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Metastable Solvent Epitaxy of SiC, the Other Diamond Synthetics

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1. Introduction

Professor W. Shockley, a Nobel prize winner and the ‘father’ of the transistor, predicted in the 1950s that SiC would soon replace Si in devices because of its superior material properties (Cited from Choyke, 1960). His prediction, however, has not yet come true because of the high cost of manufacturing SiC wafers. The current price of 2-inch SiC wafers is close to that of diamond jewellery rather than to that of Si wafers. Nevertheless, owing to its promising physical properties for high-power devices and for achieving significant reductions in CO₂ emission, many researchers and companies are purchasing many SiC wafers (Madar, 2004; Nakamura et al., 2004). Very recently, the present authors reported a novel process for fabricating epitaxial SiC (Nishitani & Kaneko, 2008). This process has the potential to cost less and to provide SiC wafers with high crystalline quality. Here we present a new synthesis process for manufacturing SiC. In this process, the driving force for crystal growth is elucidated by considering a similarity—the coexistence of the stable and metastable phases—with the diamond synthesis process reported by Bonenkirk et al. (1959) of General Electric.

In this review, the authors introduce the experimental procedures and the results of this novel process of making stable phase of SiC at first. Then they will show the speculated mechanism from the thermodynamical point of view, especially the concept of the double phase diagram and the concentration profile. In the end, understanding the phase stability of many polytypes of SiC, the reported phase diagrams, the difficulty of the equilibrium state achievement, and the results of the first principles calculations will be discussed.

2. Experimental results

2.1 Conventional methods

While SiC was first synthesized by Acheson at the end of the 19th century (Acheson, 1896), single-crystal wafers of SiC became commercially available only in the early 1990s. Boules of other important semiconductors such as Si, GaAs and InP are manufactured through a liquid process, while those of SiC are produced through a vapor process. Owing to the

peritectic reaction of SiC at above 2500 °C as shown in the phase diagram of Fig. 1, growth from a stoichiometric melt is not possible at normal pressure. The incongruent compounds are usually produced by the solution method. The low carbon solubility in liquid Si leads to insufficient mass transport of C, and this in turn results in the crystal growth in the Si-C solution method being quite slow (Hofmann & Müller, 1999). The solution method provides crystals with better crystallinity and a smaller defect density as compared to vapor methods. Thus, the solution method is preferred for preparing SiC (Hofmann & Müller, 1999). Many attempts have been made to increase the carbon solubility in liquid Si, especially by the addition of metallic elements or the application of high pressures. Nevertheless, the productivity of such methods does not exceed that of the vapor methods (Casady & Johnson, 1996; Chaussende et al., 2007; Hofmann & Müller, 1999; Wesch, 1996).

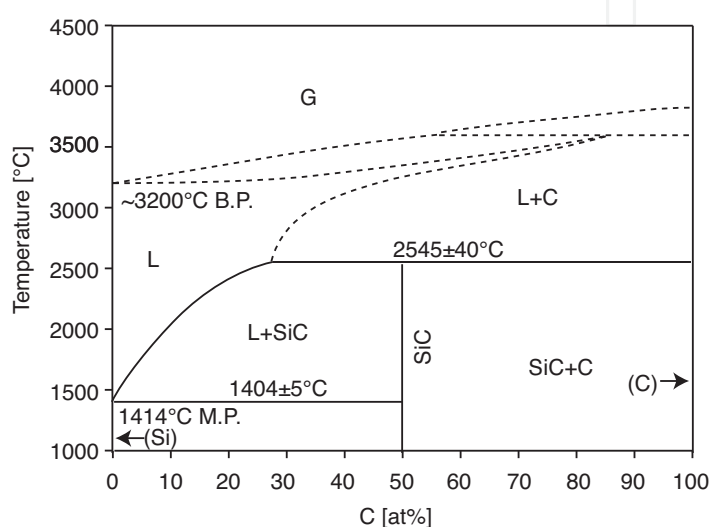


Fig. 1. Phase diagram of Si-C binary system (Olesinski & Abbaschian, 1996).

Among the vapor processes, the sublimation process, which is schematically illustrated in Fig. 2a and which was first proposed Tairov & Tsvetkov (1978), is the standard process used for the commercial manufacture of SiC. In this process, the source powder and a seed crystal are placed in a graphite crucible. SiC powder is then heated to typically 2400 °C and allowed to sublime. Si and C vapors are transferred in an inert gas, and they recrystallize on a slightly cooler single-crystal seed. The temperature dependency of the crucible is schematically shown in the left-hand panel of Fig. 2a. The main drawback of this method is that the obtained boules contain many types of grown-in defects, such as micropipes, step bunching, dislocations and stacking faults. For the reduction of these defects, the operation parameters such as the growth temperature, gas flow rate and total pressure should be controlled quite steadily for a long operation time. This long and sensitive operation performed at a high temperature and involving a temperature gradient makes SiC wafers expensive. Researches on alternative processes such as chemical vapor deposition and the vapor-liquid-solid process are continuing with the objective of improving the growth rate and crystal quality (Chaussende et al., 2007).

2.2 Metastable solvent epitaxial method

Our proposed method employs the solution growth technique. A schematic drawing of the apparatus is presented in Fig. 2b. In this method, C is transferred through liquid Si instead of Ar gas, which is used in the conventional method. Poly-crystalline 3C-SiC source and single-crystalline 4H-SiC plate are placed in a TaC crucible and liquid Si solvent is sandwiched

