of studying nano-structured  $TiO_2$ . The measurements were performed using R3 probe at  $pH \approx 5.5$ , which was close to the PZC of  $TiO_2$  (5.7) (Poznyak et al., 1999), and R2 probe at  $pH \approx 4.3$ . Near the surface of  $TiO_2$  nanostructured particles, the titration curves of the radicals are shifted to the left, whereas, for  $TiO_2$  hydrogel and xerogel (Fig.3), a shift to the right was observed. At present, we can not find any explanation for such principal different shift in the titration curves for  $TiO_2$ -based samples. The  $TiO_2$  hydrogel obtained by hydrolysis was washed out with water, removing acidic components of the initial salts. Hereafter, the samples were prepared at different drying temperatures. The samples of xerogels and nanostructured oxides were dried at  $100^\circ C$  and  $200^\circ C$ , respectively.

As was mentioned above, the shift of the titration curve for  $\gamma-Al_2O_3$  to the left corresponds to the positive charge of the particle surface that, in principle, correlates with the data obtained at  $pH$  below PZC. The left-shift (toward more acidic solutions) of the titration curve of the probe located near the surface of  $TiO_2$  nanoparticles indicates that the surface electric potential is also positive. No changes were found in the hyperfine splitting constants  $a_N$  for both deprotonated and protonated R3 radicals (Fig. 6). The curves of titration of NR



**Figure 6.** Titration curves for NR R3 in bulk aqueous solution (calibration curve) (1) and near the surface of nanostructured TiO<sub>2</sub> (2) at 25°C and  $\mu = 0.1$



**Figure 7.** Titration curves for NR R2 (1) in bulk aqueous solution and (2) near the surface of nanostructured TiO<sub>2</sub> at 25°C and  $\mu = 0.1$ . The amount of Cu<sup>2+</sup> ions sorbed in the pores of TiO<sub>2</sub> is 0.36 mmol/g of the sample

R3 in the solution (curve 1) and near the surface of the nanostructured TiO<sub>2</sub> (curve 2) begin and end at the same value of parameter  $a$  characterizing the hyperfine splitting constant  $a_N$  for the protonated and deprotonated forms of NR. This fact shows that there are no changes in the polarity of the medium on the surface of nanostructured TiO<sub>2</sub> relative to the external solution, and the term  $\Delta pK_a^{pol}$  calculated by formula (7) is reduced to zero (Khramtsov et al., 1992). Thus, the shift in  $pK_a$  is determined by the electric potential of nanoparticle surface alone ( $\Delta pK_a = \Delta pK_a^{el}$ ). The calculation by formula (6) gives  $\varphi = 31.7$  mV at the point of R3 probe location. It should be stressed that the calculated  $\varphi$  value is not the electric potential of TiO<sub>2</sub> nanoparticle surface, but it only characterizes the electric field generated by a nanoparticle at the site where the radical fragment  $-N-O\bullet$  of R3 probe is located.

### 3.1.2. Cu<sup>2+</sup>-containing samples

**For TiO<sub>2</sub> hydrogel**, pH values of the external solution, which is in contact with the hydrogel ( $pH_{ext}$ ), and of those inside the hydrogel ( $pH_{int}$ ) during the sorption of Cu<sup>2+</sup> ions from copper(II) chloride and nitrate solutions are found to be different (Fig. 5). It was observed that the anions of these copper(II) salts had no effect on the acidity of a medium. In the process of sorption Cu<sup>2+</sup> ions are involved in the chemical interaction with the active centers of the hydrogel (Kharchuk et al., 2004) that manifests itself both in the parameters of the ESR spectra of Cu<sup>2+</sup> in the hydrogel and in the impossibility of washing out sorbed Cu<sup>2+</sup> ions from the hydrogel by Na<sup>+</sup> and Ca<sup>2+</sup> nitrate solutions. According to the data obtained, initially,  $pH_{int}$  is lower by 1.5 units than  $pH_{ext}$  (Table 3). As the amount of Cu<sup>2+</sup> ions in the hydrogel grows,  $pH_{int}$  decreases (Fig. 5), and, beginning with an amount of 0.24 mmol/g of the gel, it coincides with  $pH_{ext}$ . Thus, the acidity inside the original hydrogel sample is considerably lower than that of the equilibrium external solution. Similar regularities were found for granulated organic ion-exchange resins (Kovaleva et al., 2000). For KB-2 × 4 cationite in a mixed H<sup>+</sup>–Na<sup>+</sup> form, it was found that in the process of absorption of copper, the acidity of the external solution diminishes, while the acidity inside the sorbent remains unchanged until a certain amount of Cu<sup>2+</sup> ions in the ionite is attained, with this content being dependent on the fraction of Na<sup>+</sup> form of the cationite. When this value is exceeded, the acidity inside the sorbent begins to decrease and  $pH_{ext}$  and  $pH_{int}$  values for the ionite become equal at a certain degree of saturating with Cu<sup>2+</sup> -ions. For TiO<sub>2</sub> hydrogel, the equalization of pH values is also observed, but the horizontal plateau with a constant  $pH_{int}$  value on the titration curve of NR is not observed. The constant  $pH_{int}$  values for KB-2 × 4 cationite are explained by the buffer properties of its H<sup>+</sup>–Na<sup>+</sup> mixed form. TiO<sub>2</sub> hydrogel has no buffer properties, and the acidity increases smoothly with a rise in the amount of sorbed Cu<sup>2+</sup> ions.

**For nanostructural TiO<sub>2</sub>**, in the presence of sorbed Cu<sup>2+</sup> ions, at pH close to 4.3 (Fig. 7), like in the case of pure nano-TiO<sub>2</sub> (Fig. 6), a shift of the titration curve toward lower pH values is observed; however, the hyperfine splitting constant  $a_N$  decreases (the titration curve is shifted downward). The decrease in  $a_N$  is probably related to a lower polarity (the effective permittivity of the medium) near the nanoparticle surface where the probe is located. The theory predicts that the permittivity of a medium must have the same effect on both

electroneutral (the deprotonated form of the radical) and charged (the protonated form) species (Griffith et al., 1974). We found that the rise in the  $\text{Cu}^{2+}$  amount from 0 to 0.36 mmol/g. of the gel results in an increase in the  $\Delta\text{pK}_a$  value. The shift of this curve ( $\Delta\text{pK}_a$ ) to the left is significantly smaller than for NR R3, although the change from the PZC to pH 4.3 and the incorporation of  $\text{Cu}^{2+}$  ions must enhance the positive charge of the nanoparticle surface and, according to formula (6), increase in  $\Delta\text{pK}_a$ . In accordance with formula (7), a reduction in the permittivity near the surface of  $\text{TiO}_2$  particles compared to  $\epsilon_{\text{HOH}}$ , which causes the observed decrease in  $a_N$ , must also shift the titration curve to the left.

Hence, the explicit inconsistency between the determined  $\Delta\text{pK}_a$  values for NR R2 and NR R3 radicals and the theoretical predictions is established. We relate this difference to a larger distance of NR R2 from the surface compared to NR R3; this is possible because NR R2 has a substituent in position 4, which is presumably positioned toward the surface due to the tendency of its amino groups to protonation (or complexation with  $\text{Cu}^{2+}$  ions sorbed on the surface). Thus, the  $-\text{N}-\text{O}\bullet$  fragment of the radical turns out to be removed from the surface of  $\text{TiO}_2$  by the chain length of this substituent. The electric potential induced on the  $-\text{N}-\text{O}\bullet$  fragment will decrease with the distance from the surface and will be inversely proportional to the distance from this surface, when it is represented as, for example, a plane. The small NR R3 radical can approach the nanoparticle surface much closer and, therefore, it is affected more strongly by the electrostatic potential of the  $\text{TiO}_2$  surface. The sorption of  $\text{Cu}^{2+}$  ions on the surface of nanoparticles increases the charge of the latter. The observed shift of the titration curve (Fig. 7) and the broadening of the ESR spectra of these radicals attest to the orientation of the substituent in position 4 of NR R2 toward the surface and to the interaction between amino groups of the substituent and the  $\text{Cu}^{2+}$  ions sorbed on the nanoparticle surface.

### 3.2. Organic supports of composite and hybride materials

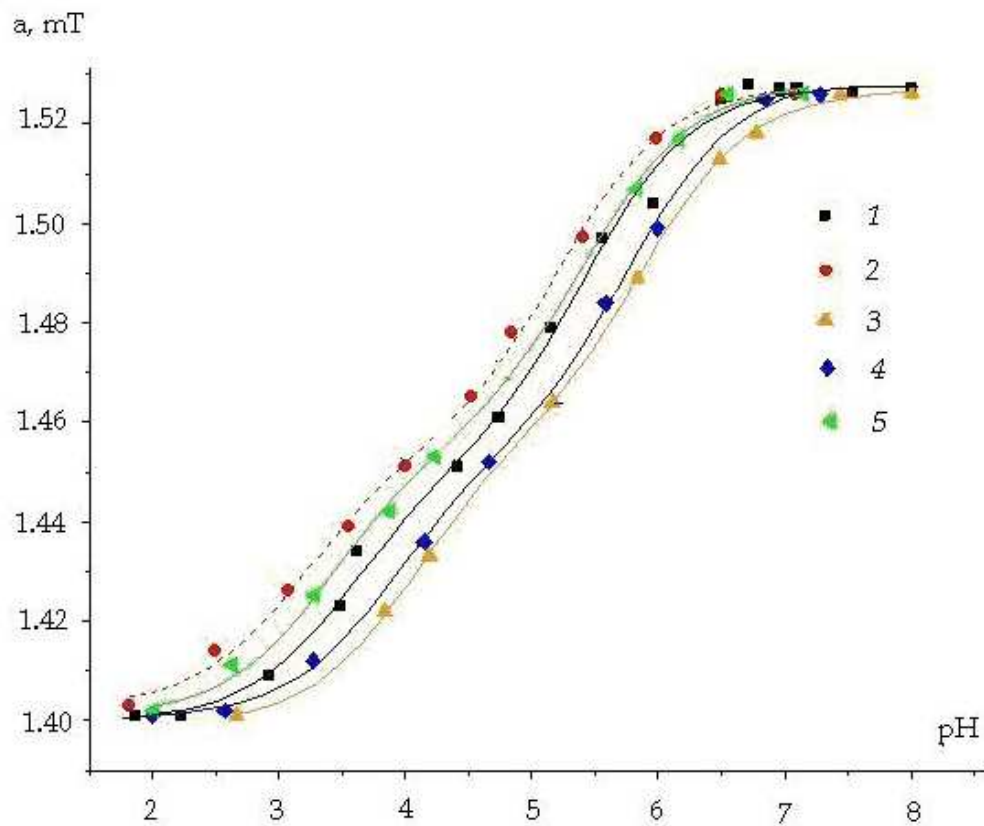
For PC and MCC, the titration curves of NR R1 were found to be shifted to the left of the calibration curve (Fig. 2, 3), indicating a positive charge on the surface (Kovaleva et al., 2000; Molochnikov et al., 2007). Hydroxyl groups of cellulose most likely play the role of surface bases whose protonation gives the surface a positive charge.

### 3.3. Solid-phase composites based on $\text{SiO}_2$ , $\text{TiO}_2$ and cellulose powder

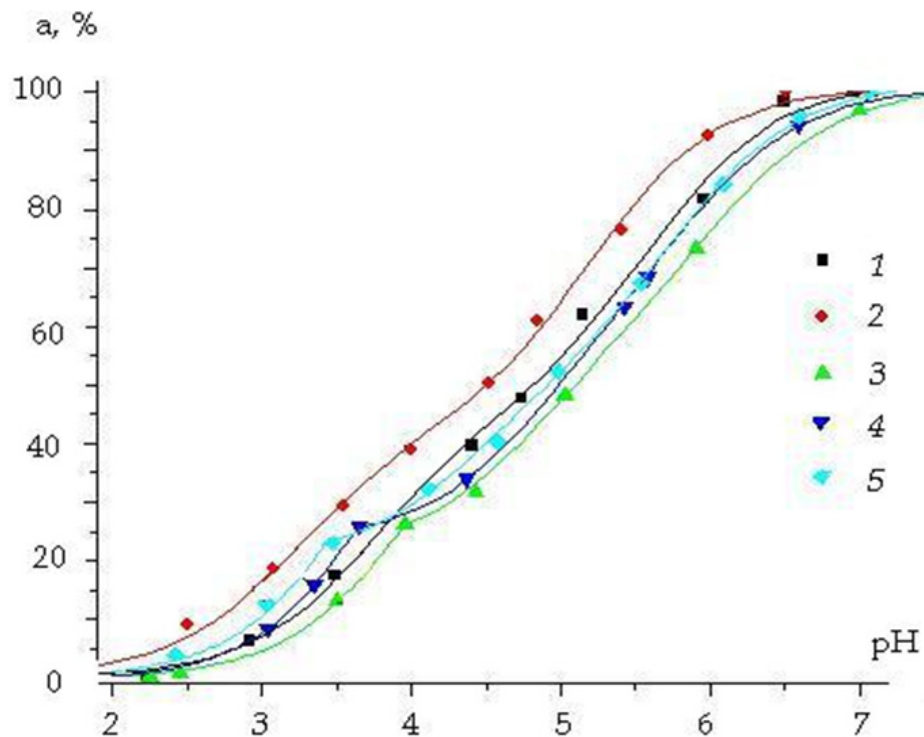
#### 3.3.1. Pure systems

As shown in Fig. 8 and 9, the titration curves of NR R1 in the CMs  $\text{TiO}_2$  – PC with 70 (not shown) and 53 % wt.  $\text{TiO}_2$  and in the  $\text{SiO}_2$  – PC with 68 and 35 % wt.  $\text{SiO}_2$  (unprotonated parts) are shifted to the right relative to the calibration curve, as in the case of  $\text{TiO}_2$  and  $\text{SiO}_2$  xerogels. For CMs based on  $\text{SiO}_2$ ,  $\text{TiO}_2$  and cellulose powder  $\Delta\text{pH} = \text{pH}_{\text{ext}} - \text{pH}_{\text{int}}$  decreased (the curves are shifted to the left relative to those of pure xerogels) as the PC content in the samples increased; this corresponds to a decrease in the negative charge of the CM surface. This is because of the lower acidity of cellulose compared to the acidity of the solution (cellulose has basic alcohol functional groups in its structure) and the positive charge of its surface (Parshina, 2011).





**Figure 8.** Titration curves for NR R1 in bulk solution (1), PC (2), TiO<sub>2</sub> xerogel (3) и CMs TiO<sub>2</sub>: PC (53% wt.TiO<sub>2</sub>)(4) and TiO<sub>2</sub>: PC (43% wt.TiO<sub>2</sub>) (5)



**Figure 9.** Titration curves for NR R1 in bulk solution (calibration curve) (1), PC (2), SiO<sub>2</sub> xerogel (3) and CMs SiO<sub>2</sub>: PC (58% wt. SiO<sub>2</sub>) (4) and SiO<sub>2</sub>: PC (35% wt. SiO<sub>2</sub>) (5).  $a, \% = ((a - a_{NRH^+}) / (a_{NR} - a_{NRH^+})) \times 100\%$



This proves that a surface of the composites studied carries lesser negative charge as compared to that of pure xerogels. The left-shift of the curve of NR R1 in CM TiO<sub>2</sub> – PC with 43% wt.TiO<sub>2</sub> relative to the calibration curve can be explained by initial positive charge of the surface of this CM due to binding H<sup>+</sup> ions. Hence, SEP of CMs based on TiO<sub>2</sub> and PC is varied over a wide range and even changes its sign from negative to positive as an increase in PC content. From the curves plotted, we notice that a decrease in percentage of TiO<sub>2</sub> xerogel in CMs from 53% to 43% leads to a pH<sub>intr</sub> greater than a pH<sub>ext</sub>. This fact can be explained by positive charge of a surface of cellulose due to the presence of alcohol groups in its structure.