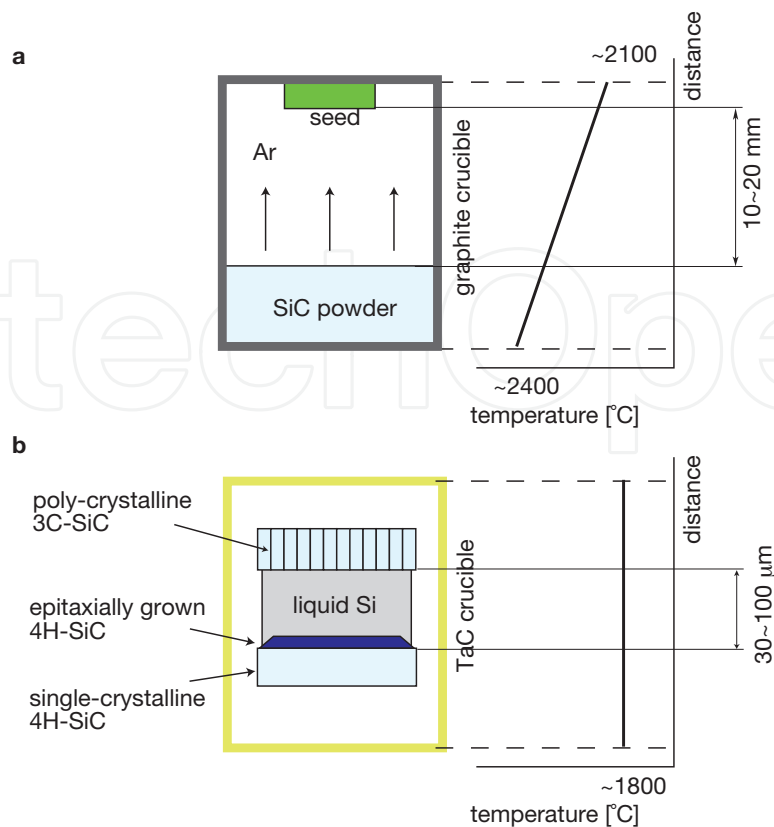


Fig. 2. Schematic illustrations of **a** the conventional sublimation method and **b** the new solvent method for the preparation of SiC. Their setups are similar and comprise the sources (SiC powder and polycrystalline 3C-SiC), transfer media (Ar and liquid Si), produced crystals (the seed and 4H-SiC fine particles) and crucibles (graphite and TaC). The characteristic differences are in the thicknesses and the temperature profiles of the transfer media, as illustrated in the panels on the right-hand side.

between them. The solvent is a very thin layer with a thickness of 30~100 μm . The temperature inside the crucible is held constant.

After holding at the growth temperature of 1800°C for 10 min, the newly grown 4H-SiC layer is observed on the original 4H-SiC single-crystalline plate as shown in Fig. 3 (Nishitani & Kaneko, 2008). The scanning electron microscopy image shows the transverse section of a sample with the sandwiched structure of SiC and Si layers. Because the substrate of 4H-SiC is n-doped, the newly grown 4H-SiC layer is easily distinguished from the original substrate. The feed of polycrystalline 3C-SiC was dense before the operation. Its boundaries dissolve preferentially, and it becomes porous during the operation. The wave dispersion X-ray (WDX) profile of carbon in Fig. 3 shows a concentration of 50% in the SiC layers and almost zero in the Si solvent. Upon reversing the configuration of 3C- and 4H-SiC plates, 4H-SiC grew epitaxially again, which indicates that no unintentional temperature gradient occurs in the crucible; thus, the temperature gradient cannot be the driving force of the solvent movement.

In the case of replacing the single-crystalline 4H-SiC to poly-crystalline 3C-SiC in the configuration of Fig. 2b, many fine particles with diameters of 30~300 μm are observed to be produced on the both sides of polycrystalline 3C-SiC source plates, as shown in Fig. 4a. Multifaceted crystals grow like mushrooms and no grain boundaries are observed in them.

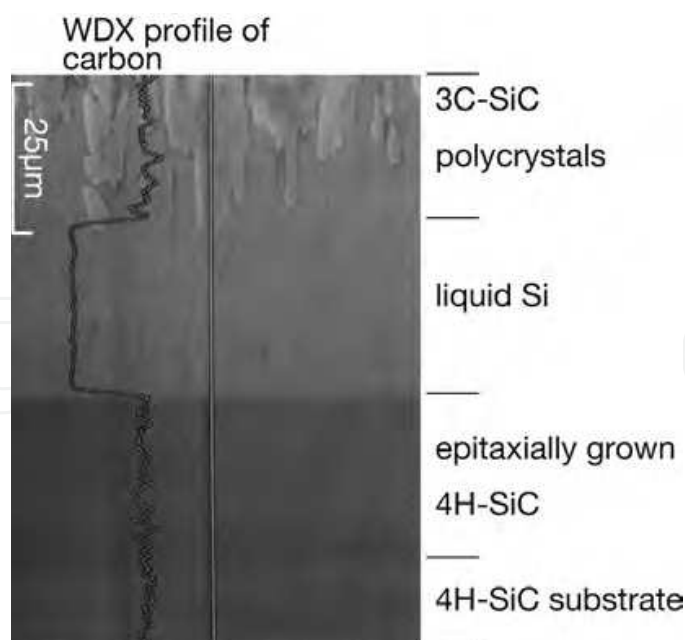


Fig. 3. Scanning electron microscopy image and wave dispersion X-ray (WDX) profile of carbon of epitaxially grown 4H-SiC in the transverse section of a sample with the sandwiched structure of SiC and Si layers(Nishitani & Kaneko, 2008).

Figure 4c shows the Raman spectra of the fine crystalline particles and the polycrystalline 3C-SiC substrate obtained using a 488-nm excitation laser at room temperature. The spectra of the substrate show broad peaks at around 768 cm^{-1} and 797 cm^{-1} , which are typical of 3C-SiC with randomly distributed stacking faults(Rohmfeld et al., 1998-I). On the other hand, the fine particles show a sharp peak at 776 cm^{-1} , which is typical of 4H-SiC. Thus, we conclude that fine single-crystal 4H-SiC particles are crystallized from polycrystalline 3C-SiC in the absence of a temperature gradient. 4H-SiC crystals grow without a temperature gradient nor a concentration difference between sources and products of SiC.

3. Speculated mechanism

3.1 Solution method

The origin of the driving force for crystal growth is the same as that in the case of diamond synthesis from a solvent of a metal-carbon binary system. Both the diamond synthesis method and the proposed method for SiC synthesis are solution growth methods; a typical example of a solution growth method is the method used for the growth of alum crystals, which is very familiar as a topic frequently dealt with in science classes, even in primary schools. Thus, a review of the growth process of alum crystals would be a good starting point for explaining as well as understanding the growth of diamond and SiC crystals. Large alum single crystals can be obtained as follows. First, alum powder is dissolved in hot water. Next, a seed tied to a long thread is inserted into the solution, and the solution is cooled and left to stand. A crystal starts growing from the seed and develops into a large single crystal.

Figure 5a shows schematic representations of the alum deposition process. The upper panel represents the dissolution of alum in water at a high temperature and the lower panel represents the deposition of alum on the seed at a low temperature. In the system shown in Fig. 5b, the source and seed are sandwiching the water solution; a temperature gradient is applied in the water solution. This is the setup in the so-called solvent zone technique,

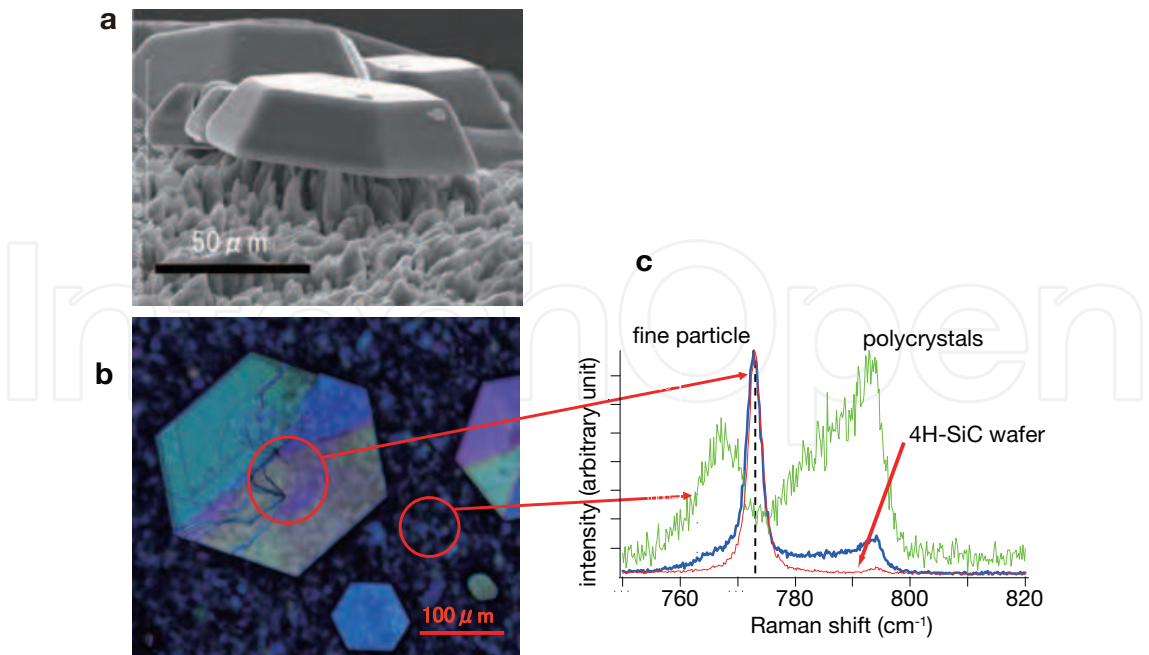


Fig. 4. **a** Side view and **b** top view of fine particles on the polycrystalline 3C-SiC source plate and **c** Raman spectra of the regions in **b** indicated by the circles. **a** is a scanning electron microscopy image and **b** is an optical microscopy image. The spectrum of the single-crystal 4H-SiC wafer is also shown in **c** as a standard.

where dissolution and deposition occur simultaneously. The driving force for crystal growth is supersaturation, which can be schematically explained by considering the solubility limit of alum or the so-called phase diagram of an alum-water system.

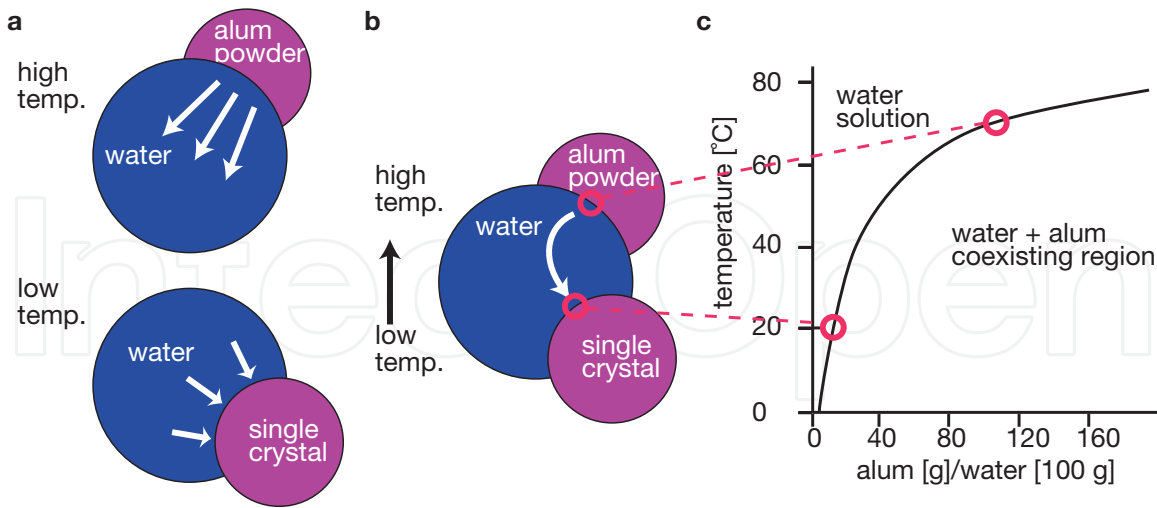


Fig. 5. Schematic drawings of **a** alum deposition and **b** the solvent zone technique in which a temperature gradient is present; **c** the phase diagram of the potassium alum-water system. In **a** and **b**, the vertical positions of the alum powder (source), water (solvent) and single crystal (seed) represent their relative energies or chemical potentials.

The solubility of potassium alum in water is measured as shown in Fig. 5c. The solubility limit, or liquidus of alum, separates the water solution region from the region where water and alum

coexist. As the temperature increases, the solubility limit shifts to higher alum contents. At 80 °C, 321 g of potassium alum dissolves in 100 g of water. Such a solution can be cooled to a temperature where the solubility limit is considerably low, and it is called a supersaturated solution. At lower temperatures around the seed crystal, the small solubility limit facilitates the deposition of alum on the seed crystal. As marked in Fig. 5c, the temperature gradient between the source and the seed alum results in different solubility limits around them; this difference is the driving force for the dissolution and deposition of alum.