For all the synthesized CMs based on SiO<sub>2</sub> xerogel and PC after complete protonating silanol groups, the titration curves of NR R1 were shifted to the left relative to the calibration curve and, hence, a surface of the samples studied remained positively charged (Fig.9). Unlike the composites based on TiO<sub>2</sub> and PC, an increase in a percentage of PC in the SiO<sub>2</sub>-PC composites leads to changing a surface charge from positive to negative with increasing pH<sub>ext</sub> (above the horizontal plateau on the titration curves) due to dissociation of functional groups. Thus, the incorporation of PC into the samples doesn't change acidity of silanol groups and doesn't make a polarizing effect on the SiO-H bond. The length of the horizontal plateau slightly increases in the accordance with the amount of silanol groups and the percentage of PC in the CMs based on SiO<sub>2</sub>. This can be caused by the increased dispersity of SiO<sub>2</sub> due to rising in S<sub>sp</sub> of the CMs (Table 1).

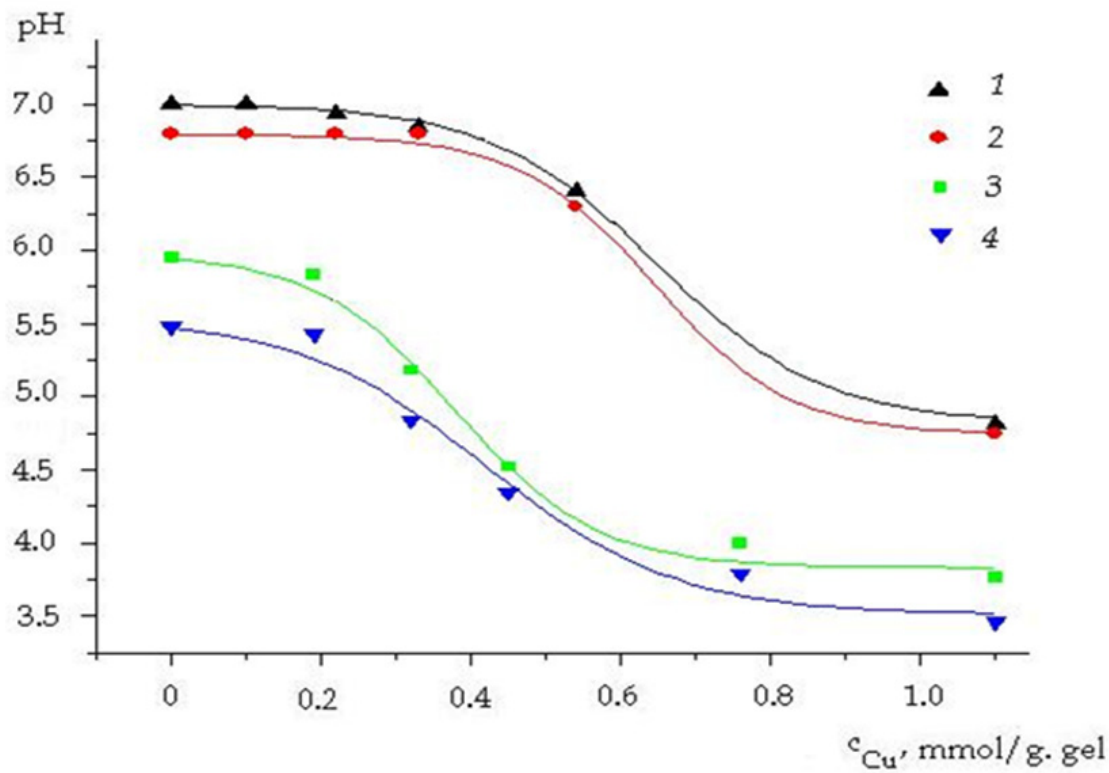
By varying the cellulose percentage in the composites and pH<sub>ext</sub> values, pH<sub>int</sub> values and SEP can be selected over a wide range. This information is needed for optimization of the conditions for pH-dependent adsorption and catalytic processes through a choice of CM with a certain pH<sub>int</sub> and SEP as catalyst substrate and adsorbent.

### 3.3.2. Cu<sup>2+</sup>-containing composites

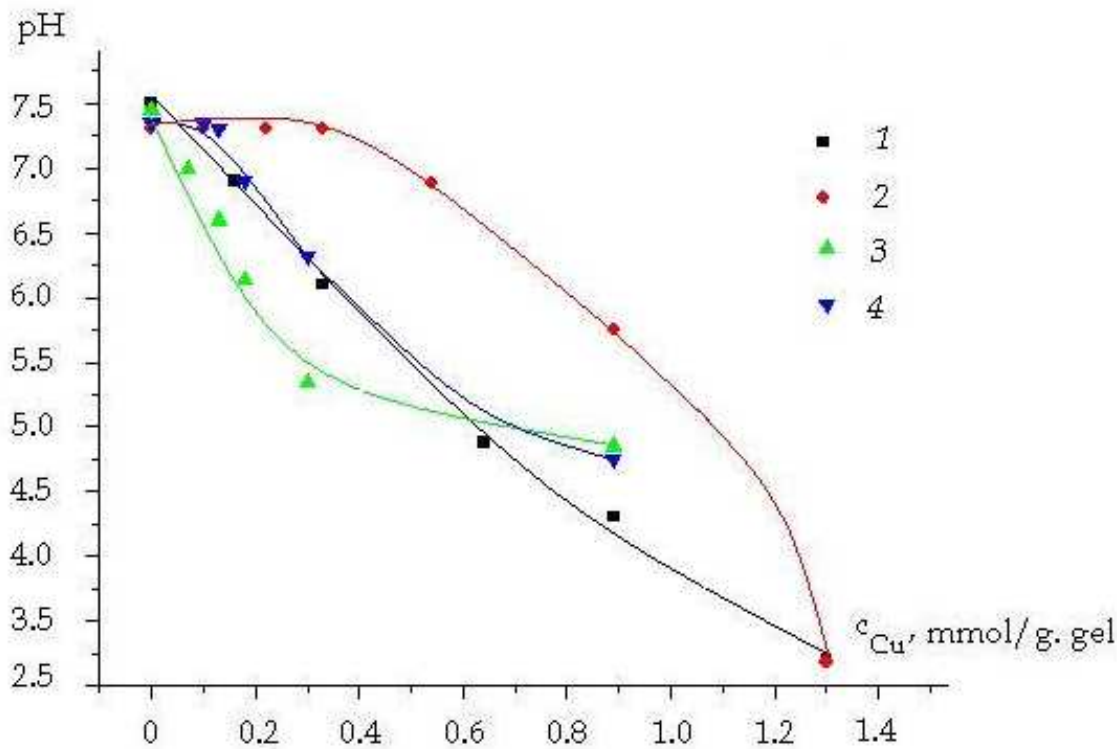
The sorption of Cu<sup>2+</sup> on the TiO<sub>2</sub> and SiO<sub>2</sub> xerogels and on the related CMs is accompanied by a change in the pH in their phases. However, the dependences of pH<sub>ext(int)</sub> vs c<sub>Cu</sub> (amount of sorbed Cu<sup>2+</sup> ions) for the above- mentioned types of xerogels and the CMs based have some differences (Fig.10, 11).

From these figures we notice that :

- pH<sub>int</sub> is different from pH<sub>ext</sub>, and with no Cu<sup>2+</sup> in an external solution (c<sub>Cu</sub> = 0), pH<sub>ext</sub> is greater than pH<sub>int</sub>.
- With an increase in the amount of Cu<sup>2+</sup> in the CMs, both the pH<sub>ext</sub> and pH<sub>int</sub> decrease.
- There is a horizontal plateau on the dependences of pH<sub>int</sub> on c<sub>Cu</sub> in the range of sorbed Cu<sup>2+</sup> from 0.15 up to 0.3 mmol Cu<sup>2+</sup>/g. TiO<sub>2</sub>, SiO<sub>2</sub> xerogel, within which pH<sub>int</sub> remains constant with increasing in c<sub>Cu</sub>.
- At greater c<sub>Cu</sub> pH<sub>int</sub> and pH<sub>ext</sub> come closer and become almost equal.