3.2 Double phase diagram

Although diamond synthesis in a metal-carbon system at high pressures is slightly complex, it is based on the same mechanism as that of alum deposition. The initial materials used for diamond synthesis in the original General Electric process were graphite and group VIII transition metals (Bonenkirk et al., 1959). At a high pressure and a high temperature, metal melts and comes into contact with graphite. The liquid metal was initially thought to function as a catalyst by the workers at General Electric, but was later clarified to act as a transport medium by Giardini and Tydings (Giardini & Tydings, 1962). Similar to Fig. 5b, a schematic drawing of a diamond synthesis system is shown in Fig. 6a. The liquid metal that is commonly used in diamond synthesis is Ni; the Ni-C phase diagram at 54 Kbar is shown in Fig. 6b (Strong & Hanneman, 1967). Although the graphite phase is the only stable phase at normal pressure, at 54 Kbar, the diamond phase is stable below 1740 K. Liquid Ni coexists with the diamond phase between 1667 K and 1728 K. Thus, the solution growth of diamond from liquid Ni is possible in this temperature range. The true trick in diamond synthesis is not the 'stability' of the diamond phase, but the 'metastability' of the graphite phase, as indicated by the dashed lines in Fig. 6b. The synthetic capsule in high-pressure anvils is small and it is very difficult to manage sensitive temperature gradients. Thus, each process of crystal growth, i.e. the melting of metastable graphite in liquid Ni, carbon transfer within the solvent and crystallization of stable diamond on the seed, occurs in a very small temperature range. The driving forces for these reactions cannot be obtained from the temperature gradient, but can be obtained from the difference in the stability between graphite and diamond. For example in Fig. 6b, at 1700 K, the solubility of metastable graphite in liquid Ni represented by the dashed line is higher than that of stable diamond denoted by the solid line. The driving force for the crystal growth of the alum shown in Fig. 5 is generated when there is a spatial or temporal temperature difference, but that in diamond synthesis can be produced even at a constant temperature.

Figure 6b shows the so-called double phase diagram that is well known in metallurgy, e.g. the Fe-C system (Hansen & Anderko, 1958), and that is commonly observed for the systems containing metastable phases (Ishihara et al., 1984-5). The mechanism of crystal growth in the presence of a solubility difference between the stable and metastable phases, which is discussed in the present study, can be observed in the work of Van Lent (1961) on the transformation of mercury/white tin amalgams to gray tin, and it was first elucidated by Hurle et al. (1967) using the so-called thin alloy zone (TAZ) crystallization.

3.3 Ostwald ripening and concentration profile

The solution growth is easily understood by the phase diagrams, but the simultaneous growth and dissolution processes of polytypes are hardly recognized. Such a puzzling process can be observed in the Ostwald ripening, which is shown schematically in Figs. 7. Ostwald ripening, also called coarsening, which occurs at the late stage of the precipitation,

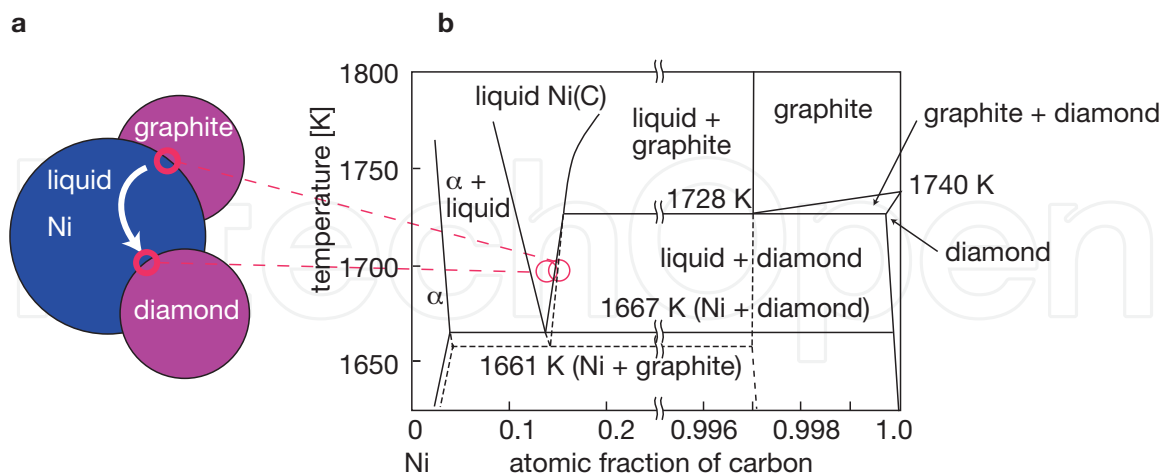


Fig. 6. Schematic drawings of **a** diamond synthesis and **b** enlarged sections of the Ni-C phase diagram at 54 Kbar. In **a**, graphite, liquid Ni and diamond are the source, solvent and seed, respectively. In **b**, the stable Ni-diamond and metastable Ni-graphite reactions are represented by solid and dashed lines, respectively (Strong & Hanneman, 1967).

is driven spontaneously by the diffusion between small (S) and large (L) precipitates due to Gibbs-Thomson effect. In this section, the matrix and precipitate phases are represented as α and β , respectively. The precipitates show the same composition of x_β but different pressures of small and large radii, which makes the different free energies, G_β^S and G_β^L , as shown in free-energy vs. concentration diagram of Fig. 7a. The common tangents of free-energy curves between matrix solution and precipitates show different angles and touch at the different concentrations x_α^L and x_α^S of the free energy curve of the matrix solution. This concentration difference of matrix appears at the interfaces of the precipitates. Fig. 7b shows the schematic diagram of the solute concentration profile of the system of α matrix and small (S) and large (L) β precipitates. The matrix concentration equilibrating with the small precipitates should be higher than that with the large precipitates. This difference drives the solute diffusion and thus the simultaneous growth and dissolution of precipitates.

3.4 Direct melting of metastable phases

Ostwald ripening is the reaction in a solid solutions, which means the life time of metastable phase may be longer at lower temperatures and lower diffusivity. Does such a metastable phase directly melt in liquid?

The typical example of the behavior of the coexistence of stable and metastable phases is observed in the Fe-C system. As mentioned before, the double phase diagram of Fe-C and Fe-Fe₃C systems is well studied and established. The melting behavior of metastable Fe₃C phase has investigated in detail by Okada et al. (1981). They measured the differential thermal analysis (DTA) curves for the white, gray and mixture cast irons at the eutectic temperature and composition region. Fig. 8 shows the summarized results of DTA curves as well as the schematic double phase diagram. The endothermic temperatures shift due to the kinetic reason of the measuring apparatus, but the corrected temperatures show the stable and metastable eutectic temperatures of 1426K and 1416K, respectively. The specimens of gray cast iron contains stable phases of graphite and fcc-Fe(austenite), where they all melt only

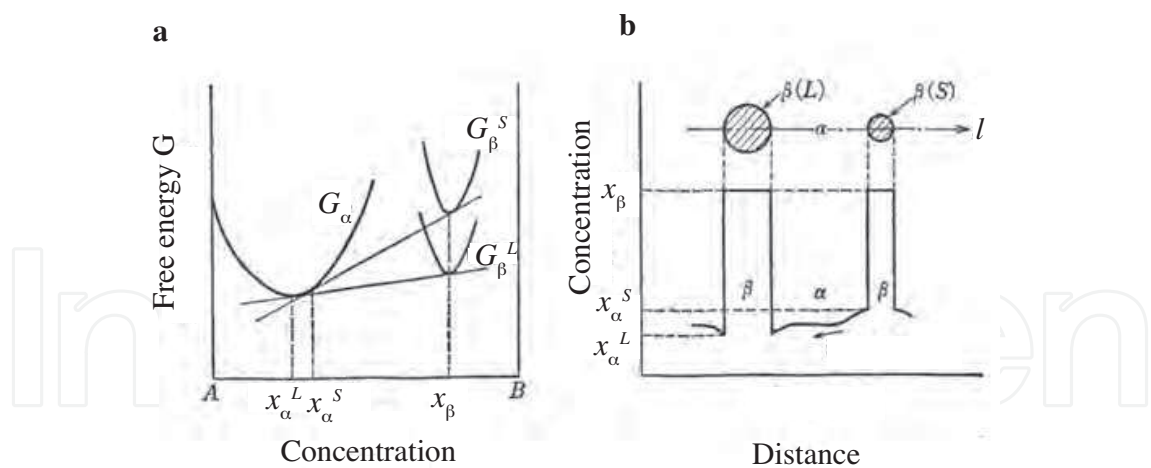


Fig. 7. Schematic drawings of **a** free energy vs. concentration diagram and **b** concentration profile of the Ostwald ripening.

at the stable eutectic temperature of 1426K. On the other hand, the specimens of white cast iron shows the double peaks of endothermic reactions at the slow heating rates. Okada et al. (1981) found that at the first peak the metastable Fe_3C melts but soon graphite solidifies and then remelt at the second peak. At the faster heating rates, only the melting of the metastable Fe_3C phase occurs. For the specimens of mixture cast iron, the reactions are complicated but the melting and solidifying occur simultaneously. These experimental results indicate that the metastable phase is so stable that can melt directly.

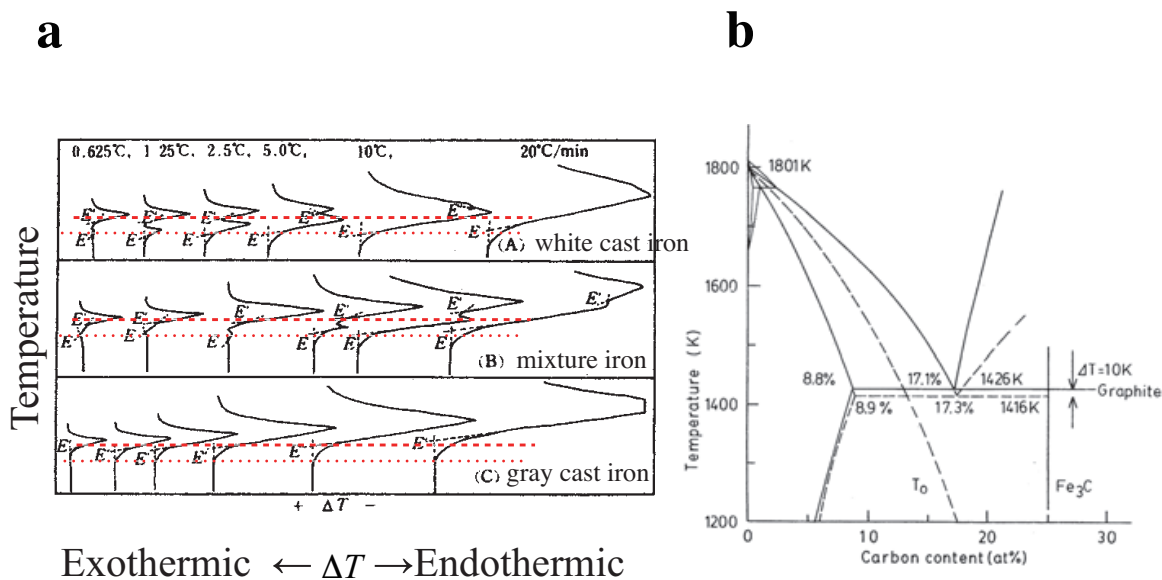


Fig. 8. **a** DTA curves of cast irons(Okada et al., 1981) and **b** the double phase diagram of equilibrium Fe-Graphite and metastable Fe- Fe_3C systems.

3.5 Speculated mechanism

From the experimental result shown in Fig. 4, 4H-SiC is expected to be more stable than 3C-SiC. The Si-C system should show a double phase diagram, as schematically shown in Fig. 9a. The corresponding free-energy vs. concentration diagram is also illustrated in Fig. 9b.

The solubility limit of each phase is determined by the tangent common to the free-energy curves of the coexisting phases. At the temperature indicated by the dotted line, the solubility limit of metastable 3C-SiC is x_1^{3C} , which corresponds to the dashed line of the liquidus in Fig. 9a. The solubility limit of stable 4H-SiC is x_1^{4H} , which corresponds to the solid line of the liquidus. The concentration gradient in the layers consisting of 3C-SiC and 4H-SiC and a very thin layer of liquid Si in between is obtained. The schematic carbon concentration profile in the liquid Si layer is shown in Fig. 9c. The concentration gradient at the metastable solubility limit of 3C-SiC leads to the extraction of C from the source plates. This concentration gradient also causes carbon diffusion across the liquid Si solvent and to the interface of the seed, where C is deposited due to the supersaturation of 4H-SiC. Although the small solubility limit of C in liquid Si, which is the cause of the slow growth of SiC in the conventional liquid method, still remains, the small thickness of the Si solvent in the new method leads to a sufficient concentration gradient for the growth of 4H-SiC crystals.

4. Phase stability of SiC polytypes

4.1 Phase diagram assessment of Si-C system

The key data in order to rationalize this novel process is the phase stability or the hierarchy of SiC polytypes. The reported phase diagrams, however, are somewhat conflicting.

The standard data book Chase (1998) shows the standard formation enthalpies for α and β phases, as follows:

$$\Delta_f H(\alpha - \text{SiC}, 298.15\text{K}) = -71.546 \pm 6.3\text{kJ/mol}$$

$$\Delta_f H(\beta - \text{SiC}, 298.15\text{K}) = -73.220 \pm 6.3\text{kJ/mol}.$$

Although the data shows that the β (cubic) phase is more stable than the α (hexagonal) phase at 298.15K, the difference of the measured values are within the measurement errors. The α (hexagonal) phase indicated in Chase (1998) is 6H, but also mentioned that the many polytypes have not been adequately differentiated thermodynamically. The heat capacity and Gibbs free energy are also reported as shown in Fig. 10. The measured values and the adapted functions in Chase (1998) suggest that α (hexagonal) phase is less stable up to 2000K, and they concluded unlikely the transformation to β (cubic) phase at about 2300K.