**Figure 10.** pH of external solution ( $pH_{ext}$ ) (curves 1,3) and near the surface ( $pH_{int}$ ) (curves 2,4) of  $TiO_2$  xerogel (curves 1, 2) and the  $TiO_2$  : PC (47% wt.  $TiO_2$ ) composite (curves 3,4) vs.  $c_{Cu}$



**Figure 11.** pH of external solution ( $pH_{ext}$ ) (curves 1,3) and near the surface ( $pH_{int}$ ) (curves 2,4) of  $SiO_2$  xerogel (curves 1,2) and the  $SiO_2$  : PC (68% wt.  $SiO_2$ ) composite (curves 3,4) vs.  $c_{Cu}$

The increase in acidity of external solution, first of all, is caused by the hydrolysis of  $\text{Cu}^{2+}$  at the initial  $\text{pH}_{\text{ext}}$  equal to 6-7 and is accompanied by intense release of  $\text{H}^+$  ions. The lesser  $\text{pH}_{\text{int}}$  values as compared to  $\text{pH}_{\text{ext}}$  for both types of xerogels and the related CMs can be explained by a surface negative charge. A surface of the samples studied attracts  $\text{H}^+$  ions in its double electrical layer to compensate this charge. Also an increase in acidity of both external and internal (near the surface) solutions is caused by the fact that the sorption of  $\text{Cu}^{2+}$  ions is accompanied by intense release of  $\text{H}^+$  ions into them. The increase in the acidity of a medium due to the competition of  $\text{H}^+$  ions almost completely suppresses the ion exchange sorption of  $\text{Cu}^{2+}$ . An analogous tendency was previously observed for organic ion-exchange resins (as grains) and  $\text{TiO}_2$  hydrogel (Kovaleva et al., 2000 ; Molochnikov et al., 2004).

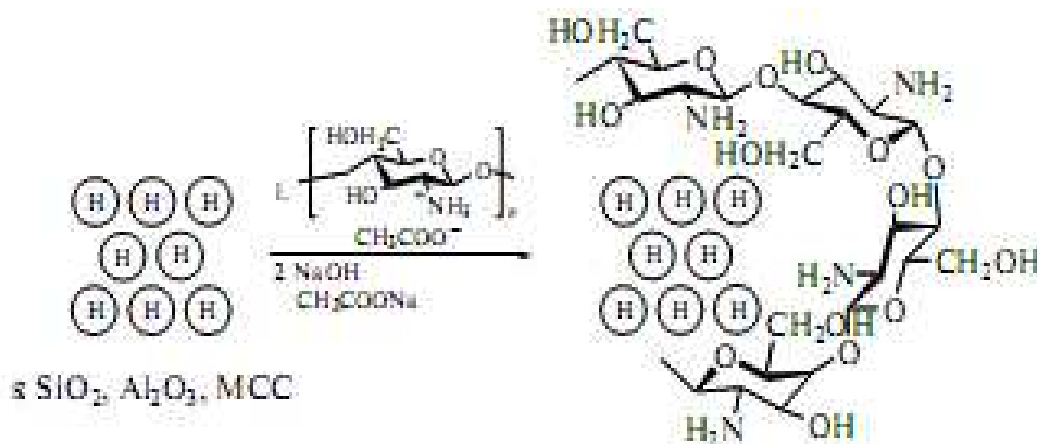
The presence of the horizontal plateaus on the curves  $\text{pH}_{\text{int}}$  vs  $\text{ccu}$  (Fig. 9) indicates the buffer properties of the  $\text{TiO}_2$  xerogel and the  $\text{TiO}_2$  : PC samples by analogy with the same properties of the mixed  $\text{H}^+\text{--Na}^+$  form of the KB-2  $\times$  4 cationite (Kovaleva et al., 2000). The greater is percentage of PC in a sample, the shorter is the horizontal plateau on the above-mentioned curves. The buffer effect of the  $\text{TiO}_2$  xerogel and the related CMs is caused by the existence of some amount of deprotonated functional groups in the samples before the sorption of  $\text{Cu}^{2+}$ . The complexation of  $\text{Cu}^{2+}$  ions with these groups occurs first of all and does not lead to the release of  $\text{H}^+$  ions, which would reduce both  $\text{pH}_{\text{ext}}$  and  $\text{pH}_{\text{int}}$ . In addition, releasing of hydroxide groups as a result of dissociation of molecules of  $\text{Cu}(\text{OH})_2$  responsible for complexation with functional groups of a surface of xerogel can retard decreasing pH with increasing  $\text{ccu}$ .

Unlike  $\text{TiO}_2$  systems, for  $\text{SiO}_2$  and the related CMs  $\text{pH}_{\text{int}}$  values were found to be more than  $\text{pH}_{\text{ext}}$  ones (Fig. 11). Apparently, this difference can be explained by the significant differences in  $\text{pK}_a$  of functional groups of  $\text{SiO}_2$  and  $\text{TiO}_2$  oxides.  $\text{pK}_a$  values of silanol groups ( $3.95 \pm 0.07$ ) are significantly lower than  $\text{pH}_{\text{ext}}$  from which sorption of  $\text{Cu}^{2+}$  has been conducted. Therefore, silanol groups are completely dissociated and  $\text{H}^+$  ions don't exchange with  $\text{Cu}^{2+}$  ions during complexation. Hence, no functional groups can neutralize  $\text{OH}^-$ -groups forming during breakdown of  $\text{Cu}(\text{OH})_2$  molecules. As a result,  $\text{pH}_{\text{int}}$  values (near a surface of samples) have high values.

As the titration curves for  $\text{TiO}_2$ -based CMs don't contain the horizontal plateaus as well as pure systems, hence,  $\text{pK}_a$  values of active acidic centers of a surface of these materials are out of zone of sensitivity of the NR used (greater than 7) (Fig. 8). Therefore, these centers are mainly in the protonated form before sorption of  $\text{Cu}^{2+}$ -ions. During complexation  $\text{Cu}^{2+}$ -ions exchange with  $\text{H}^+$  ions of functional groups releasing them into external solution. Hence,  $\text{pH}_{\text{int}}$  became less than  $\text{pH}_{\text{ext}}$ . It shows up in a general decrease pH (both  $\text{pH}_{\text{int}}$  and  $\text{pH}_{\text{ext}}$ ) with increasing  $\text{ccu}$  (Fig. 10 ).

### 3.4. Hybrid organo-inorganic materials based on the chitosan– $\text{Al}_2\text{O}_3$ , chitosan– $\text{SiO}_2$ , and cellulose –chitosan–systems

Hybrid systems were obtained in accordance with the scheme (Fig.12)



**Figure 12.** Scheme for obtaining hybrid organo-inorganic materials. s -support

The compositions of the obtained hybrid systems were characterized by elemental analysis (Table 2). According to the obtained data, the reactivity of substrates was different and declined in the order  $\text{MCC} > \text{Al}_2\text{O}_3 > \text{SiO}_2$ , since the compositions of the obtained hybrids in the case of inorganic oxides do not correspond to the molar ratios of the initial components.

The presence of chitosan molecules on the surface of the obtained systems was confirmed by the IR spectra, which contain characteristic absorption bands at 1652 and 1555  $\text{cm}^{-1}$  corresponding to residual acetoamide groups of the polymer.