The most widely adapted phase diagram should be that by Olesinski & Abbaschian (1996) as shown in Fig. 1, where the β (cubic) phase is more stable than the α (hexagonal) phase at any temperatures below the periodic temperature of the decomposition of SiC, 2545°C. Although the evaluators of Olesinski & Abbaschian (1996) mentioned nothing on the types of α (hexagonal) phase, the same authors reported the co-existence of polytypes of α phases, 6H, 15R, and 4H(Olesinski & Abbaschian, 1984). Furthermore, it also mentioned on the report of Verma & Krishna (1966), the existence of α stability above 2000°C. On the other hand, Fromm & Gebhardt (1976) reported the different type of phase diagram as shown in Fig. 11, where the phase transition from β to α phases occurs at around 2000°C.

Solubilities of carbon in liquid silicon measured by Hall (1958), Scace & Slack (1959), Dash (1958), Dolloff (1960), Nozaki et al. (1970), Oden & McCune (1987), Suhara et al. (1989), Kleykamp & Schumacher (1993), Iguchi & Narushima (1993), Ottem (1993), and Yanabe et al. (1997) are summarized as in Figs. 12. Two reported phase diagrams as shown in Fig. 1 and Fig. 11 are based on the data given by Dolloff (1960). Dolloff (1960)'s data, however, are distinctively different from the others, where the solubility limits are larger than the others.

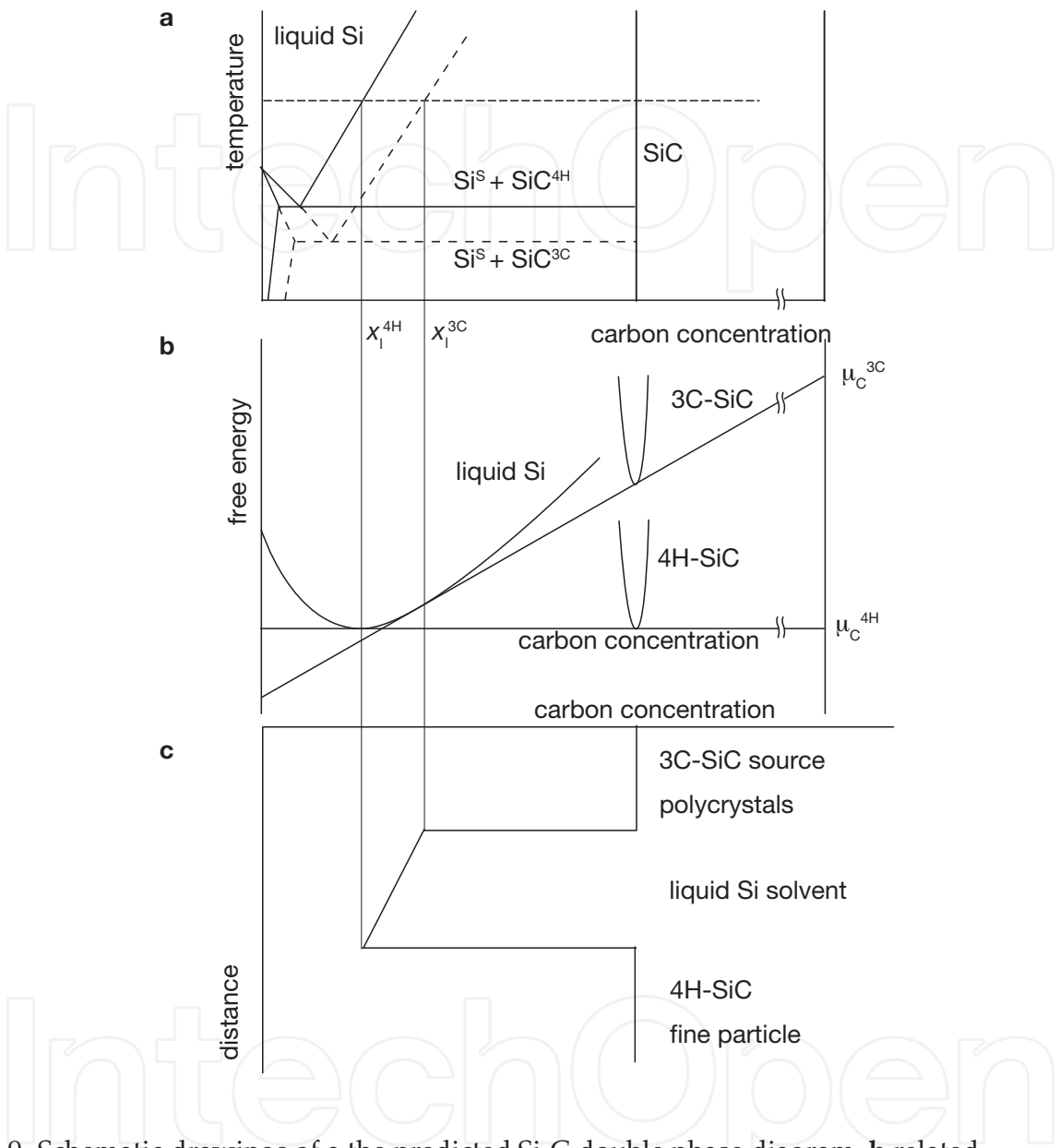


Fig. 9. Schematic drawings of **a** the predicted Si-C double phase diagram, **b** related free-energy vs. concentration diagram and **c** carbon concentration profile in the liquid Si solvent between the 3C-SiC source and the 4H-SiC fine particles. The metastable eutectic temperature of the reaction liquid Si \rightarrow Si^S + SiC^{3C} is lower than the stable eutectic temperature of the reaction liquid Si \rightarrow Si^S + SiC^{4H}, where Si^S denotes solid Si. The chemical potentials of C, μ_c , are given by the intersections of the common tangents with the pure-carbon line in **b**, and are spatially different in the liquid Si solvent contacting with 3C-SiC and 4H-SiC in **c**. The configuration of **c** is related to that of the panel on the left-hand side in Fig. 2b.