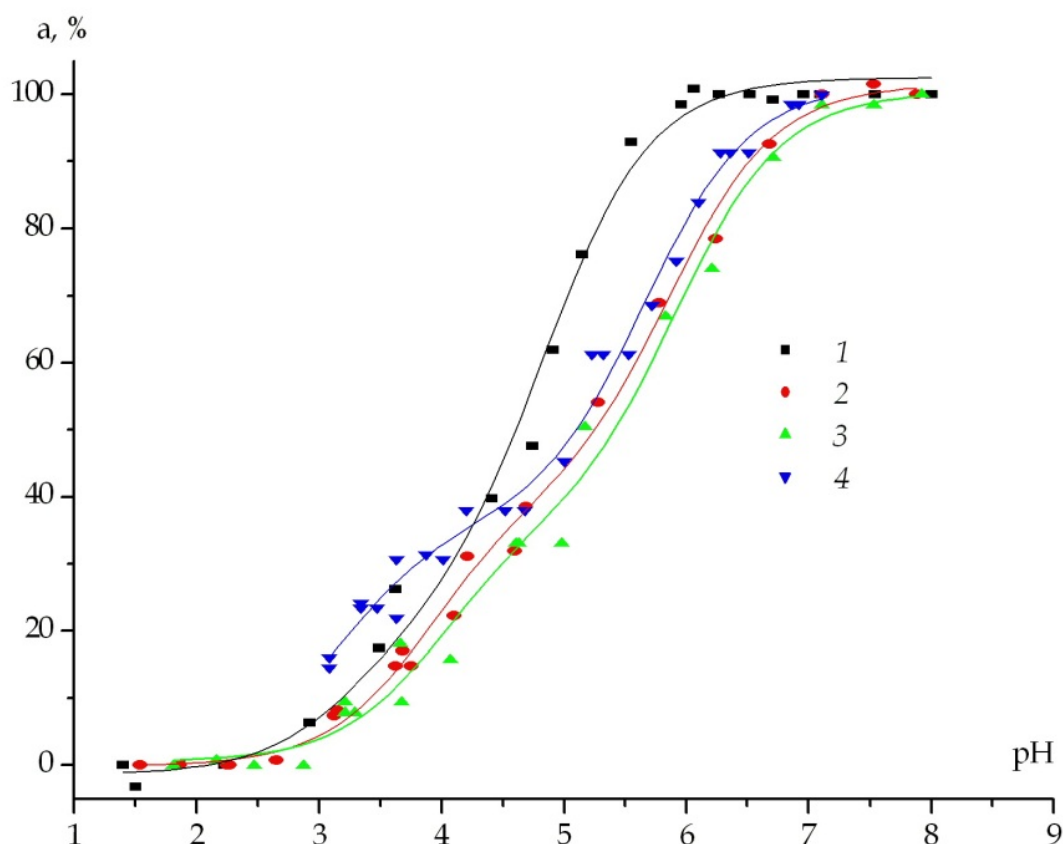
Since the obtained composites have to be electroneutral, we must assume that in the case of  $\text{SiO}_2$ , the substrate, being a weak acid, forms an ionic bond with the chitosan's amino group. In case of  $\text{Al}_2\text{O}_3$ , the same situation is possible, but some of the chitosan's amino and hydroxyl groups are involved in complex-forming with aluminum. In the case of MCC, the interaction occurs through hydrogen bonds.

NR were used as pH-probes to obtain a more detailed characterization of the surface structure of hybrid systems, their acid–base properties.

A comparison of the ESR spectra of NR R1 in solution and in the samples studied showed that in all cases, there was an isotropic signal which indicated the lack of covalent bonds between the NR and the sample surface and the presence of this radical near the surfaces of particles (Fig. 4, spectrum 1).

### 3.4.1. Pure hybride systems

Analysis of the structure of the  $\text{Al}_2\text{O}_3$ –chitosan system according to the adsorbed probe molecules of the nitroxyl radical showed that coating the initial  $\alpha\text{-Al}_2\text{O}_3$  substrate with chitosan leads to a slight increase in the surface negative charge, as was confirmed by the shift of the NR titration curve to the right (Fig. 13). The titration curves of  $\alpha\text{-Al}_2\text{O}_3$  and  $\text{Al}_2\text{O}_3$ –chitosan are not parallel to the calibration curve at low pH values due to the gradual dissolving of  $\alpha\text{-Al}_2\text{O}_3$  particles in weakly acidic media, as has been noted during investigations of samples containing aluminate ions (Iller, R., 1979 ; Golovkina, 2009).

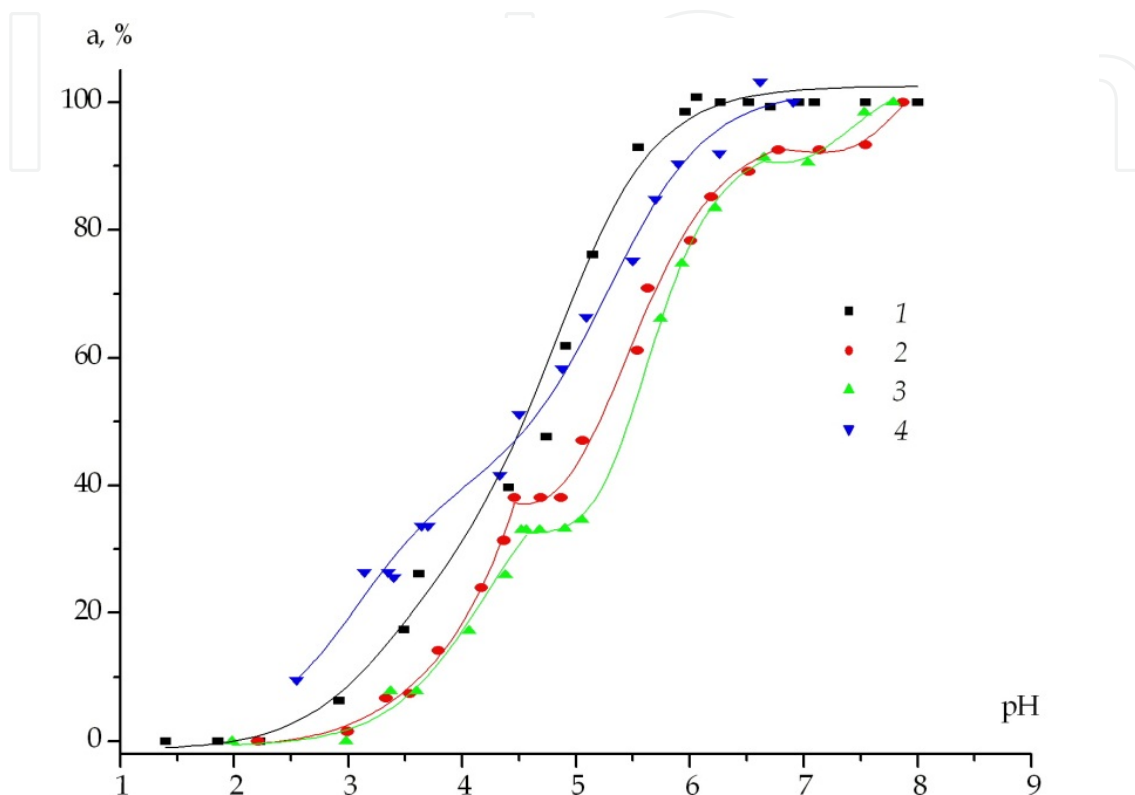


**Figure 13.** Titration curve for NR R1 in aqueous bulk solution ( calibration curve) (1), near the surface of  $\text{Al}_2\text{O}_3$  (2), hybrid  $\text{Al}_2\text{O}_3$ - chitosan (3) and  $\text{Al}_2\text{O}_3$ - chitosan- $\text{Co}^{2+}$  (4) systems .  $a, \% = ((a-\text{anRH}^+)/(\text{anR}-\text{anRH}^+)) \times 100\%$

For the  $\text{SiO}_2$ -chitosan system, the NR titration curve overlaps the corresponding curve for  $\text{SiO}_2$  values within the horizontal section of the titration curve for silandiol groups in the range of high pH (Fig. 14). We can correspondingly claim that in this pH region, the surfaces of the samples all have the same charge, demonstrating that chitosan has no influence on the  $\text{SiO}_2$  surface. If pH is decreased below 6, the titration curve for the  $\text{SiO}_2$ -chitosan system shifts to the right of the corresponding curve for  $\text{SiO}_2$ , indicating the negative charge of the surface. The horizontal plateau of the titration curve for silanol groups is lower in the case of the  $\text{SiO}_2$ -chitosan system (pH 4.5–5), and the  $\text{pK}_a$  value for silanol groups in the presence of chitosan falls slightly. A similar reduction in this parameter has been observed during modification of the surface of mesoporous molecular sieves based on  $\text{SiO}_2$  with aluminate and borate ions (Golovkina et al., 2009). Below the titration curve plateau for silanol groups (pH < 4.2), the titration curves of  $\text{SiO}_2$  and  $\text{SiO}_2$ -chitosan overlap again, indicating a similar surface charge.