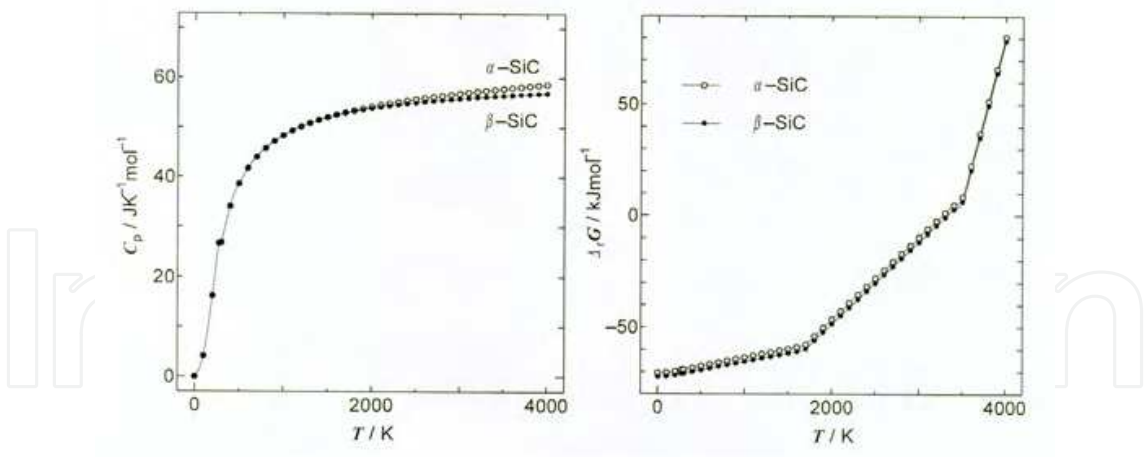


Fig. 10. Heat capacity and Gibbs energy of SiC(Chase, 1998).

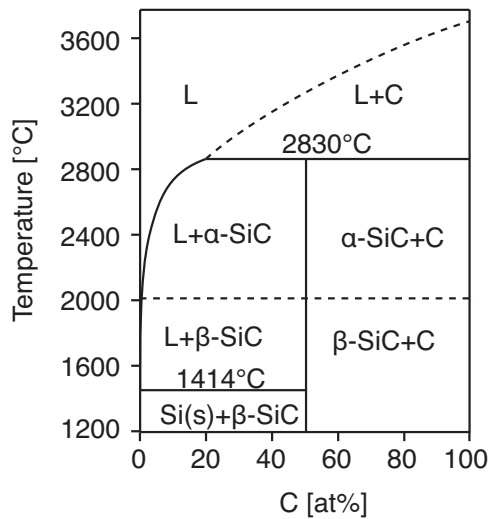


Fig. 11. Phase diagram of Si-C binary system including the phase transition from β to α phase around 2000°C(Fromm & Gebhardt, 1976).

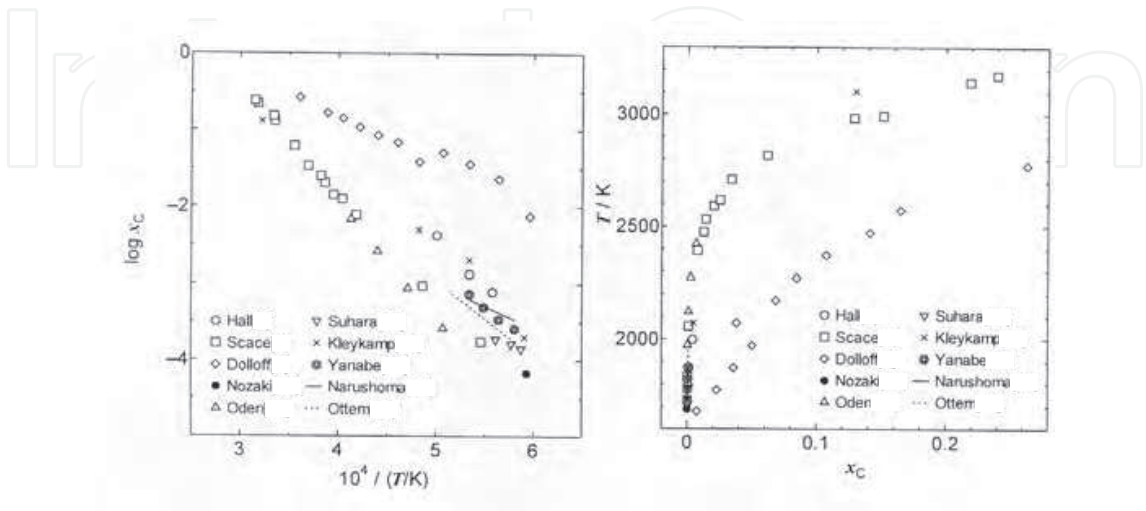


Fig. 12. Solubility of carbon in liquid silicon.

The reported phase stability between α (hexagonal) and β (cubic) might be 6H and 3C. If this assumption is true, 4H stability has not been shown experimentally. Furthermore, the distinct difference of solubility limits indicates that the coexistence of stable and metastable phases.

4.2 Difficulty of the equilibrium state

Although the conflicts of phase diagrams shown above remain, there have been many attempts of experimental and theoretical researches on the kinetic process of crystal growth of SiC polytypes. Famous stability diagrams of SiC polytypes proposed by Knippenberg (1963) and Inomata et al. (1968) show that the crystalized phases are controlled both by the temperature and growth rate of the operations.

The limit to slow growth rate of kinetic processes or results of long period holding should be equal to static results. But it is very difficult to dissolve whole amount of SiC crystals due to small solubility limit of SiC in liquid Si. If there remain seeds of metastable phases during the previous processes, it is difficult to remove all of them. The metastable phase also grows due to the co-existence of less stable phases as shown in Ostwald ripening, or from super-saturated liquid Si. Furthermore, the required high temperature and inert environment make the static conditions very difficult.

Inomata et al. (1969) performed careful experimental observations on the relationship between the polytypes of SiC and the supersaturation of the solution at 1800°C with the solution method, and have shown the following results;

1. β -SiC crystallizes from highly supersaturated solution. The crystals obtained at the condition of low supersaturation, however, consist of mainly α -SiC such as 4H, 15R and 6H.
2. Relative amount of 4H increased markedly with decreasing the supersaturation.
3. From the results stated above, it is concluded that 4H is the most stable structure at 1800°C among the basic polytypes of SiC, 3C, 4H, 15R and 6H.

Those results indicate that the difference between 4H and 6H is crucial for determining the hierarchy of SiC polytypes.

Izhevskiy et al. (2000) summarized not only the kinetic observations, but also pointed out the impurity effects, especially nitrogen affects the transformations among 6H, 3C and 4H SiCs. Not only through the contamination of the higher temperature operations, but also from the starting materials made by Acheson method, specimens contain non-negligible nitrogen. Very recent improvements on materials and apparatuses make it possible to avoid nitrogen inclusions and get the hierarchy of pure SiC polytypes experimentally soon.