The deposition of chitosan on MCC causes the titration curve to shift to the right of to both the calibration curve and the NR titration curve of the initial MCC sample (Fig. 15), demonstrating the negative charge of the surface, as was observed for the  $\text{Al}_2\text{O}_3$ -chitosan and  $\text{SiO}_2$ -chitosan systems. A peculiarity of this substrate is the recharging of the surface, from positive for MCC to negative for the MCC-chitosan system during chitosan

deposition. In this case, the type of change in the surface charge occurring upon a reduction in pH is the same for both the MCC and the MCC–chitosan systems. Changing symbotically with the curve typical of the initial MCC sample, the titration curve for the MCC–chitosan system almost overlaps the calibration curve in the acidic region ( $\text{pH} \leq 4.5$ ). In this pH range, the potential of the MCC–chitosan system surface is therefore close to zero.

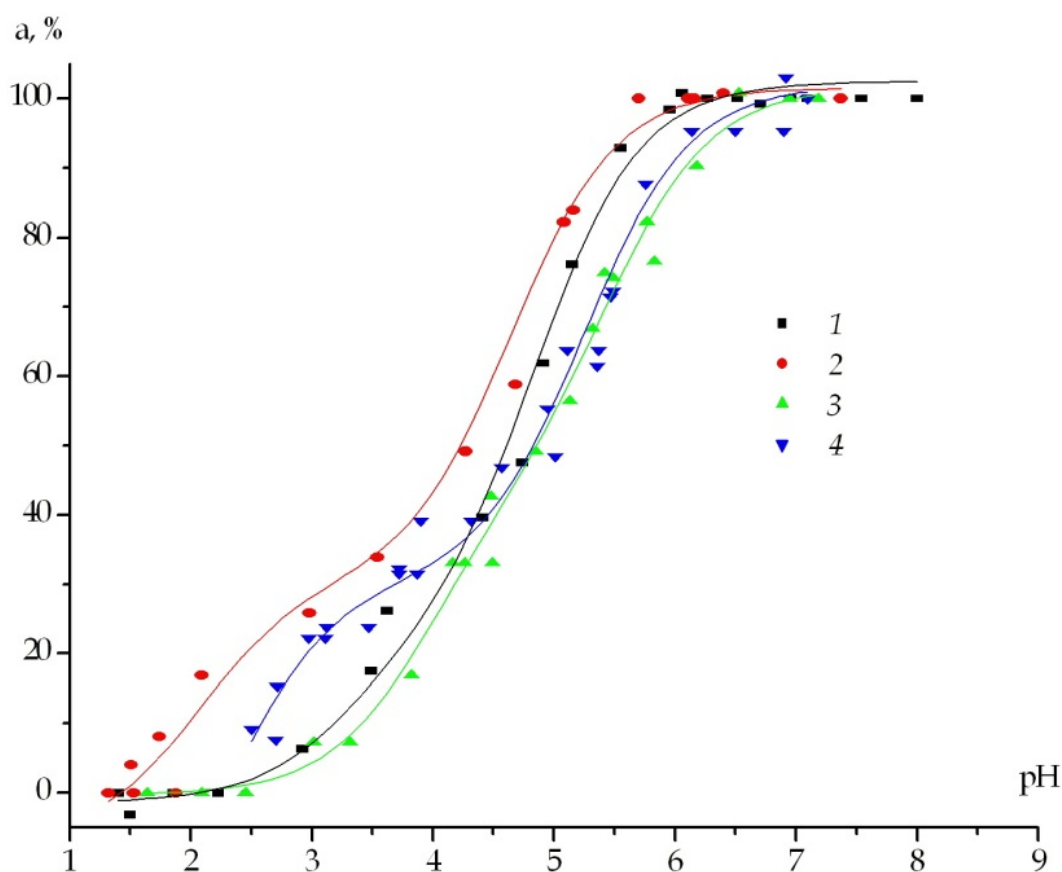


**Figure 14.** Titration curve for NR R1 in aqueous bulk solution ( calibration curve) (1), near the surface of  $\text{SiO}_2$  (2), of the hybrid  $\text{SiO}_2$ –chitosan (3), of the  $\text{SiO}_2$ –chitosan- $\text{Co}^{2+}$  (4) systems.  $a, \% = ((a-\text{anRH}^+)/(\text{anR}-\text{anRH}^+)) \times 100\%$

The deposition of chitosan on the substrate always causes the titration curve of the radical near the surface of hybrid material to shift to the right; i.e., it leads to a negative charge on the surface. While the deposition of chitosan leads to relatively slight changes in the surface potential in the case of inorganic substrates, these changes are so great in the case of MCC that they even lead to changes in the surface charge.

Some of differences in the behavior of titration curves for chitosan-containing materials on the inorganic substrates that occur in the range of high pH values could be due to the close  $\text{pK}_a$  values of the silandiol groups of the  $\text{SiO}_2$  substrate ( $\text{pK}_a = 6.5\text{--}6.8$ ) and of the chitosan ( $\text{pK}_a = 6.42$  (Skorik et al., 2003)). When the pH falls, amino groups of chitosan or silandiol groups become protonated and form hydrogen bonds with one another. As a result, the charge of the initial substrate surface and that of the hybrid material become similar. During further protonation of silandiol and amino groups at the same pH range, this interaction becomes impossible, and the surface charge of the  $\text{SiO}_2$ –chitosan hybrid material becomes more negative than that of the  $\text{SiO}_2$ .





**Figure 15.** Titration curve for NR R1 in aqueous bulk solution ( calibration curve) (1), near the surface of MCC (2), of the hybride MCC– chitosan (3), of the MCC– chitosan- $\text{Co}^{2+}$  (4) systems .  $a, \% = ((a - a_{\text{NRH}^+}) / (a_{\text{NR}} - a_{\text{NRH}^+})) \times 100\%$

This is in complete agreement with the relative positions of the radical titration curves for  $\text{Al}_2\text{O}_3$  and  $\text{Al}_2\text{O}_3$ –chitosan. It must be emphasized that the  $\text{pK}_a$  dissociation value of hydroxyl groups of  $\text{Al}_2\text{O}_3$  is higher, and lies outside the radical sensitivity range.

Deposits of chitosan on the MCC substrate breaks it into smaller particles (Table 2), due most likely to the recharging of the surface from positive to negative. Chitosan molecules are appreciably smaller in size than MCC molecules, that is why a negative charge leads to increased repulsion inside large MCC particles, to their disintegration, and finally to an increase in the total surface area of the material. Chitosan thus plays the role of a disaggregating surfactant. In case of inorganic substrates, the deposition of chitosan leads to an increase in the particle size, and consequently to a decrease in their surface area (Table 1). Chitosan is therefore a weakly aggregating surfactant as to inorganic substrates.