4.3 First principles calculations

For some cases of hardly measuring experimental value, the first principles calculations give some hints of the puzzles, and have been applied on the topic of the hierarchy of SiC polytypes. Liu & Ni (2005) have summarized the results of the first principles calculations of SiC. All calculations show that β -SiC is less stable than α -SiCs of 2H, 4H and 6H. The hierarchy between 4H and 6H is subtle; two of nine calculations shows that 6H SiC is most stable, but the majority of the results indicates that 4H SiC is most stable. Of course the calculating results should be judged by the precisions, the energy differences, however, are too small from 0.2 to 2 meV/Si-C pair to identify. Although the reliability of the first principles calculations of the hierarchy of SiC polytypes are insufficient, it is important that the calculating results show against the experimental results. Those are ground state results, which means that is

only reliable at low temperatures. For including the finite temperature effect, the vibrational entropy effect is crucial, because the configurational entropies should be similar among these SiC polytypes due to the similar local configurations of tetrahedrons. One calculating result including the vibrational effect is shown by Nishitani et al. (2009) as in Fig. 13.

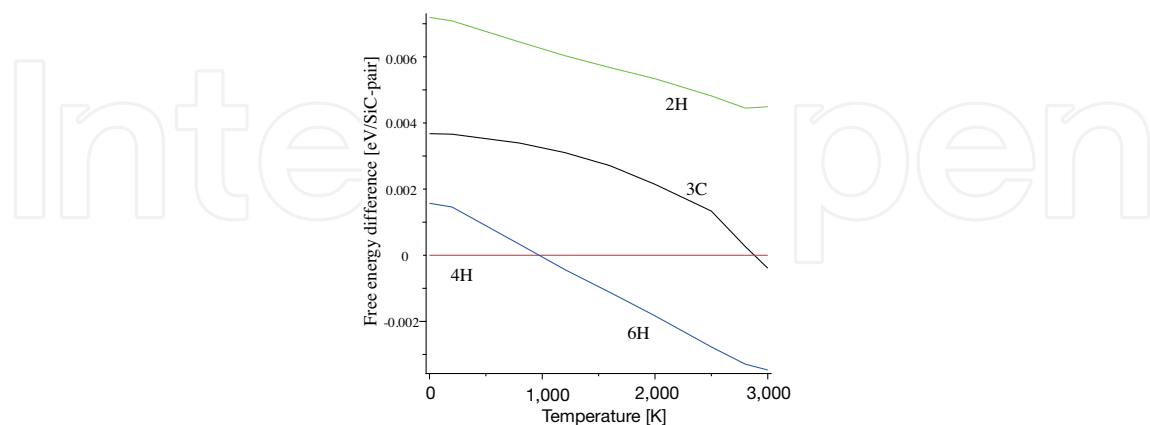


Fig. 13. First principles calculations of temperature dependency of free energy difference of 6H, 3C, and 2H against 4H SiC (Nishitani et al., 2009). Finite temperature effects are included through the vibrational free energy calculated by Phonon codes (Medea-phonon, n.d.; Parlinski et al., 1997).

These first principles calculations were carried out using the Vienna Ab initio Simulation Package (VASP) code (Kresse & Furthmüller, 1996a;b; Kresse & Hafner, 1993; 1994). The interaction between the ions and valence electrons was described by a projector augmented-wave (PAW) method (Kresse & Joubert, 1999). A plane-wave basis set with a cutoff of 400 eV was used. The exchange-correlation functional was described by the generalized gradient approximation (GGA) of the Perdew-Wang91 form (Perdew & Wang, 1992). Phonon calculation was performed by a commercial pre-processor of Medea-phonon (Medea-phonon, n.d.) with the direct method developed by Parlinski et al. (1997). The volumes and/or c/a ratios were fitted to the most stable point at each temperature.

Fig. 13 shows the temperature dependencies of free energy of 6H, 3C, and 2H SiC polytypes measured from 4H SiC. 4H SiC is most stable at low temperatures, but 6H SiC is most stable at higher temperatures. 3C SiC is less stable against 4H or 6H SiC except at very high temperature region. Those results are consistent with the other speculations but the precisions of the calculations are not enough. Although the more precise calculations will alter the results of hierarchy of polytypes, their result pointed out the possibility of the phase transition in the Si-C system from the first principles calculations.

5. Conclusions

We have utilized a new method for manufacturing SiC from liquid Si; in this method, single crystals of 4H-SiC are obtained from polycrystalline 3C-SiC source in the absence of a temperature gradient. The origin of the driving force for crystal growth is the same as that in the case of diamond synthesis from a metal-carbon solvent, and it is elucidated by considering the stable-metastable double phase diagrams. This similarity in the growth mechanism indicates that the methods developed for diamond synthesis can be directly used for growing large-size SiC crystals from a metastable solvent of Si.

6. References

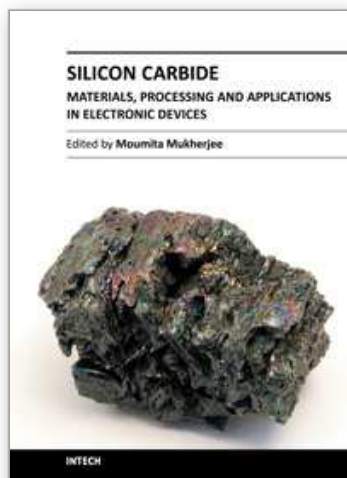
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Silicon Carbide (SiC) and its polytypes, used primarily for grinding and high temperature ceramics, have been a part of human civilization for a long time. The inherent ability of SiC devices to operate with higher efficiency and lower environmental footprint than silicon-based devices at high temperatures and under high voltages pushes SiC on the verge of becoming the material of choice for high power electronics and optoelectronics. What is more important, SiC is emerging to become a template for graphene fabrication, and a material for the next generation of sub-32nm semiconductor devices. It is thus increasingly clear that SiC electronic systems will dominate the new energy and transport technologies of the 21st century. In 21 chapters of the book, special emphasis has been placed on the “materials” aspects and developments thereof. To that end, about 70% of the book addresses the theory, crystal growth, defects, surface and interface properties, characterization, and processing issues pertaining to SiC. The remaining 30% of the book covers the electronic device aspects of this material. Overall, this book will be valuable as a reference for SiC researchers for a few years to come. This book prestigiously covers our current understanding of SiC as a semiconductor material in electronics. The primary target for the book includes students, researchers, material and chemical engineers, semiconductor manufacturers and professionals who are interested in silicon carbide and its continuing progression.

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