The  $\text{pK}_a$  value for amino groups of chitosan is 6.42 (Skorik et.al, 2003). In the investigated pH range ( $\text{pH} \leq 6$ ), polymer molecules must consequently be positively charged due to the formation of  $-\text{NH}_3^+$  ammonium groups. If chitosan molecules were adsorbed on the surface of inorganic oxides so that the protonated amino groups were on the material surface, this kind of adsorption would be reflected in the titration curves of the modified samples and

would cause the titration curve to shift to the left of the titration curves of the initial oxides. Since this did not happen, we can assume that during the formation of hybrid material, chitosan amino groups form weak hydrogen bonds or Van der Waals bonds with functional groups of substrates (organic or inorganic), leading to an arrangement of chitosan molecules in which glucosamine rings are turned in the direction of the substrate.

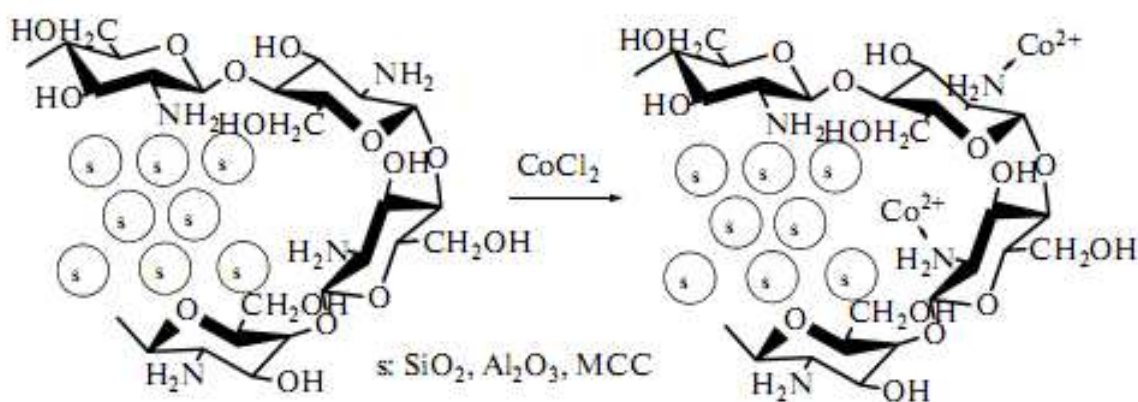
### 3.4.2. $\text{Co}^{2+}$ -containing hybride systems

The presence on the support surface of a polymer that is capable of acquiring metal ions of high coordination numbers is needed to fix these ions more firmly on metal-containing hybride materials. This role was played by chitosan on the surfaces of the inorganic ( $\text{SiO}_2$ ,  $\text{Al}_2\text{O}_3$ ) and organic (MCC) supports.  $\text{Co}^{2+}$  ions were sorbed from an aqueous ethanol solution (Fig. 16).

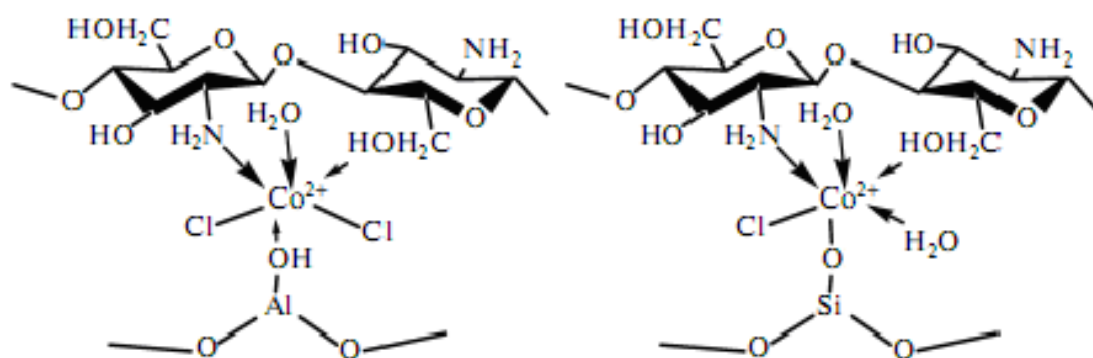
The composition of the studied systems was characterized by elemental analysis. According to the data, the sorption capacity of the hybrid systems relative to  $\text{Co}^{2+}$  ions is different and declines in the following order:  $\text{SiO}_2 > \text{Al}_2\text{O}_3 > \text{MCC}$ . We should note that in this case, there is a difference between the mechanisms of binding ions for different supports, since their surfaces have the same coating influencing the process.

As follows from the elemental analysis, chitosan plays the role of a complexing agent. In this case, all amino groups are involved in the coordination by cobalt ions (ratio  $\text{Co}:\text{NH}_2 = 1:1\text{--}1.5$ ).

In the case of the  $\text{Al}_2\text{O}_3$ -chitosan system, some of the amino groups do not participate in the complexation, but the  $\text{Co} : \text{Cl}$  ratio corresponds to the composition of the initial salt that provides the coordination mechanism for binding cobalt ions (Fig.17). In the case of the  $\text{SiO}_2$ -chitosan system, the  $\text{Co} : \text{Cl}$  ratio is  $\sim 1$ . The need to obey the law of electroneutrality for the obtained hybrid systems requires assuming that the support, in addition to being a weak acid, also plays a coordinating role with respect to  $\text{Co}^{2+}$  ions, which accords with the covalent mechanism of binding  $\text{Co}^{2+}$  ions (Fig. 17). In the case of the MCC-chitosan system, metal ions are sorbed as a basic salt, and hydroxyl groups on the surface provide weaker coordination binding than  $\text{Al}_2\text{O}_3$ .



**Figure 16.** Scheme for synthesis of cobalt-containing hybrid systems. s - support



**Figure 17.** Scheme for the mechanism of binding  $\text{Co}^{2+}$  by hybrid systems

The elemental composition of the surface shows that chitosan does not cover it completely and some part of the functional groups remains on the support surface in free form. The more complex structure of the surface of hybrid system was characterized by investigating its acid–base properties via ESR spectroscopy of the NR used as pH probes. The titration curves of NR R1 on the surface of the  $\text{Co}^{2+}$ -containing system are shown in Figs. 13–15. The shift in these curves relative to the calibration curve left or right allows us to determine whether the surface has a positive or negative charge, respectively (Kovaleva et al., 2000 ; Molochnikov et al., 2007).

The technique for introducing  $\text{Co}^{2+}$  ions into the hybrid materials from the ethanol solutions could lead to the formation of sediments of basic cobalt chloride or chloride–alcoxide micelles on the surface of these materials. In the case of the MCC–chitosan (Fig. 15) and  $\text{Al}_2\text{O}_3$ –chitosan (Fig. 13) systems, the titration curves of the cobalt-containing materials at  $\text{pH} > 5$  are shifted slightly to the left relative to the titration curves of the initial samples. Consequently, the occurrence of basic cobalt chloride on the surface of these materials does not lead to a change in the surface charge, but changes its value slightly. This is due to the effective neutralization of a positive charge of  $\text{Co}^{2+}$  ions by negative chloride and hydroxide ions. In order to confirm our hypothesis, we present the published data on studying the sorption of  $\text{Co}^{2+}$  ions from aqueous solutions and on the nature of the interaction between the metal center and amino groups of chitosan. According to (Minimisawa et al., 1999) maximum adsorption starts to decrease with increasing pH due to the formation of cobalt hydroxocomplexes. The maximum sorption of cobalt ions by chitosan found at pH 6–8 in (Silva et.al, 2008) is in good agreement with the data from (Minimisawa et al., 1999) and is determined by the formation of  $\text{Co}(\text{OH})_2$  phase or slightly soluble basic salts. No chemical interaction with amino groups of chitosan occurs in this case (Zhao et al., 1998).

The  $\text{SiO}_2$ –chitosan system behaves differently. Even at the highest pH values, the titration curve of the sample modified with cobalt (Fig. 17) was appreciably shifted to the left relative to the NR curve of the initial sample, although it remains to the right of the calibration curve. Accordingly, the formation of the chloride hydroxyl cobalt micelles immediately leads to a considerable reduction in the negative surface charge of the  $\text{SiO}_2$ –chitosan system. The different behavior of the titration curve for system III (Table 2) is most likely associated

with the initially lower amount of  $\text{Cl}^-$  ions in the cobalt micelles ( $\text{Co} : \text{Cl} = 1 : 1$ ), i.e., to its higher amount of  $\text{OH}^-$  ions. Acid sites on the  $\text{SiO}_2$  surface (silanol groups) also likely interact with the basic cobalt chloride particles precipitating on the surface, thereby replacing the  $\text{Cl}^-$  ions. In both cases, the titration curve must shift left due to the neutralization of the negative surface charge. In the study of the samples in the aqueous medium at high pH values, the transformation for systems I–III (Table 2) thus occurs; hydration of the surface, followed by hydrolysis leading to the formation of colloidal particles based on chloro-hydroxo complexes of  $\text{Co}^{2+}$  ions, is observed.

These micelles of  $\text{Co}^{2+}$  chloro-hydroxocomplexes begin to dissolve at  $\text{pH} \leq 5$  for all investigated systems. The dependence of sorption of cobalt ions on pH (Minimisawa et al., 1999) shows that the maximum adsorption starts to decline at  $\text{pH} < 5$ , and  $\text{Co}^{2+}$  ions in the solution are in the form of aqua complexes (Zhao et al., 1998). The latter are sorbed by the chitosan primary amino groups, thereby charging the surface positively; in this case, the titration curves of systems I–III (Table 2) (Figs.13-15) are shifted left relative to the corresponding curves for the initial samples and to the calibration curve. A further reduction in pH leads to the neutralization of  $\text{OH}^-$  groups, and no further changes in pH near the particle surface are observed upon a change in  $\text{pH}_{\text{ext}}$ .

This indicates the presence of a horizontal plateau on the titration curve. The value of the horizontal section is lower for system III than for systems I and II (Table 2), since some of the  $\text{OH}^-$  groups are replaced by residues of silicic acid. At high pH values,  $\text{Co}^{2+}$  ions initially cause a substantial decrease in the negative surface charge for the case of system III; their transformation into the form coordinated by the chitosan -  $\text{NH}_2$  groups thus influences the surface charge to a lesser degree upon declining pH. Our results indicate the participation of chitosan amino groups in the complexation with  $\text{Co}^{2+}$  ions. During the interaction of the hybrid materials with the  $\text{Co}^{2+}$ -containing solution, at least a part of glucosamine rings are consequently turned outward and are capable of becoming ligands.

#### 4. Conclusions

pH-sensitive NRs gave reliable information on the local acidity of solutions in and the charge of a surface on pure and metal containing inorganic and organo-inorganic materials and systems and allowed to estimate an electric potential near the surface of  $\text{TiO}_2$  nanoparticles.

The differences between the acidities of external solutions ( $\text{pH}_{\text{ext}}$ ) and inside pores (or near the surface) of all the studied materials and systems ( $\text{pH}_{\text{int}}$ ) were found.

The method of spin pH probes allowed to determine the ionization constants of characteristic functional groups of  $\text{SiO}_2$ -based systems from the horizontal plateaus corresponding to the constant  $\text{pH}_{\text{int}}$  in the samples.

An increase in concentration of  $\text{H}^+$  ions (a decrease in  $\text{pH}_{\text{int}}$ ) in solutions located inside  $\alpha\text{-Al}_2\text{O}_3$ ,  $\text{TiO}_2$  hydrogel and near the surface of the BS-50 type  $\text{SiO}_2$ ,  $\text{TiO}_2$  and  $\text{SiO}_2$  xerogels ; the

related CMs and hybrid materials; metal-containing systems, as compared to those of external solution can be explained by releasing  $H^+$  ions due to dissociation of acidic functional groups, exchange them with metal ions and the partial desruption of hydrogen bonds. It leads to negative charge of a surface of the above-mentioned objects. A decrease in concentration of  $H^+$  ions (an increase in  $pH_{int}$ ) as compared to those of external solution were characteristic for  $\gamma-Al_2O_3$  and cellulose matrixes. This resulted from binding  $H^+$ -ions by the surface of  $\gamma-Al_2O_3$  and MCC and PC with basic functional groups such as  $-AlOH$ ,  $-AlO^-$  and  $OH^-$ , respectively. As a result, a surface gains a positive charge.

The sorption capacity of  $Cu^{2+}$  ions depends on a surface charge of the oxides gels, xerogels and the related CMs studied and decreases as a negative surface charge reduces. The sorption of  $Cu^{2+}$  ions on the surface of nanoparticles of nanostructural  $TiO_2$  increases the charge of the latter. An increase in a percentage of PC in the  $SiO_2$ -PC composites leads to an increase in the amount of silanol groups as a result of increasing in dispersivity of  $SiO_2$  particles and specific surface ( $S_{sp}$ ) of the samples, and to reducing a negative surface charge up to zero, and even its reversing. It led to the formation of  $Cu(OH)_2$ .

The deposition of chitosan on the substrate always creates a negative charge on the surface. While the deposition of chitosan leads to relatively slight changes in the surface potential in the case of inorganic substrates such as  $Al_2O_3$  and  $SiO_2$ , these changes are so great in the case of MCC that they even lead to changes in the surface charge.

The charge of the surface of  $Co^{2+}$ -modified organo-inorganic hybrid materials at different  $pH_{int}$  was found to effect on the composition and structure of  $Co^{2+}$ -containing surface compounds.

The modification of a surface of powder cellulose with nanostructured  $SiO_2$  and  $TiO_2$  xerogels, aluminum oxides, silica and MCC with acidic functional groups and chitosan makes it possible to adjust the local acidity and surface charge over a wide range.

The study of the surface of organo-inorganic composites and hybrid materials and systems using pH-sensitive nitroxyl radicals allows also to reveal regularities in changing their properties during further modification. In addition, this method enables us to describe qualitatively the processes of structure formation in these systems and their effect on catalytic activity in different pH-dependent reactions.

The calculated  $\varphi$  value (31.7 mV) was found not to be the electric potential of  $TiO_2$  nanoparticle surface, but it only characterizes the electric field generated by a nanoparticle at the site where the radical fragment  $-N-O\bullet$  of NR is located. Once the anisotropic spectra of NR in nanostructured oxides are simulated, an electical potential of a surface can be determined. When measuring the SEP of solids, the knowledge of the distance between a radical and a surface is of principal importance. The fixation of pH-sensitive NRs on the surface of nanoparticles with linkers of a known length can solve this problem. This will allow one to calculate the potential immediately on the surface of nanoparticles and to compare the calculation results with the experimental data on the electrokinetic potentials.



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## Acknowledgement

The authors are very grateful to :

- Prof. A.M. Volodin and Prof. A.I. Kulak for synthesizing and supplying aluminum oxides and nanostructured TiO<sub>2</sub> both original and modified with F<sup>-</sup> and SO<sub>4</sub><sup>2-</sup> groups,
- Dr. I.A. Kirilyuk and Prof. I.A. Grigor'ev for synthesis of pH-sensitive nitroxide radicals,
- Dr. A.B. Shishmakov, Mrs. Yu. V. Mikushina and Dr. E. V. Parshina for synthesizing TiO<sub>2</sub> and SiO<sub>2</sub> hydrogels and xerogels, pure and Cu<sup>2+</sup>-containing solid-phase composites based on nanostructured SiO<sub>2</sub> and TiO<sub>2</sub> and powder cellulose,
- Mr. A.V. Mechaev for synthesis and characterization of pure and Co<sup>2+</sup>-containing hybrid organo-inorganic materials based on the chitosan-SiO<sub>2</sub>, chitosan- Al<sub>2</sub>O<sub>3</sub>, and -chitosan-cellulose systems,
- Prof. A.I. Kokorin , Dr. A.V. Pestov and Prof. Yu.G. Yatluk for fruitful discussion of the study results.

The studies presented in this manuscript were financially supported by the Ministry of the Education of the Russian Federation (2007-2011, Theme 01.2.007-06425).

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